

The role of manganese in CoMnO_x catalysts for selective long-chain hydrocarbon production via Fischer-Tropsch synthesis

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Cobalt is an efficient catalyst for Fischer–Tropsch synthesis (FTS) of hydrocarbons from syngas ($\text{CO} + \text{H}_2$) with enhanced selectivity for long-chain hydrocarbons when promoted by Manganese. However, the molecular scale origin of the enhancement remains unclear. Here we present an experimental and theoretical study using model catalysts consisting of crystalline CoMnO_x nanoparticles and thin films, where Co and Mn are mixed at the sub-nm scale. Employing TEM and in-situ X-ray spectroscopies (XRD, APXPS, and XAS), we determine the catalyst's atomic structure, chemical state, reactive species, and their evolution under FTS conditions. We show the concentration of CH_x , the key intermediates, increases rapidly on CoMnO_x , while no increase occurs without Mn. DFT simulations reveal that basic O sites in CoMnO_x bind hydrogen atoms resulting from H_2 dissociation on Co^0 sites, making them less available to react with CH_x intermediates, thus hindering chain termination reactions, which promotes the formation of long-chain hydrocarbons.

The Fischer–Tropsch synthesis (FTS) reaction converts syngas ($\text{CO} + \text{H}_2$) to hydrocarbons, with cobalt (Co) being one of the most efficient catalysts^{1–3}, with the most desirable reaction products being long-chain hydrocarbons, rather than methane⁴. Numerous studies have shown that the product selectivity for C_{5+} and longer hydrocarbons can be enhanced by addition of manganese (Mn)^{5–25}. Investigation by Weckhuysen and co-workers revealed the critical role played by $\text{MnO}/\text{Mn}_{1-x}\text{Co}_x\text{O}$ phases and by the reducible TiO_2 support

in improving selectivity towards long-chain hydrocarbons^{7–10}. Bell et al. further elucidated the significance of the Lewis acid–base interaction in promoting the activation of CO at the Co–MnO interface^{12–16,22}. Kruse et al. reported that formation of alcohols and aldehydes can be favored on CoMnO_x catalysts^{17,19–21}, proposing a synergistic effect between Mn_2O_3 oxide and cobalt carbide (Co_2C) driving the reaction^{17,21}. Sun et al. proposed that Co_2C is an active phase of CoMn catalyst for the selective conversion of syngas to

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lower olefins, with Mn acting as a structure promoter in the formation of Co_2C nano-prisms from the $\text{Co}_x\text{Mn}_{1-x}\text{O}$ precursor^{18,24}. So far however, the nature of the catalytic active phase in Co/Mn compounds, and the molecular scale mechanism of the FTS reaction is still under debate. To advance in this field, a better knowledge of the atomic scale structure, chemical state, and nature of the intermediate species present on the catalyst surface is necessary, a goal hampered by the heterogeneity of the widely used powder catalysts that make atomic-level investigations difficult.

To resolve this problem, we used two model catalysts^{26,27}: one in the form of crystalline nanoparticles (NPs) of CoMnO_x , the other in the form of amorphous few-nanometers-thin films of Co, MnO, and CoMnO_x , grown by physical vapor deposition on silicon wafers and on silicon nitride (SiN_x). An important feature of our catalysts is that in the CoMnO_x NPs and in the thin films the atomic components are mixed at the sub-nanometer (sub-nm) scale, which we verify by TEM and X-Ray diffraction, rather than segregate in different domains sharing boundaries when in contact. Our model catalysts facilitate microscopic and spectroscopic studies under reaction conditions, which reveal the structure of the active sites and the special role of Mn. Our spectroscopic techniques include Ambient Pressure X-ray Photoelectron Spectroscopy (APXPS), and X-ray Absorption Spectroscopy (XAS), in the presence of syngas reactants. The results show that the presence of Mn substantially enhances the formation of CH_x species, which explains the high selectivity for long-chain hydrocarbon products. With DFT calculations we explain the role of MnO_x

in controlling the chemical potential of H, which binds to basic O sites of MnO and thus reducing the supply needed for chain termination reactions, while the increase in the concentration of CH_x species favors C-C coupling and chain growth, in agreement with Weckhuysen et al.¹⁰.

Results and discussion

Synthesis and characterization of CoMnO_x model catalysts

Crystalline nanoparticles of mixed Co and Mn oxide (CoMnO_x NPs) were synthesized by hot injection of $\text{Co}_2(\text{CO})_8$ and $\text{Mn}_2(\text{CO})_{10}$ into a solution of octyl-ether in the presence of oleic acid^{28–30}. The bulk composition of the NPs was measured by Inductively Coupled Plasma Optical Emission Spectrometry (ICP-OES). High-resolution scanning transmission electron microscopy (STEM) shows that the CoMnO_x NPs are crystals of ~10 nm diameter (Fig. 1a). Two atomic plane distances of 0.248 nm and 0.217 nm are observed, consistent with the (-111) and (002) lattice spacings of CoO (Fig. 1b), indicating that they are solid solutions of Mn in the CoO matrix. The initial oxidation state of CoMn is due to reaction with oleic acid and by exposure to air.

Thin films of CoMnO_x were also prepared, using sequential evaporation of Co and Mn onto a Si wafer and/or onto SiN_x TEM windows. High-angle-annular-dark-field (HAADF), STEM, and energy-dispersive X-ray spectroscopy (EDS) images of CoMnO_x NPs, and of CoMnO_x thin films after the activation process described below, are shown in Fig. 1c–j. The images show that Co and Mn are homogeneously mixed at the sub-nm scale.

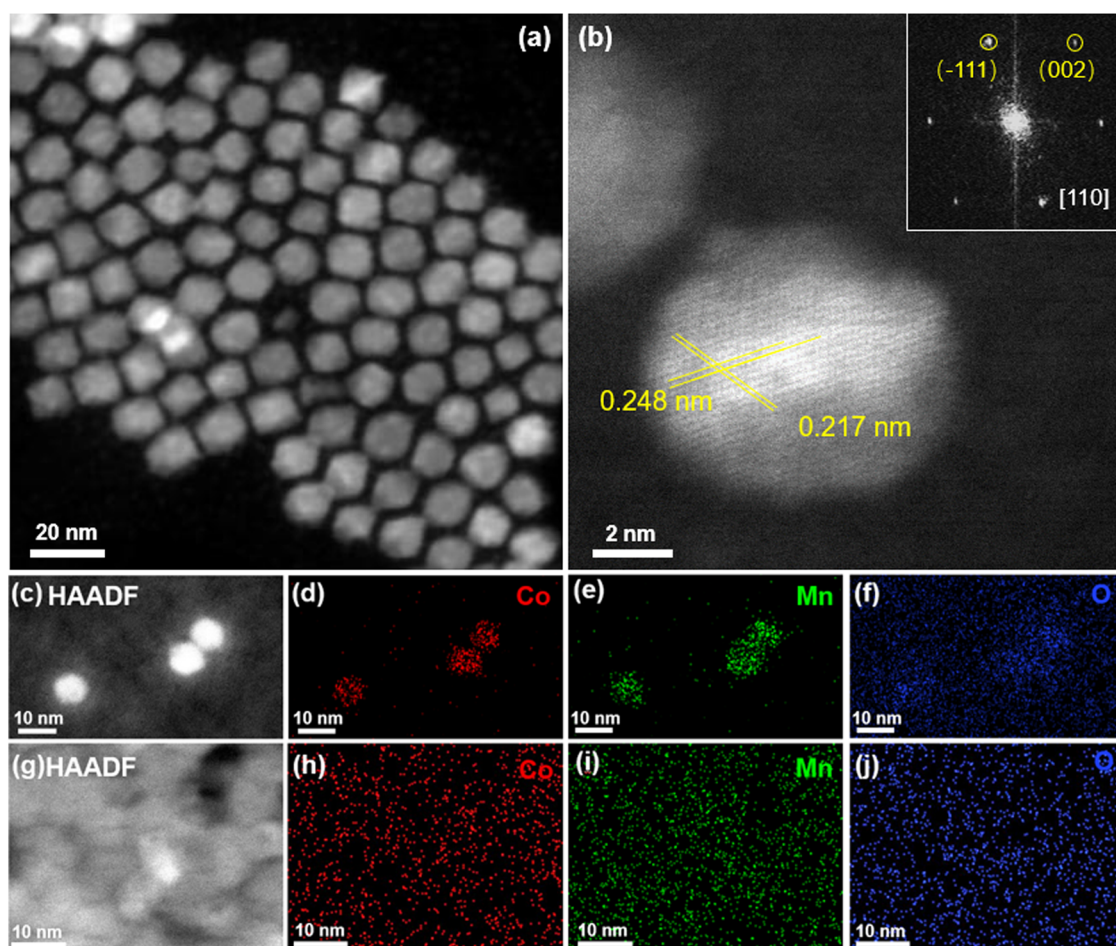


Fig. 1 | STEM-HAADF and EDS analysis of CoMnO_x NPs and thin film. **a** STEM-HAADF image of as-synthesized CoMnO_x crystal NPs; **(b)** High resolution image of a CoMnO_x NP, with an inset showing its FFT pattern; **(c–f)** STEM-HAADF images and EDS maps of CoMnO_x NPs; **(g–j)** STEM-HAADF image and EDS maps of a 10 nm thick

amorphous CoMnO_x film grown on a SiN_x membrane. The images **(c–j)** are acquired after reduction by heating to 450 °C in 1 Bar of H_2 . These images show that in both cases Co and Mn are intimately mixed at the sub-nm scale.

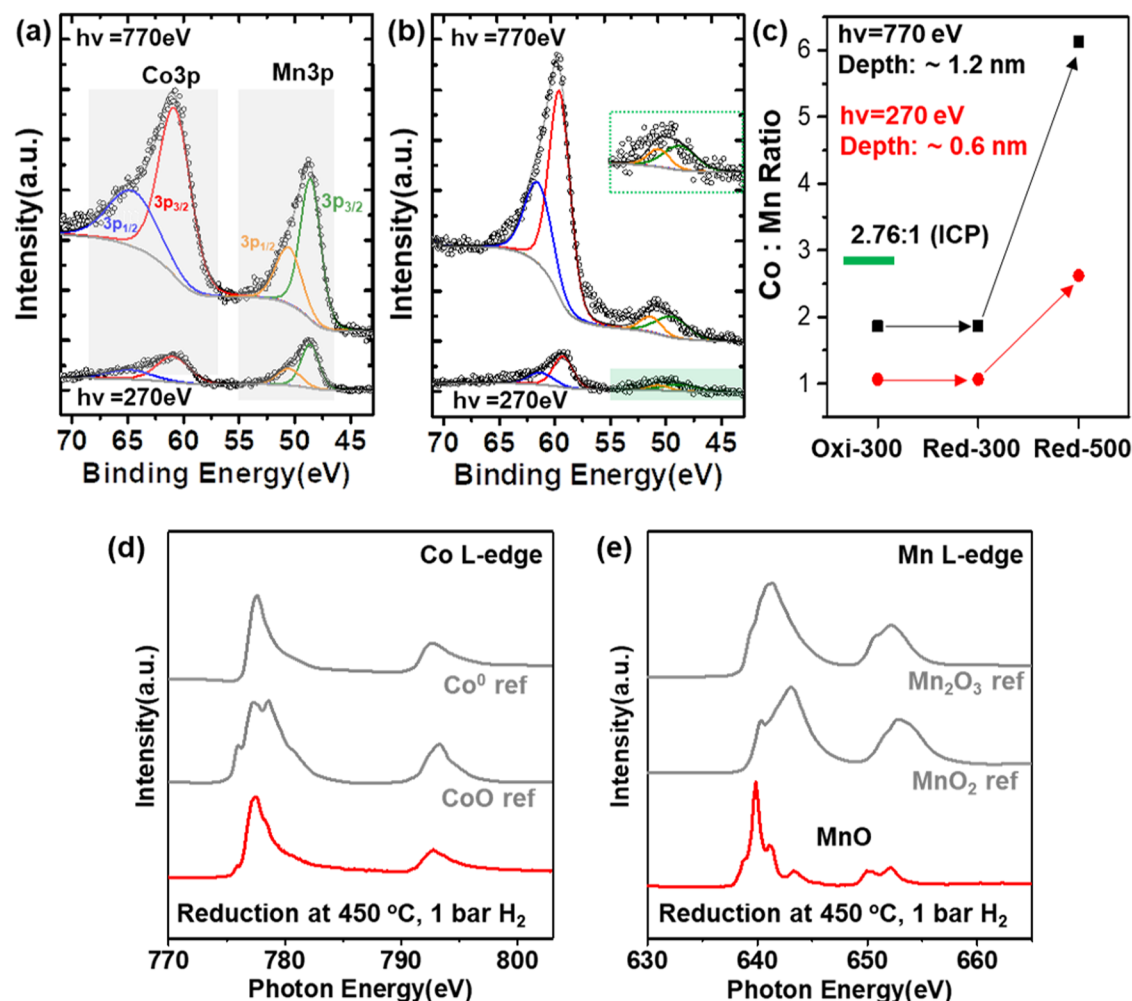


Fig. 2 | Surface composition of CoMnO_x NPs from Co3p and Mn3p XP spectra. **a** under 100 mTorr O₂ at 300 °C. **b** under 100 mTorr H₂ at 500 °C. Two photon energies, 770 eV and 270 eV were used (top and bottom spectra) to obtain a depth distribution profile based on the different mean-free paths of 1.2 and 0.6 nm respectively of the corresponding photoelectrons. The colored curves are fitting peaks for the 3p_{1/2} and 3p_{3/2} components for Co and Mn. The inset in the green box

shows an expanded view of the Mn 3p XPS (hv = 270 eV) region in (b). **c** Co:Mn ratio in the top surface region (~0.6 nm, red points) and near surface region (~1.2 nm, black points) after oxidation and reduction treatments. The top surface of the activated catalyst contains Co⁰ and Mn²⁺ in a 2.5:1 ratio. **d, e** Co and Mn L-edge total electron yield XAS from CoMnO_x NPs (red), after activation by heating in 1 bar of H₂ at 450 °C. Spectra from known compounds (gray) are shown for comparison.

In-situ X-ray Diffraction (XRD) patterns of the CoMn NPs, shown in Fig. S1, reveal the CoMn spinel structure after O₂ calcination at 300 °C, which produced MnO after reduction in H₂ at 450 °C, suggesting segregation of Co from the bulk to the surface region, consistent with the in-situ XPS results discussed below.

Catalytic activity measurements

The FTS activity of the activated CoMnO_x NPs model catalysts was measured using a fixed-bed reactor described in the Methods section. Product distributions are shown in Fig. S2 for two Co/Mn composition ratios (left), and for pure Co on Al₂O₃ (right), along with the prediction from the Anderson-Schulz-Flory (ASF) model for a chain growth probability α of 0.5¹. Under steady state of 10% CO conversion, 220 °C, 1 bar, and H₂/CO ratio of 2, the CoMnO_x NPs catalysts with a Co:Mn ratio between 1.6 and 2.8 displayed higher selectivity for production of hydrocarbons with chain lengths ≥ 5 carbon units (C₅₊: 48%, CH₄: 18%), than Co/Al₂O₃ catalyst (C₅₊: 25%, CH₄: 38%). The product distribution on these CoMnO_x NPs obeys the ASF distribution with $\alpha = 0.76$, in agreement with previous results^{6,23}. These results show that our CoMnO_x NPs model catalysts perform similarly to industrial powder-form catalysts in terms of product distribution, with enhanced

percentage of hydrocarbon chains lengths of more than 5 carbons units. Similar catalytic activity measurements could not be carried out on the CoMnO_x films due to the large difference in catalyst area exposed to reactants, which was nearly 4 orders of magnitude higher for the NPs than for the films. However, in spite of this difference, as we show in the following sections, that the elementary steps of the FTS reaction are the same in the thin films and in the nanoparticles.

Catalyst surface structure after activation

APXPS characterization results from CoMnO_x NPs after activation, consisting of heating at 300 °C in 100 mTorr O₂, to remove contaminants, and subsequent heating to 500 °C in 100 mTorr H₂ are shown in Fig. 2. The oxidation process resulted in a mixture of Co²⁺, Co³⁺, Mn²⁺, and Mn³⁺ states as in Co₃O₄ and Mn₃O₄ respectively, as shown by comparing the XAS results of our catalyst with standard compounds (Fig. S3a, b). The reduction was carried out under 1 Bar of H₂, but the APXPS experiments were carried out in the vacuum chamber where the H₂ pressure was 100 mTorr. The structure of the activated CoMn was the same in both. A composition depth profile of the surface region was obtained from the intensities of the Co 3p and Mn 3p XPS peaks acquired at photon energies of 770 eV and

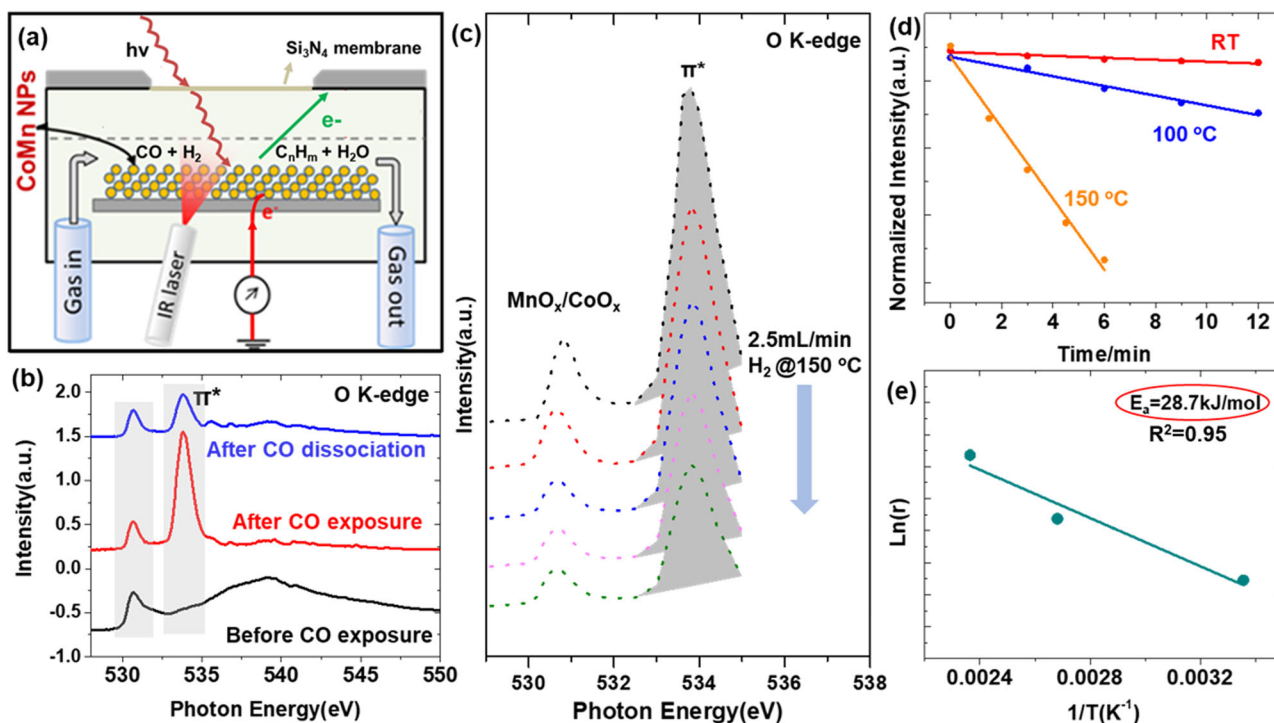


Fig. 3 | Reactions between CO and H₂ on CoMnO_x crystal NPs. **a** Schematic illustration of the reaction cell used for operando reaction studies using TEY-XAS. The yellow circles represent CoMnO_x NPs on a gold foil covering a Si wafer. **b** O K-edge XAS of CoMnO_x before (black), after CO exposure (red), and after H₂ introduction (blue). The peak at 531 eV is due to lattice O in CoMnO_x. The peak at 534 eV corresponds to the X-rays excitation of electrons from the O 1s level of CO to the antibonding orbital (π^*). This peak has contributions from the molecular CO adsorbed on Co⁰, and from carbonates on CoMnO_x. The broad peak around 540 eV

is due to transitions to unoccupied σ^* bands³⁵, which disappear due to adsorbed CO donating electrons to these orbitals. After introduction of H₂, the peak at 534 eV decreases due to H₂-assisted CO dissociation, leaving only contributions from the carbonate species. **c** Time evolution (2 min intervals) of the CO π^* peak intensity under 1 bar of H₂ flowing at 2.5 mL/min at 150 °C. **d** Intensity of the CO π^* orbital peak as function of time for three temperatures, which measures the H₂-CO reaction rate. **e** Arrhenius plot from the π^* orbital peak intensity decay rate.

270 eV (top and bottom spectra in Fig. 2a, b), which generate photoelectrons of -700 eV and -200 eV energies with mean free paths of -1.2 nm and -0.6 nm, respectively. We refer to these as 'near surface' and 'top surface' regions. The red/green traces for the 3p_{3/2}, and blue/orange traces for the 3p_{1/2} indicate their spin-orbit components. For CoMnO_x NPs with Co:Mn ratio of 2.76:1 (from ICP-OES), the peak intensities after oxidation indicate a Co:Mn ratio of 1:1 for the top surface region, and 2:1 for the near surface region, indicating surface enrichment of Mn. The reduction under 100 mTorr of H₂ at 300 °C brought no appreciable changes in surface composition (Fig. S4a), and only chemical changes of Co and Mn, which were both reduced to the 2+ state (Fig. S4b, c). However, when the temperature was raised to 450 °C and higher, the surface composition changed significantly, as indicated by the increase of Co 3p XPS peak intensity (Fig. 2b), which now corresponds to a Co/Mn ratio of 6 in the near surface region and 2.5 in the top surface region, respectively (Fig. 2c). The higher temperature reduction changed the oxidation state of Co from 2+ to metallic, while Mn remained in the 2+ state, as can be seen in the XAS of the samples, along with reference spectra from Co, CoO, Mn₂O₃, and MnO₂ (Fig. 2d, e). STEM-EDS maps demonstrate the preservation of the CoMnO_x bulk structure during oxidation (Fig. S3c) and reduction pretreatment (Fig. 1d-f, h-j), indicating that the pretreatments mostly affect the redistribution of Co and Mn in the surface region. The surface composition of reduced CoMnO_x was also confirmed by the C1s APXPS results using CO as probe molecule which adsorbs molecularly only on Co⁰ sites and becomes carbonate on MnO sites (Fig. S5). The two-step reduction pattern is similar to that reported for CoMnO_x catalysts prepared by co-impregnation^{8,31}.

Reactions between CO and H₂ on CoMnO_x NPs, CoMnO_x, MnO, and Co films

To study the reaction between CO and H₂ on CoMnO_x, we used a reaction cell like that shown schematically in Fig. 3a. The cell is closed by a 100 nm thick Si₃N₄ window that separates the volume inside, filled with reaction gases, from the beamline vacuum chamber. The CoMnO_x NPs, were deposited on a gold foil on top of a Si wafer. They were first reduced by exposure to 1 bar of a 10% H₂/Ar mixture at 450 °C, heated by an 808 nm IR photodiode laser via an optical fiber from the back, and then cooled to reaction temperature^{28,32}. The NPs chemical composition was determined by XAS recorded in the Total Electron Yield (TEY) mode by measuring the sample to ground current²⁸, which is surface sensitive to a depth of few nm^{33,34}. The O K-edge XAS of the activated sample before CO introduction, is shown by the bottom black curve in Fig. 3b. The peak at ~531 eV corresponds to 1s-level electrons from lattice O in CoMnO_x excited to states at the bottom of the oxide conduction band^{28,35}, and the broad peak near 540 eV to excitation to higher unoccupied σ^* bands of CoMnO_x³⁵. After exposure to CO, which adsorbs molecularly only on Co⁰ sites, an intense peak at ~534 eV is observed, corresponding to the excitation of O 1s-level electrons of CO to the empty π^* antibonding orbital of the molecule²⁸ (Fig. 3b, red curve). On oxide sites, however, CO adsorbs forming carbonates, which also contribute to the 534 eV peak³⁵. After flowing H₂ at 2.5 mL/min at 1 Bar pressure for 12 min with the sample at 150 °C, the intensity of the peak at ~534 eV dropped significantly as a function of time (Fig. 3c) due to the H₂-assisted CO dissociation reaction²⁸. The remaining peak, after all molecular CO has dissociated (Fig. 3b, blue curve), corresponds to unreacted carbonates on CoMnO_x³⁵. Figure 3d shows the evolution of the π^* peak intensity as a function of reaction

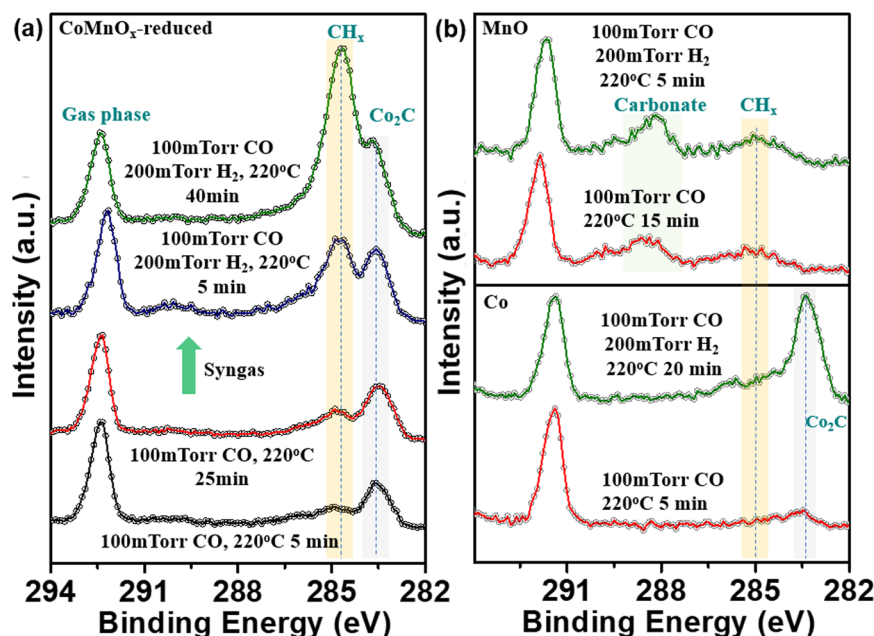


Fig. 4 | XPS showing surface carbonaceous species on CoMnO_x, Co, and MnO thin films. **a** From the bottom: activated CoMnO_x under 100 mTorr CO at 220 °C for 5 min (black), after 25 min (red), and under syngas (100 mTorr CO + 200 mTorr H₂) for 5 min (blue) and for 40 min (green). The absence of a peak at -286.0 eV from adsorbed CO is due to its low coverage at 220 °C resulting from desorption/equilibration with the gas phase, and from dissociation by reaction with H₂.

Most noticeable is the increase in intensity of the peak from CH_x species at -285 eV. **b** Bottom: metallic Co film exposed to 100 mTorr CO at 220 °C for 5 min (red), and to syngas (100 mTorr CO and 200 mTorr H₂) for 20 min (green). CO dissociation by H₂ produces cobalt carbide (-283.5 eV) on Co but no stable CH_x. Top: pure MnO film under 100mTorr CO at 220 °C for 15 min (red) and under syngas for 5 min (green) showing a smaller peak at -285 eV, likely due to adventitious contamination.

time for 3 temperatures: RT, 100 °C, and 150 °C. The linear relationship indicates a constant reaction rate. An Arrhenius plot of the rate, Fig. 3e, gives an activation energy of -28 kJ/mol, a value substantially smaller than that of the typical FTS reaction (-80 kJ/mol)⁶, showing a reduction of the barrier for H₂ induced CO dissociation on the Co-MnO surface. Besides, the H₂-assisted carbonate dissociation reaction is unlikely to occur. According to the DFT simulations, the adsorption of CO as carbonate is much less likely than molecular adsorption on the Co phase of the catalyst (-0.28 eV versus -1.66 eV). Therefore, these carbonates adsorbed on the oxidic part are likely spectators, as any catalytic cycle involve carbonate will require the activation from the metallic Co phase with an energetic cost of at least 1.38 eV. It agrees with the in-situ O 1s XAS results as after the H₂ introducing, the unreacted carbonates still remain on CoMnO_x catalyst surface.

In addition to XAS, the chemical state of the CoMnO_x NPs under FTS reaction conditions was followed by APXPS. After reduction by exposure to 100 mTorr H₂ at 500 °C (Fig. S6a, black curve), the spectra show that most of the Co atoms in the top surface region (-4.6 Å), are in the metallic state, while within the same depth, (Fig. S6b), Mn remains in the 2+ state. Following exposure of the reduced sample to 100 mTorr of CO at 220 °C, a fraction of the metallic cobalt was oxidized to Co²⁺ due to CO dissociation²⁸, as indicated by the increased shoulder intensity at -780.2 eV from Co²⁺ (Fig. S6a, red curve). Under 300 mTorr of syngas at a ratio CO:H₂ = 1:2, the oxide peak increased substantially and becomes the dominant peak (Fig. S6a, blue curve). The increased oxidation is the result of H₂-assisted dissociation of CO, confirmed by Co²⁺ peak in the Co L-edge XAS (Fig. S7a). During this period, no significant chemical change in the MnO was observed (Figs. S6b, S7b). More important, a depth profile using the XPS peaks of Co 3p and Mn 3p (Fig. S7c, d) demonstrates that under syngas reaction conditions at 220 °C, the composition of top and near surface regions remains unchanged during reaction, indicating the stability of the Co-MnO structure. It also shows that oxidized Co, from H₂-assisted CO dissociation, is subsequently reduced by H₂. This structural stability

under the FTS reaction conditions was further confirmed by in-situ XRD in Fig. S1 and ex-situ TEM in Fig. S8, which demonstrates that although some morphological changes occur, the Co and Mn atoms remain intimately mixed at the sub-nm scale.

The nature of the species formed on the catalyst surface during the reaction of H₂ with CO on each Co, MnO and CoMnO_x films is further revealed by APXPS in the C 1s region, shown in Fig. 4. Starting with a 100 mTorr of CO gas environment, when the temperature was raised to 220 °C, the carbonates at -289.0 and -289.5 eV desorbed from on MnO and CoMnO_x (Fig. S5), while a peak at -283.5 eV from Co carbide increases noticeably on Co and somewhat less on CoMnO_x, where there is an apparent increase due to its overlap with the -285 eV peak from CH_x species. These C atoms for the reactions originate from H₂-assisted CO dissociation²⁸. The most remarkable change is the growth of the carbon peak at -285 eV due to CH_x species, which occurs only on CoMnO_x but not on Co or MnO. There is however a small contribution to that peak in Co and MnO from contaminant species from background gases^{36–40} that produce species with peaks between 284–285 eV. After 40 min of reaction, top green curve in Fig. 4a, the CH_x peak is the dominant and most stable carbonaceous species on the surface, as shown by the higher temperature, above 400 °C, needed to desorb it, as shown in the sequential spectra acquired versus time (Fig. S9). The increase in concentration, and higher thermal stability of the CH_x species indicates that it belongs to the longer chain intermediates formed by the chain-growth reaction^{41,42}, because longer chain molecules bind more strongly to the surface than shorter ones, with methane binding the weakest. Highly significant is also the observation that under the same conditions, no significant growth of CH_x species is observed on pure metallic Co or on MnO (Fig. 4b).

Theoretical simulations

To better understand the molecular scale origin of the higher selectivity of CoMnO_x catalysts towards C₅₊ hydrocarbon products in the FTS reaction, we performed DFT simulations on a model catalyst with a

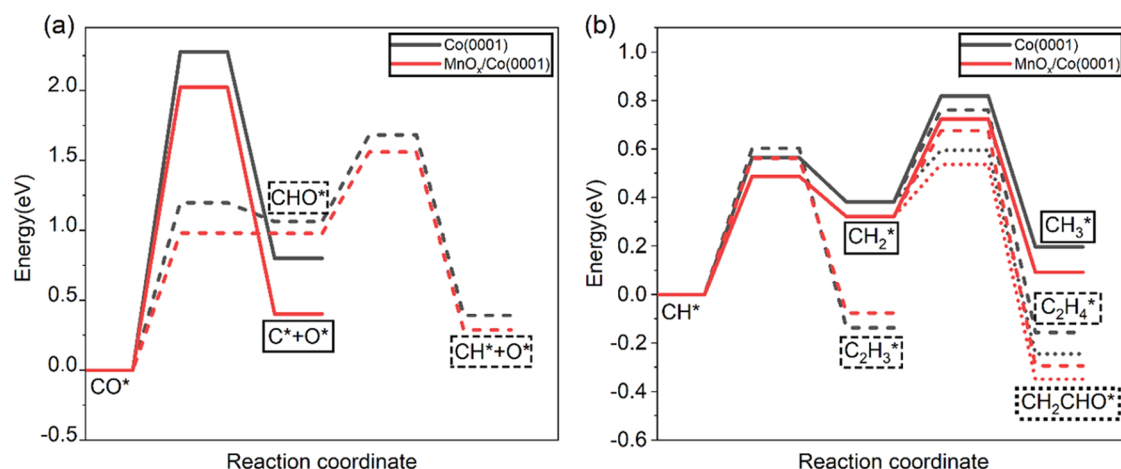


Fig. 5 | Energy diagrams. **a** CO dissociation reaction: Solid lines for thermally activated process; dashed lines for H₂-assisted process; The black line is for pure Co, and the red for MnO_x/Co(0001). **b** C-H and C-C bond formation reactions: Solid

lines for CH* + H* reactions. Dashed lines for CH_x* + CH₂* reactions. Dotted lines for CH₂* + CHO* reaction. Relative energies between elementary steps were shifted to align the initial state for energy barrier comparisons.

structure and composition based on the experimental data described above: a Co:Mn ratio of ~2.5 in the topmost layer with Co in the metallic state (Co⁰), and Mn with a double positive charge (Mn²⁺). A model that fits this ratio and oxidation state is one where the components Co, Mn, and O are arranged in a compact structure. And interfacial structure was determined by combining global optimization (See SI computational details) and the calculation of stability under reaction conditions (Fig. S10), as that illustrated in Fig. S10a (center) with a unit cell in the topmost layer containing 12 Co atoms, 4 Mn atoms and 4 O atoms. A Bader charge analysis shows that the Mn in the model has a charge of +1.20 |e|, compared with that in MnO (+1.47 |e| for Mn in the bulk, and +1.39 |e| for Mn on the surface). For comparison the Mn charge in Mn₂O is +0.72 |e| for Mn in the bulk. The average charge of surface Co atoms is +0.00 |e|, and the average charge of the Co atoms near MnO_x is +0.07 |e|, indicating that Co is mostly metallic but slightly polarized near the MnO interface. The projected density of states (PDOS) of the Co d-band in Fig. S11, shows only a small difference between that from Co atoms in Co(0001) and in CoMnO_x. This structure is stable against changes in oxidation state by either gain or loss of oxygen atoms illustrated in Fig. S10a–c, respectively.

The adsorption energies of relevant key fragments were then calculated on three simple surfaces (Table S1): MnO(100), metal Co(0001), and our model MnO_x/Co(0001). The adsorption of H*, CO*, CH*, and CH₂* on MnO(100) was found to be weaker than on Co(0001) by 1.26 eV, 1.64 eV, 2.54 eV, and 1.86 eV, indicating that the species prefer to adsorb on Co instead of MnO. In the MnO_x/Co(0001) model, 6 sites near the interface are considered, as shown in Fig. S10a, which are top sites (marked by circles), hcp and fcc 3-fold sites (marked by down-pointing and up-pointing triangles). On MnO_x/Co(0001), most of the adsorbates bind preferentially to the Co sites instead of MnO sites with an adsorption energy only slightly different from that in pure Co(0001). Some carbon fragments (CH* and C₂H₄*) do not adsorb on MnO_x sites of the MnO_x/Co(0001) model, while for CO* and CH₂*, the adsorption is more endothermic than on the Co sites by 1.60 eV and 1.45 eV. For H*, O*, and C* the adsorption is only moderately more endothermic than on the Co sites by 0.15 eV, 0.37 eV, and 0.25 eV, respectively. The diffusion of H from Co to O sites on MnO_x/Co(0001) was found to be sensitive to the local interface structure when the oxygen content changes. For example, the energy barrier of H diffusion from Co to O sites near the interface changes from 1.47 to 0.85 eV (Fig. S12), indicating that the diffusion is easier on configurations with certain local oxygen environments.

DFT calculated energy diagrams for the reaction steps are shown in Fig. 5. In panel (a) the continuous lines correspond to thermal CO dissociation, while the broken lines are for H₂-assisted dissociation. The continuous lines are for reactions on Co sites in Co(0001) (black), and for Co sites in the MnO_x/Co(0001) surface (red). The energy barrier of the rate determining step (RDS) for H₂-assisted dissociation on Co(0001) is 1.20 eV, while the energy barrier of CO thermal dissociation is 2.28 eV. On MnO_x/Co(0001), both energy barriers are decreased, the former to 0.98 eV, and the latter to 2.03 eV.

A charge density difference analysis shows that CO has a stronger interaction with Co on the MnO_x/Co(0001) surface than with Co on the Co(0001) surface (Fig. S13). Assuming the pre-exponential factor of the reaction rate to be the same for the Co sites on Co(0001) and Co sites on MnO_x/Co(0001), the reaction rate on MnO_x/Co(0001) is two orders of magnitude higher than on pure Co(0001) at -220 °C. Therefore, the presence of the interface will facilitate CO dissociation and increase CHO* formation, which further decomposes to CH*.

Reactions involving carbon fragments can proceed in two different ways (Fig. 5b): (i) oligomerization, i.e., C-C coupling, and chain growth, or (ii) termination by hydrogenation of the carbon moieties. The competition between these two processes leads to the ASF distribution⁴³. The hydrogenation of CH* and CH₂*, and the coupling of CH* + CH₂*, CH₂* + CH₂*, CH₂* + CHO*, reactions were calculated to understand the effect of Mn in the formation of C-H and C-C bonds. The CH₂* + CHO* coupling reaction has an energy barrier of 0.21 eV on Co(0001), and 0.22 eV on MnO_x/Co(0001), indicating that C-C coupling can occur through the reaction, depending on the relative population of both species. Since CHO* easily dissociates into CO* and H*, the coverage of the intermediate CHO* will be low, leading to a low rate of this bimolecular reaction, and for this reason the reaction is not discussed further in the following. In general, the C-C formation releases more energy than the C-H formation, indicating that C-C bond formation is thermodynamically favorable (Fig. 5b, Table S2). Starting from CH* fragments, the energy barrier towards C-H formation (0.57 eV on Co, 0.49 eV on MnO_x/Co(0001)) is lower than the C-C formation (0.60 eV on Co(0001), 0.56 eV on MnO_x/Co(0001)). However, once CH₂* is formed, the energy barrier for C-C coupling (0.38 eV on Co, 0.36 eV on MnO_x/Co(0001)) is lower than C-H formation (0.44 eV on Co(0001), 0.40 eV on MnO_x/Co(0001)), indicating that C-C formation is likely to occur with CH₂*.

From the simulations, the following picture of the reaction emerges (Fig. 6) that explains the impressive performance of CoMnO_x. First, syngas species adsorb on the metal and H₂ dissociates, with some

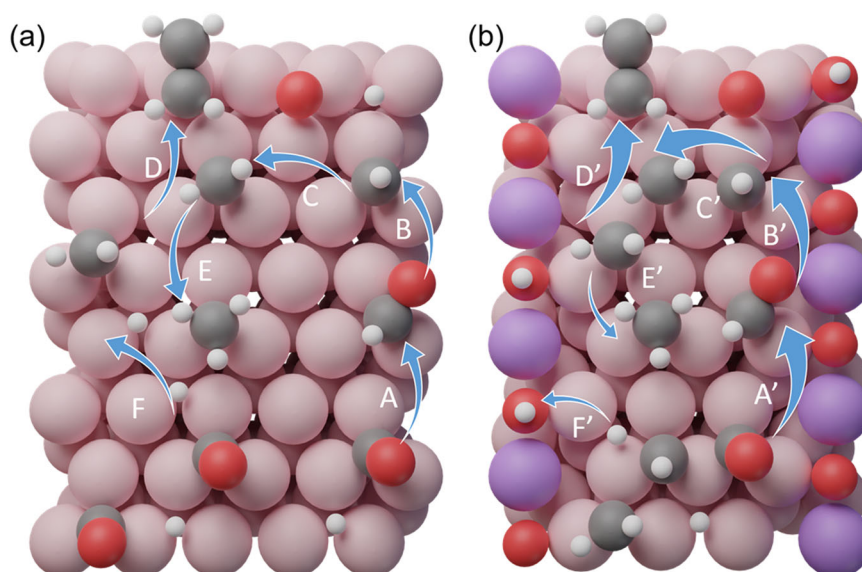


Fig. 6 | Illustration of the proposed reaction mechanism. **a** on Co(0001) and **(b)** on CoMnO_x. Co atoms shown in pink, Mn in purple, O in red, H in white, and C in gray. The arrows' size follows qualitatively the value of the reaction rate for

comparison on different catalysts. A more detailed illustration of the reaction steps is shown in Fig. S16.

H atoms being captured on O sites in MnO_x to form hydroxyls species (F' in panel b), in agreement with the appearance of a shoulder in the XPS peak of O 1s at -531 eV (Fig. S14). Hydroxyls from the Cobalt surface may also contribute to this peak⁴⁴. Once the H atoms bind to basic centers (O) of the CoMnO_x system, they are sequestered and unavailable for C-H formation (E') (Figs. 6 and S15). The interface shows good activity in activating CO via the H₂-assisted pathway (A'-B'), which increases the rate and overall activity of formation of CH_x* species (C'). Second, chain growth steps (D') are more likely on CoMnO_x compared to pure Co, from both coverage and energetic considerations: CH_x* adsorption on MnO is unfavorable, therefore facilitating the increase of CH_x* coverage on Co. In addition, the barrier for C-C coupling is lower on CoMnO_x than that on pure Co, which favors chain growth. The step-termination reactions are less likely on CoMnO_x because a fraction of H atoms are bound to the oxide and some consumed to remove some of the O generated by CO dissociation. Thus, the MnO_x acts as a buffer for the chemical potential of both oxygen and H under reaction conditions making them optimal to reduce side reactions.

It has been proposed in recent literature that Co₂C, in the form of nanoprisms exposing (101) and (020) facets, are active for selective olefin conversion¹⁸. However, the CoMnO_x NP crystal catalysts used here mainly produces C₅₊ hydrocarbons, indicating that the Co₂C, present in small amounts in pure Co and even less in CoMnO_x, as shown by the C 1s APXPS in Figs. 4 and S5, has a negligible contribution to the FTS products in our case. To explain this we performed DFT calculations using the stoichiometric surface Co₂C(011) surface, which has the lowest surface energy⁴⁵. The results, shown in Fig. S17, indicate that for all the pathways shown in Fig. 6, the energy barriers on Co₂C(011) are higher than on Co(0001).

In summary, by using well-defined CoMnO_x crystalline nanoparticles and amorphous thin films as model catalysts where Co and Mn are not separated but mixed at the sub-nm scale, together with in-situ spectroscopic characterization by APXPS and XAS under operating conditions, and with the help of DFT calculations, we unraveled the elemental distribution of Co and Mn atoms on the surface of the catalyst and their oxidation state, showing also that the bulk structure remains practically unchanged. The catalytically active phase of CoMnO_x under FTS reaction conditions is composed of metallic Co in

contact with MnO. The reaction proceeds by H₂-assisted CO dissociation and leads to a large increase in the production of CH_x* intermediates, much more than in pure Co, favoring chain growth reactions. The main advantage of CoMnO_x is that Mn remains always oxidized with its O providing basic sites that bind H, thus lowering the amount available for CH_x*+H* coupling, leading to chain termination and methane formation, while increasing the concentration of CH_x* species, which favors the C-C coupling between towards oligomerization and chain growth.

Methods

Preparation of model catalysts

The CoMnO_x NPs were synthesized as follows: a solution of 0.5 mL of oleic acid in 15 mL of dioctyl ether was stirred under vacuum and heated to 50 °C for 30 min. Next, a mixture of 205 mg of Co₂(CO)₈ and 117 mg of Mn₂(CO)₁₀ (corresponding to a feeding ratio of Co:Mn = 2:1) in 2 mL of dioctyl ether was injected in the solution. The solution was heated at 5–8 °C/min and kept at 290 °C for 1 h, followed by cooling down to room temperature, with isopropanol added to precipitate the nanoparticles. The nanoparticles were purified by 2 cycles of centrifugation and dispersed in hexane. This synthesis produced CoMn nanocrystals with diameters of ~10 nm and Co:Mn composition ratio of 2.76:1, as measured by Inductively Coupled Plasma Optical Emission Spectrometry (ICP-OES).

CoMnO_x thin film were prepared by sequential evaporation of Co and Mn onto a Si wafer and onto SiN_x TEM windows and heated to 300 °C under 1 bar O₂ atmosphere. Their thickness, determined from Quartz Crystal Microbalance (QCM) measurements, was 100 nm for the Si wafer substrates and 10 nm for the silicon nitride TEM windows, respectively. Other two film compositions, Co and Mn, were prepared side-to-side on the same Si wafer. In this manner comparison of their structure under the same conditions, and under the same oxidation/reduction procedures as those used for the CoMnO_x NPs, was possible.

TEM characterization

The as-prepared CoMnO_x NPs were drop-cast onto a SiN_x TEM window (10 nm thickness, VWR Scientific). High resolution TEM studies were performed in an aberration corrected FEI Titan 80300 operated at

300 kV and equipped with a CEOS GmbH double hexapole aberration corrector providing angstrom level resolution in scanning imaging modes. Prior to TEM characterization, the CoMnO_x NPs underwent activation, consisting of two steps: first removal of the carbon by heating in a gas flow of 20% O₂/Ar mixture, at 25 mL/min, 1 bar, at 300 °C for 2 h. The second step is a reduction in a flow of 20% or 10% H₂/He mixture at 25 mL/min, 1 bar, for 2 h. The reduction step was done at two different temperatures of 300 °C and 450–500 °C.

FTS catalytic activity measurements

The FTS catalytic activity of the CoMnO_x NPs was tested using 70 mg of NP particles supported on alumina powder by sonication with a loading of 10 wt%. The fixed-bed reactor used was heated to 450 °C with a ramp of 5 °C /min for 2 h in flowing H₂ and subsequently cooled to 220 °C in flowing He. After activation the gas flow or was gradually switched from He to syngas over 10 min. The syngas was composed of H₂, CO (H₂/CO at a ratio of 2:1), and 7% Ar (Praxair, 99.999% purity) used as an internal standard. The catalyst temperature was measured using a K-type thermocouple positioned in the bed center. An Agilent 6890 N gas chromatograph was used to monitor the chemical composition of the reactor effluent. The gas chromatograph was equipped with a capillary column connected to a flame ionization detector to measure the hydrocarbon products, and a packed column connected to a thermal conductivity detector for measuring H₂, Ar, and CO.

In-situ APXPS and TEY-XAS characterization

For in-situ APXPS and TEY-XAS experiments, the CoMnO_x NPs were drop-cast onto an Au foil and activated as indicated above. The APXPS results were acquired in beamline 9.3.2 of the Advanced Light Source (ALS) of the Lawrence Berkeley National Laboratory (LBNL). TEY-XAS experiments were performed using a home-built gas cell in beamline 8.0.1.4 of the ALS with an energy resolution of 0.2 eV for the Co L-edge, Mn L-edge and O K-edge. The reaction cell is closed by a 100 nm thick Si₃N₄ window that separates the volume inside, filled with 1 Bar of reaction gases from the high vacuum chamber^{28,32}. The CoMnO_x NPs, supported on the Au foil, were heated using an 805 nm IR laser via an optical fiber that illuminates the sample from the back. Copper gas lines were used to introduce the CO and H₂ reactants with Cu carbonyl traps kept at 240 °C to remove any carbonyls present. All measurements were conducted under flowing gas conditions.

Data availability

The model system, structures of adsorption and reaction are available in ioChem-BD database⁴⁶ (<https://doi.org/10.19061/iochem-bd-6-272>).

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Author contributions

M.S. oversaw the project. H.C. and X.Z. conduct the in-situ APXPS/XAS/XRD experiment supported by the beamline scientists: M.B., A.H., I.W., H.O., J.G., and S.H.L. Data analysis were done by H.C. and M.S. Simulations were carried out by Z.L. and N.L. Sample preparation, characterization and reactivity measurement were done by H.C., P.F.P., X.Z., J.W., J.O.M., J.Y., J.S., E.P., S.C., L.T., N.L., Y.S., Z.Z., F.Y., E.W., and V.A. The paper was written by H.C., Z.L., N.L., and M.S., with contributions from P.Y. and H.Z. and other co-authors. A.T.B. helped to review and edit the manuscript.

Competing interests

The authors declare no competing interests.

Additional information

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