

# Dual transition metal electrocatalysis for difluoromethylation of Aryl halides using potassium difluoroacetate

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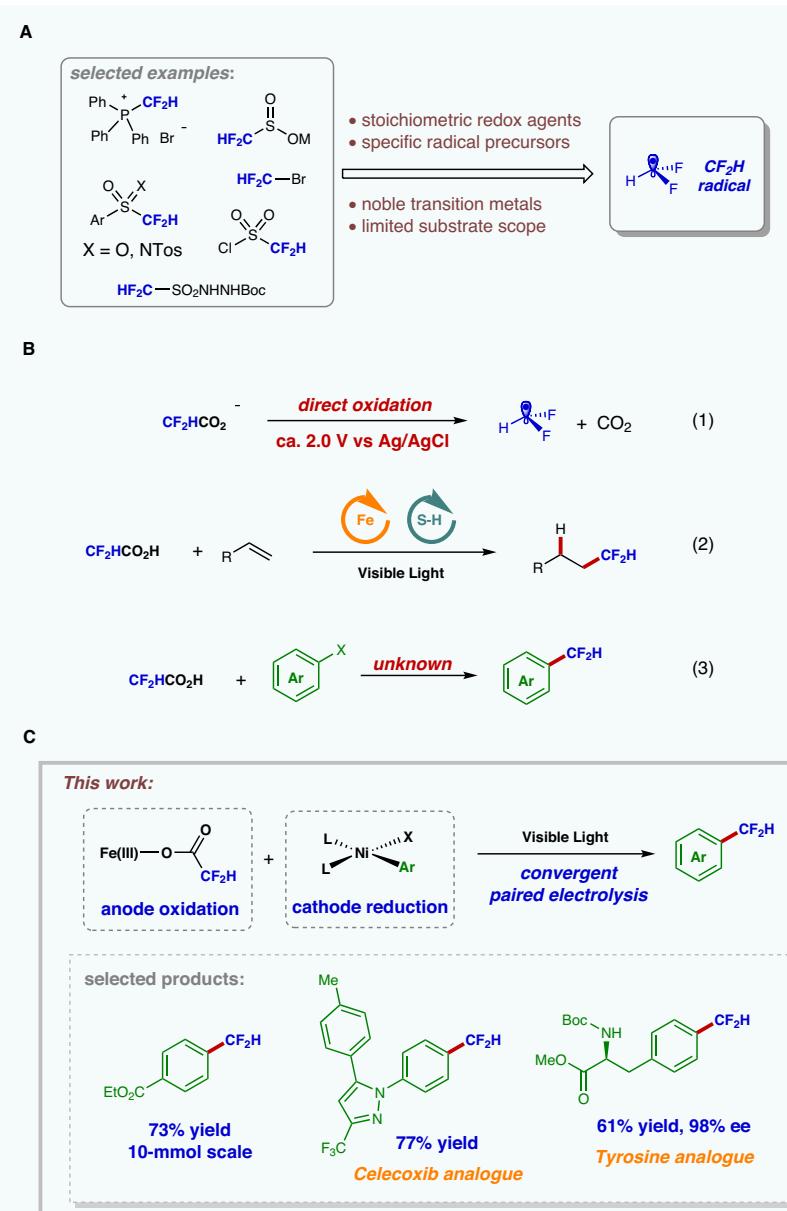
The development of robust, practical, and chemoselective methods for introducing the difluoromethyl ( $\text{CF}_2\text{H}$ ) group into organic molecules is highly sought after in the fields of pharmaceutical and agrochemical design. Herein, we report a Fe/Ni dual-transition-metal electrocatalytic strategy for difluoromethylation of (hetero)aryl halides, using difluoroacetate—the most abundant source of the  $\text{CF}_2\text{H}$  group—as an effective difluoromethylating reagent. A diverse array of aryl and heteroaryl halides, bearing synthetically useful functional groups, can be readily converted into the corresponding difluoromethylated products with good efficiency. This difluoromethylation protocol is readily scalable and is successfully applied to the preparation and late-stage functionalization of bioactive molecules.

Owing to the unique stability, physicochemical, and biological properties of fluorinated compounds, the development of new and efficient approaches for the installation of fluorinated alkyl groups into organic molecules remains highly desirable in the synthetic community<sup>1–7</sup>. In this context, the  $\text{CF}_2\text{H}$  group has recently attracted considerable attention in medicinal chemistry<sup>8–10</sup>. The highly polarized C–H bond within the  $\text{CF}_2\text{H}$  group endows it with a potent capacity to participate in hydrogen bonding interactions, making it a viable bioisostere for alcohol, thiol, and amino groups in drug molecular design. Moreover, the  $\text{CF}_2\text{H}$  group can also modulate lipophilicity and influence the conformation of molecules containing this functionality. As a consequence, the  $\text{CF}_2\text{H}$  group could be found in various biologically active compounds, such as drug candidates<sup>6</sup> and agrochemicals<sup>7</sup>, and there is a growing demand to develop sustainable and practical protocols for the introduction of the  $\text{CF}_2\text{H}$  group with readily available difluoromethylating agents<sup>11–22</sup>. As an elegant example, Xu and co-workers devised a new difluoromethylating reagent,  $\text{CF}_2\text{HSO}_2\text{NHNHBoc}$ , which is air-stable and can be readily obtained in one synthetic step from commercially available reagents. Under electrochemical ferrocene-catalyzed conditions,  $\text{CF}_2\text{H}$  radical can be generated without the need for chemical oxidants and has been successfully applied in an alkyne addition followed by a challenging 7-membered ring-forming homolytic aromatic substitution reaction for the synthesis of fluorinated dibenzazepines<sup>18</sup>. Indeed, difluoromethylation with  $\text{CF}_2\text{H}$  radical

constitutes an attractive strategy, and recent advances in this direction have led to the invention of a wide range of new difluoromethylating agents, enabling mild and selective ways to construct  $\text{C}(\text{sp}^3)\text{–CF}_2\text{H}$  and  $\text{C}(\text{sp}^2)\text{–CF}_2\text{H}$  bonds<sup>23–26</sup>. However, these methods are often constrained by the requirement for stoichiometric redox reagents to produce  $\text{CF}_2\text{H}$  radicals via single-electron oxidation or reduction reactions, which might decrease functional group compatibility and lead to the generation of wasteful and often environmentally hazardous byproducts (Fig. 1A).

Difluoroacetic acid is a chemical feedstock that can be considered the most promising source of the  $\text{CF}_2\text{H}$  group due to its availability and stability. However, the demanding oxidation potential of  $\text{CF}_2\text{HCOO}^-$  [ $E_{\text{p}/2} = 2.09$  V versus Ag/AgCl, MeCN] substantially hampers its application in radical difluoromethylation [Fig. 1B, eq. (1)]<sup>27</sup>. Indeed, direct radical decarboxylation of difluoroacetate for the generation of  $\text{CF}_2\text{H}$  radicals typically necessitates harsh reaction conditions, such as strong chemical oxidants or large operating cell potentials, which would limit the reaction efficiency and restrict the substrate scope and functional group tolerance<sup>28–30</sup>. Recently, photoredox iron catalysis has been demonstrated in hydrofluoroalkylation of alkenes with diverse fluoroalkyl carboxylic acids as fluoroalkyl radical precursors through an inner-sphere electron transfer activation mechanism<sup>31,32</sup>, in which a weakly oxidizing iron catalyst could facilitate the radical decarboxylation of redox-reticent alkyl carboxylic acids under mild

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**Fig. 1 | Background information and our strategy for difluoromethylation of (hetero)aryl halides. A** Radical difluoromethylation through the development of new  $\text{CF}_2\text{H}$  radical precursors. **B** Radical decarboxylative difluoromethylation with

difluoroacetate. **C** Electrophotocatalytic transition metal-catalyzed difluoromethylation with difluoroacetate.

reaction conditions [Fig. 1B, eq. (2)]<sup>33–35</sup>. While decarboxylative arylation of aliphatic carboxylic acids has garnered considerable interest in recent years<sup>36–42</sup>, achieving direct decarboxylative arylation of difluoroacetate with aryl halides for the synthesis of difluoromethylarenes is still an elusive goal in the synthetic chemistry community [Fig. 1B, eq. (3)]. Given the pronounced significance of  $C(sp^2)-CF_3H$  motif in medicinal chemistry<sup>43–53</sup>, the direct and efficient use of cost-effective and atom-economic difluoroacetate as the  $CF_3H$  source for the preparation of difluoromethylarenes would be a significant advancement in synthetic chemistry.

Electrophotocatalysis has recently emerged as a powerful synthetic platform, unlocking radical transformations beyond the reach of photochemistry or electrochemistry alone<sup>54-59</sup>. For instance, this strategy has been successfully applied in chemoselective and stereoselective C(sp<sup>3</sup>)-H functionalization in a more sustainable manner<sup>60-69</sup>. In particular, studies on decarboxylative functionalization of alkyl carboxylic acids have shown that the corresponding

alkyl radicals could be generated electrophotochemically under mild and environmentally friendly conditions<sup>70–77</sup>. Furthermore, their subsequent interaction with transition metal catalysis, such as nickel powered by cathode reduction, can be effectively harnessed for the development of radical-based redox-neutral coupling reactions<sup>78–86</sup>.

Here, we show a Fe/Ni dual-transition-metal electrocatalytic protocol for difluoromethylation of (hetero)aryl halides using readily available potassium difluoroacetate as the difluoromethylating reagent. Specifically, electrophotocatalytic Fe-catalyzed decarboxylation of difluoroacetate is applied for the generation of highly reactive  $\text{CF}_2\text{H}$  radical at the anode. Concurrently, electrochemical Ni catalysis, driven by cathode reduction, has been utilized for the activation of aryl halides. This catalytic system allows the construction of  $\text{C}(\text{sp}^2)\text{-CF}_2\text{H}$  bonds from a broad spectrum of aryl and heteroaryl halides along with potassium difluoroacetate with high reaction efficiency (Fig. 1C).

**Table 1 | Reaction discovery and optimization<sup>a</sup>**

<b>Entry</b>	<b>Variation from above conditions</b>	<b>Yield (%)<sup>b</sup></b>
1	none	84 (80) <sup>c</sup>
2	no phthalimide	20
3	TBABF <sub>4</sub> as electrolyte	40
4	no TBABr	42
5	$E_{\text{cell}} = 2.0 \text{ V}$	51
6	FeCl <sub>2</sub> or FeBr <sub>2</sub> instead of Fe(OTf) <sub>2</sub>	54, 44
7	no iron catalyst	0
8	no nickel catalyst	0
9	no electricity	0
10	no light	0

<sup>a</sup>Reaction conditions: **1** (0.2 mmol, 1.0 eq.), CF<sub>2</sub>HCO<sub>2</sub>K (1.0 mmol, 5.0 eq.), DMSO (4.0 mL), undivided cell. <sup>b</sup>Yields were determined by <sup>1</sup>H NMR using 1,1,2,2-tetrachloroethane as the internal standard.

<sup>c</sup>Isolated yield.

## Results and Discussion

### Reaction discovery and optimization

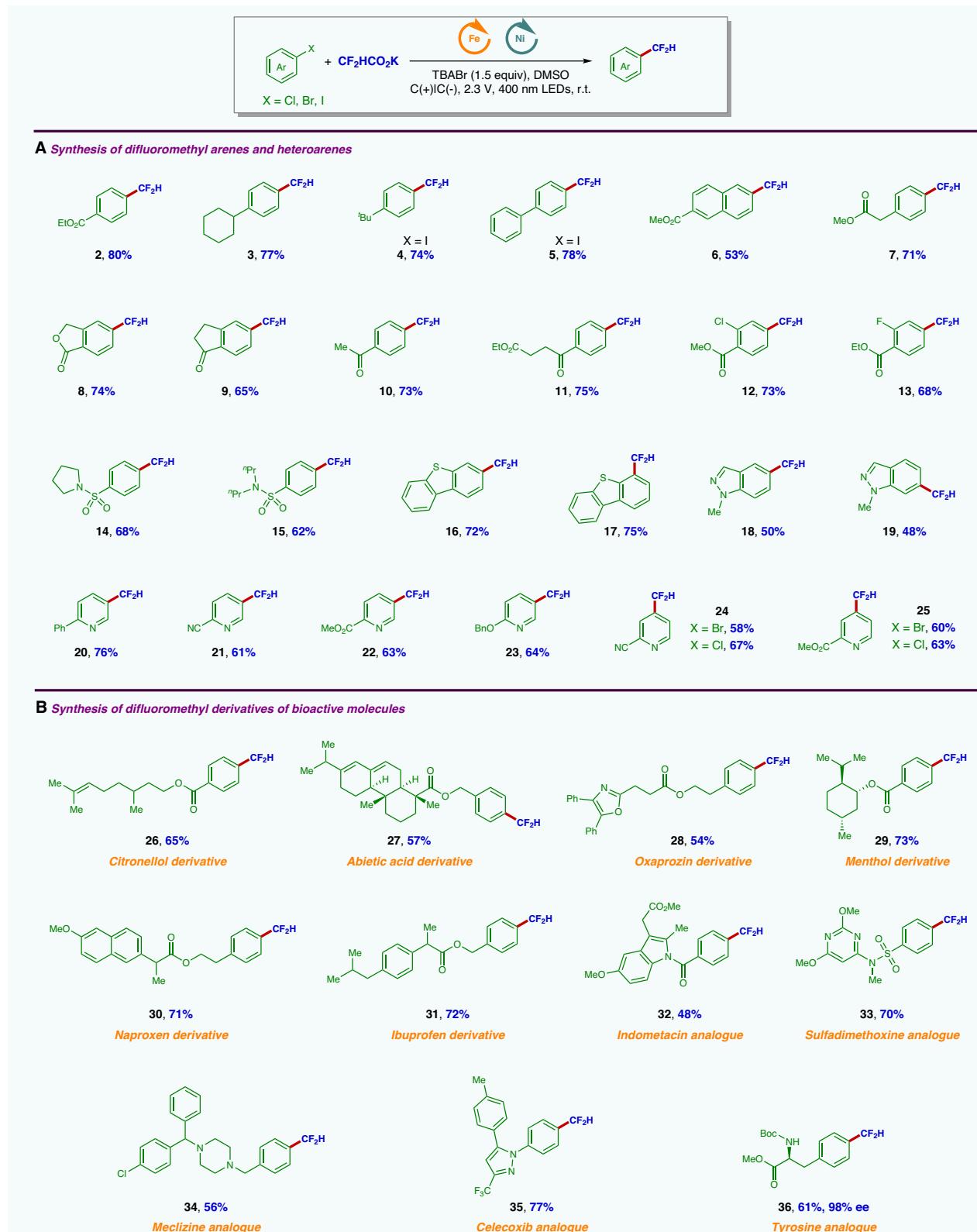
We began our investigation by a strategic combination of iron and nickel catalysis under electrophotocatalytic conditions. Achieving the desired difluoromethylation of aryl halides using our dual-transition-metal electrocatalytic strategy would necessitate the simultaneous activation of two transition metal catalysts at both electrodes, in which each catalyst independently operates on its respective redox event<sup>78,79</sup>. Therefore, the reaction was carried out in a straightforward, undivided cell setup, equipped with two carbon-felt electrodes and maintained at an applied cell potential ( $E_{\text{cell}}$ ) of 2.3 V (corresponding to an initial anodic potential of ca. 0.50 V versus the ferrocenium ion/ferrocene redox couple) while being irradiated with 400 nm light-emitting diodes (LEDs). The use of phthalimide as an additive was believed to stabilize the putative arylnickel(II) species, thus improving the reaction efficiency of radical-based convergent paired electrolysis<sup>39,42,81,82</sup>. With 4,4'-di-*tert*-butyl-2,2'-bipyridine ('B bpy) as ligand to the nickel center, the reaction of ethyl 4-bromobenzoate **1** with potassium difluoroacetate was able to yield the desired decarboxylative coupling product **2** with 80% isolated yield (Table 1, entry 1). As expected, reaction in the absence of phthalimide led to a dramatically diminished yield (entry 2). Although electrolyte could be replaced with TBABF<sub>4</sub> or even be omitted from the reaction system, the addition of TBABr as electrolyte was essential for high reaction efficiency (entries 3 and 4). The applied cell potential of the reaction could be lowered to 2.0 V, yet still giving the coupling product in 51% yield (entry 5). Other iron salts, such as FeCl<sub>2</sub> and FeBr<sub>2</sub>, also proved to be competent catalyst precursors in this catalytic system (entry 6). A series of control experiments was conducted by omitting each component, thereby elucidating the pivotal roles of iron, nickel, electrical energy, and light irradiation in facilitating the desired transformation (entries 7–10).

### Substrate scope

With the optimized reaction conditions in hand, we then evaluated the generality of the aryl halide component in this dual transition metal-catalyzed electrophotocatalytic difluoromethylation protocol. As summarized in Fig. 2, a broad range of aryl halides with different substituents were found to be suitable coupling partners in the

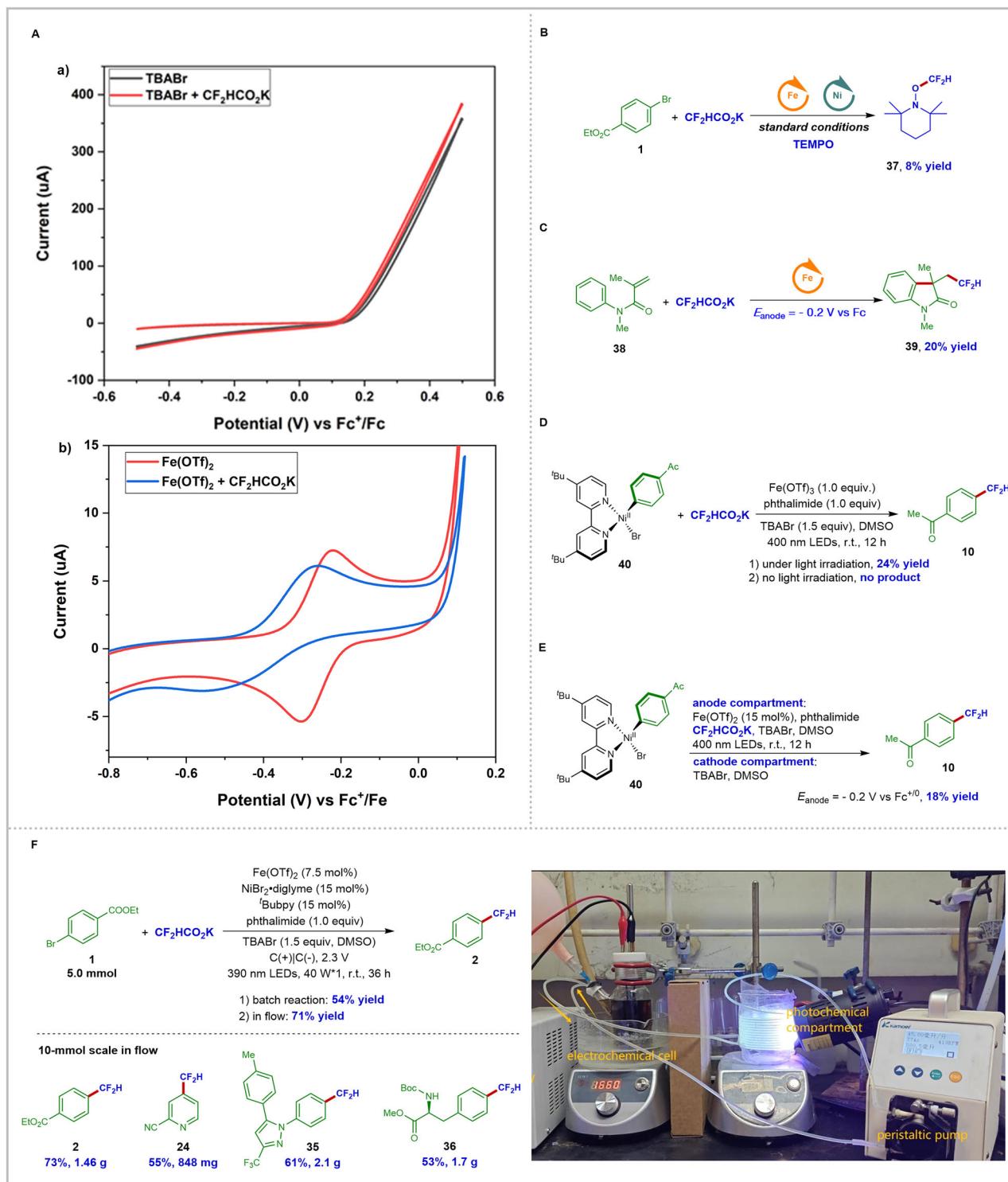
reaction, generating the desired difluoromethylarenes in moderate to good yields (Fig. 2A, 2–25, 48–80% yield). Aryl bromides with both electron-withdrawing and electron-neutral groups performed well to furnish the decarboxylative coupling products with good reaction efficiency (2 and 3, 80 and 77% yield, respectively). In addition, aryl iodides were also found to be competent substrates for the reaction, giving the corresponding difluoromethylarenes with good reaction yields (4 and 5, 74 and 78% yield, respectively). Importantly, the very mild reaction conditions imparted by electrophotocatalytic transition metal catalysis made accessible a broad scope of difluoromethylarenes with a diverse range of functional groups, including esters (2, 6, 7, 8, 11, 12, and 13), ketones (9, 10, and 11), and halides (12 and 13). In particular, the halide functional handles might be subsequently elaborated using orthogonal technologies, e.g., aromatic substitution and cross-coupling reactions to produce more synthetically important molecular structures. In addition, medicinally relevant sulfonamide could be smoothly incorporated into the substrates, with no significant effect on the reaction efficiency (14 and 15). We further expanded the reaction scope to various other heteroarenes including dibenzothiophenes (16 and 17), indazoles (18 and 19), and pyridines (20–25) given their widespread prevalence in pharmaceutical chemistry. All of these aryl bromides were successfully transformed into the corresponding difluoromethylarenes with synthetically useful reaction efficiency. Notably, activated aryl chlorides could also be employed as coupling partners in the reaction to yield the desired products with comparable yields (24 and 25).

The versatility of this method was further highlighted by its capability for late-stage difluoromethylation, providing streamlined access to molecules of pharmaceutical relevance (Fig. 2B). First, difluoromethylated derivatives generated from bioactive molecules including citronellol (26), abietic acid (27), oxaprozin (28), menthol (29), naproxen (30), and ibuprofen (31) could be readily obtained. Notably, functionalities that are susceptible to radical addition, such as alkene (26) and diene (27), can be successfully engaged in this radical decarboxylative arylation. Further, CF<sub>2</sub>H-containing analogs of indometacin (32), sulfadimethoxine (33), meclizine (34), and celecoxib (35) were successfully synthesized from the corresponding aryl bromides. Finally, direct difluoromethylation of phenylalanine-derived aryl bromide led to the formation of product 36 in 61% yield without erosion of

**Fig. 2 | Substrate scope of decarboxylative difluoromethylation of aryl halides.**

**A** Synthesis of difluoromethyl arenes and heteroarenes. **B** Synthesis of difluoromethyl derivatives of bioactive molecules. All yields are of isolated products. Unless otherwise noted, reaction conditions were as follows: Aryl halides

(0.2 mmol, 1.0 eq.),  $\text{CF}_2\text{HCO}_2\text{K}$  (1.0 mmol, 5.0 eq.),  $\text{Fe}(\text{OTf})_2$  (0.03 mmol, 15 mol%),  $\text{NiBr}_2\text{-diglyme}$  (0.03 mmol, 15 mol%), 'Bubpy' (0.03 mmol, 15 mol%), TBABr (0.3 mmol, 1.5 eq.), phthalimide (0.2 mmol, 1.0 eq.), DMSO (4.0 mL), carbon felt anode, carbon felt cathode, under  $\text{N}_2$ , undivided cell, 400 nm 10 W for 12 h.

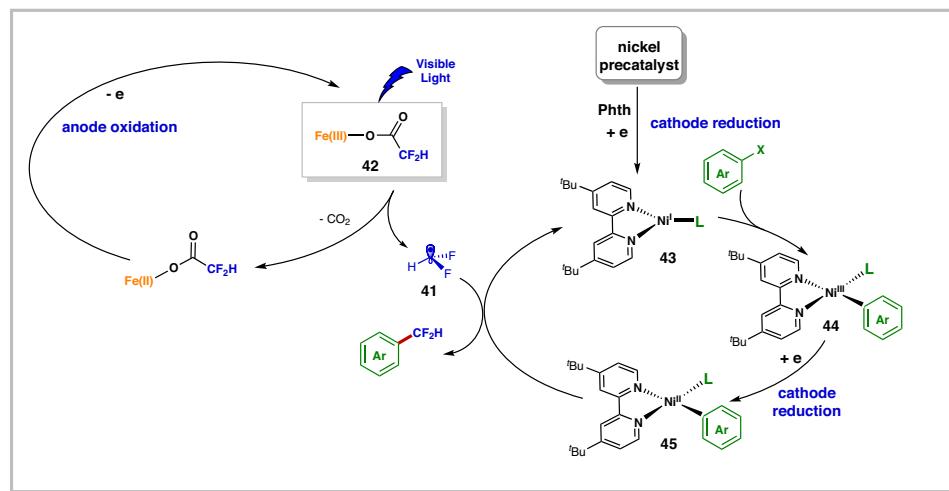


**Fig. 3 | Mechanistic studies and scale-up synthesis in flow. A** Cyclic voltammetry studies. **B** Radical trapping experiment. **C** Radical clock experiment. **D** Stoichiometric experiment without electricity. **E** Stoichiometric experiment using divided cell. **F** Continuous flow electrolysis enables scale-up synthesis.

the stereoselectivity; due to the capacity of the  $\text{CF}_2\text{H}$  group to serve as a lipophilic hydrogen bond donor, it could potentially function as a tyrosine analog in medicinal chemistry. These reaction results, together with those depicted in Fig. 2A, clearly demonstrated the promising application of this dual transition metal-catalyzed electrochemical decarboxylative difluoromethylation method, especially considering the exceptional compatibility with functionally diverse groups that are significant in the context of medicinal chemistry.

### Mechanistic studies and reaction scale-up in flow

Cyclic voltammetry studies of potassium difluoroacetate with  $\text{TBABF}_4$  as the supporting electrolyte indicated that no redox event could be observed between 0 and 0.8 V (Supplementary Fig. S4). However, with  $\text{TBABr}$  as the electrolyte, a distinct anodic event was observed with an onset potential of 0.12 V, due to the oxidation of the bromide anion in the reaction system (Fig. 3A, B). The iron catalyst showed a quasi-reversible redox feature at  $-0.20$  V, which we attributed to the  $\text{Fe}^{III}/\text{Fe}^{II}$



**Fig. 4 | Proposed catalytic cycles.** The possible reaction mechanism for Fe/Ni dual-transition-metal electrocatalytic system for difluoromethylation of aryl halides. Phth phthalimide.

redox couple of the iron catalyst itself (Fig. 3A, B, red line). Further addition of potassium difluoroacetate resulted in a cathodic shift of the redox wave. This shift likely arose from the formation of Fe(II) carboxylate, and the generation of Fe(III) species was facilitated by the coordination of the carboxylate to the Fe center (Fig. 3A, B, blue line). The formation of the CF<sub>2</sub>H radical was supported by the radical trapping experiment; in the presence of the radical scavenger 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO), the desired decarboxylative difluoromethylation of aryl bromide 1 was completely inhibited, with the TEMPO-trapped adduct 37 being formed in 8% yield (Fig. 3B). The electrophotochemical Fe-catalyzed mechanism allows the radical decarboxylation of difluoroacetate to occur in a manner that is much more controlled than those through direct voltage-gated pathways, leading to the observed mild reaction conditions with high chemoselectivity. To support this hypothesis, we designed and performed a controlled potential electrolysis experiment, in which the application of a  $-0.20$  V anodic potential proved sufficient in promoting the difluoromethylation/cyclization reaction of 38, giving the difluoromethylated oxindole 39 in 20% yield (Fig. 3C).

To confirm the formation of arylnickel(II) species in the reaction, we synthesized an arylnickel(II) complex 40 and performed a series of stoichiometric experiments. First, reactions of potassium difluoroacetate with nickel complex 40, in the presence of a stoichiometric amount of Fe(OTf)<sub>3</sub>, revealed that the difluoromethylated product 10 can only be generated upon light irradiation (Fig. 3D). These results clearly indicated the essential role of light irradiation for the reaction. When the reaction was performed in the anode compartment at a controlled potential of  $-0.2$  V, utilizing a setup that isolated the anode from the cathode with a sintered glass frit to prevent bulk mixing of liquids within both compartments, the desired product 10 was obtained with 18% yield (Fig. 3E). This experimental result is in accordance with our mechanistic hypotheses that the arylnickel(II) species formed at cathode is able to transfer into the bulk solution and productively react with the CF<sub>2</sub>H radical generated by electrophotochemical Fe-catalyzed decarboxylation in the anode compartment, thus achieving a challenging radical-based convergent paired electrocatalytic system<sup>51,52</sup>.

Finally, we endeavored to showcase the synthetic utility of this dual transition metal electrocatalytic approach for the difluoromethylation of aryl halides by establishing a gram-scale procedure. However, initial attempts to scale up the reaction to 5 mmol using batch cells resulted in a significantly lower isolated yield compared to that obtained from the small-scale reaction (54 vs 80%). We speculated

that the reduced reaction efficiency might stem from limitations in light absorption, which could impact the photoredox process. Consequently, a continuous flow electrochemical setup featuring a larger photochemical chamber was designed to enhance the kinetics of the photoredox step in the catalytic cycle. After optimizing several reaction parameters, the reaction yield was improved to 71%. Notably, applying this continuous flow protocol enabled the reaction to be readily scaled up to 10 mmol scale without further optimization of the reaction conditions by using 5% mmol loading of the Fe catalyst, yet maintaining comparable reaction efficiency (Fig. 3F). The success of the continuous flow electrophotochemical protocol further confirms that the electrochemically generated transition metal complexes are sufficiently stable to migrate into the bulk solution, thereby enabling subsequent photoredox processes for radical decarboxylative difluoromethylation of aryl halides. More importantly, it is also consistent with our proposed stepwise energy input activation mechanism for the reaction.

Collectively, our experimental observations and mechanistic data support the mechanistic scenario proposed in Fig. 4. The electrophotochemical Fe-catalyzed radical decarboxylation of difluoroacetate at the anode produced the CF<sub>2</sub>H radical 41 in a controllable and homogeneous fashion. The coordination of the carboxylate to the Fe center makes the generation of photoredox-active iron carboxylates 42 more feasible. Concurrently, catalytically active Ni(I) species 43, produced by cathode reduction, would react with aryl halide via oxidative addition to give 44. After gaining one electron from the cathode, the generated arylnickel(II) species 45 would be sufficiently stable to migrate into the bulk solution. The phthalimide additive (Phth) may act as a ligand to stabilize the presumed arylnickel(II) species 45. At this stage, radical addition of the CF<sub>2</sub>H radical 41 to nickel complex 45 delivered the desired difluoromethylation product upon reductive elimination and regenerated Ni(I) species 43 for the next catalytic cycle.

In summary, we have successfully developed a practical electrophotochemical method for the difluoromethylation of (hetero)aryl halides, using readily available potassium difluoroacetate as the CF<sub>2</sub>H radical precursor. This dual transition metal electrocatalytic system exhibits exceptional functional group compatibility and has been effectively demonstrated through the late-stage functionalization of pharmaceutically relevant molecules. Additionally, a continuous flow electrochemical setup has been devised for gram-scale synthesis, showing promise for practical synthetic applications. Given the widespread applicability of the CF<sub>2</sub>H group in medicinal chemistry, we

anticipate that researchers in both laboratory and process would find this work to be of high synthetic value.

## Methods

### General procedure for decarboxylative difluoromethylation of aryl halides

**Pre-catalyst solution.** A pre-catalyst solution was prepared by combining  $\text{NiBr}_2$ ·diglyme (10.6 mg, 15 mol%) and 4,4'-di-*tert*-butyl-2,2'-bipyridine (8.0 mg, 15 mol%) in 4 mL of anhydrous dimethyl sulfoxide (DMSO), stirred for 30 min under a nitrogen atmosphere before use.

**Reaction setup.** An oven-dried two-neck tube was equipped with a stir bar, a rubber septum, a threaded Teflon cap fitted with electrical feedthroughs, two carbon felt anodes ( $0.8 \times 1.7 \text{ cm}^2$ , connected to the electrical feedthrough via a 9.0 cm in length, 2.0 mm in diameter graphite rod). Under nitrogen atmosphere, aryl halide (0.2 mmol, 1.0 eq.),  $\text{CF}_2\text{HCO}_2\text{K}$  (134.1 mg, 1.0 mmol, 5.0 eq.), TBABr (96.7 mg, 0.3 mmol, 1.5 eq.), phthalimide (29.4 mg, 0.2 mmol, 1.0 eq.) and  $\text{Fe}(\text{OTf})_2$  (10.6 mg, 0.03 mmol, 15 mol%) were added, followed by subsequent addition of the pre-catalyst solution. The reaction mixture was then sparged with nitrogen for 5 minutes and maintained under nitrogen atmosphere with a balloon. The reaction was irradiated with LEDs (10 W, 400 nm, 1.0 cm away) under the vessel and electrolysis was initiated at a constant cell potential of 2.3 V. After 12 h at room temperature, the photolysis and electrolysis were terminated, the tube cap was removed, and the electrodes were rinsed with ethyl acetate ( $\text{EtOAc}$ ), which was combined with the crude mixture. The organic layers were further washed with brine, dried over anhydrous  $\text{Na}_2\text{SO}_4$ , and concentrated under reduced pressure. The crude material was purified by flash column chromatography on silica gel to give the desired product.

## Data availability

Materials and methods, optimization studies, experimental procedures, mechanistic studies,  $^1\text{H}$  NMR spectra,  $^{13}\text{C}$  NMR spectra and mass spectrometry data generated in this study are provided in the Supplementary Information. Data supporting the findings of this manuscript are also available from the corresponding author upon request.

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## Author contributions

N. F. conceived and directed the project. N. Z. performed the experiments and developed the reactions. M. C. helped collecting some

experimental data. N. F. wrote the manuscript. N. Z and M. C. prepared the supplementary information and revised the manuscript. All authors have read and approved the final version of the manuscript.

## Competing interests

The authors declare no competing interests.

## Additional information

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