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# PEGylated chitosan xerogels for localized periodontal therapy: design, characterization, and drug release

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It is essential to adopt effective therapy strategies for periodontal diseases to achieve optimal results while avoiding adverse effects on the system. This study has developed various PEGylated chitosan-based biodegradable xerogels for localized release of doxycycline hydrate (DH) to treat periodontal infectious diseases. The xerogels were formulated using the solvent casting method, and the solvent (0.25 M HCl) was slowly evaporated at ambient conditions. Two different molecular weights were employed for chitosan and polyethylene glycol, and twelve combinations, including the placebos and controls, were prepared for the formulation of xerogels. Different physical and chemical characteristics of the prepared DH xerogels were studied, such as drying time and rate, thickness, moisture content, swelling index, organoleptic characteristics, scanning electron microscopy, FTIR spectrometry, differential scanning calorimetry, drug release and kinetics, and antibacterial activity. The results revealed that the drug transforms from a crystalline to an amorphous state, thus rapidly releasing the drug (> 60% in 30 min. in all xerogels), followed by a sustained release up to 10 h. The release kinetics results revealed that the drug followed the Korsmeyer–Peppas model. It is concluded that the formulated DH-loaded xerogels showed promising results for use in the periodontal pockets to treat various infectious diseases.

**Keywords** Doxycycline hydrate, Xerogels, Chitosan, Polyethylene glycol, Hydrochloric acid

Xerogels have garnered significant attention recently as novel drug delivery systems. They are extensively studied for their potential in numerous drug delivery systems<sup>1–16</sup>, including a few reported applications in the oral cavity<sup>17–25</sup>.

The selection of biomedical materials, particularly for applications in periodontal pathologies, is paramount due to their intrinsic characteristics. Not only the type of polymers and drugs but also the nature of the solvent affect the physicochemical characteristics of the dosage form. During the sol–gel process, the solvent facilitates the hydrolysis and condensation reactions necessary to form the gel network<sup>26,27</sup>. It also affects the porosity, surface area, and mechanical strength of the resulting xerogel<sup>28–30</sup>. The solvent volatility and interactions with the gel components are crucial for efficient solvent removal, ultimately determining the xerogel structure and drug release profiles<sup>29,30</sup>. Previously, our group has studied the sustained release effect of doxycycline hydrate (DH) from PEGylated acetic acid-based chitosan xerogels, and the rationale for selecting those polymers has already been discussed<sup>31</sup>. However, the choice of solvent plays a critical role in preparing xerogels, significantly influencing their structural and functional characteristics<sup>32,33</sup>. Despite these promising results, there remains a need to explore alternative solvents that might offer improved control over xerogels' physicochemical characteristics and drug release behavior.

Currently, no xerogel dosage form of DH is available in the market. Internationally, two different novel commercial products, PerioChip® (i.e., a polymeric chip of chlorhexidine gluconate) and Atridox® (a 10% solution of doxycycline and polymers, both given separately in prefilled syringes), are available for the treatment of periodontal diseases. Due to their high costs, they are unaffordable to most patients; thus, dental practitioners rarely use these products. In contrast, the xerogels developed in this study present a cost-effective alternative.

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The present study aims to create a relatively economical xerogel system using hydrochloric acid for local drug delivery of DH. This study would help develop an understanding of the solvent effect on the physicochemical properties and drug release behaviors of PEGylated chitosan xerogels.

## Materials and methods

### Materials

Doxycycline hyolate (DH,  $\geq 98\%$ ), polyethylene glycol (PEG,  $M_w$ : 8000 and 20,000), and chitosan (CHT,  $\geq 95\%$  acylated; medium: 100–200 mPa.s and high: 200–400 mPa.s) were purchased from Shanghai Macklin Biochemical Technology Co., Ltd. (Shanghai, China). All the other chemicals and solvents used were of analytical grade. Freshly prepared double reverse osmosis deionized water with zero total dissolved solids (TDS) was used throughout the study (FineTech Water Treatment, Pakistan).

### Preparation of placebo xerogels

All the solutions of each ingredient were prepared in 0.25 M hydrochloric acid (HCl) in the desired concentrations (Table 1) with continuous magnetic stirring (ACE-HS-5L, ACE GmbH, Germany). A 3% solution of each medium (MMW) and high molecular weight (HMW) CHT was prepared using 0.25 M HCl and stirred for 90 min at  $60 \pm 5$  °C. Various combinations of polymers were prepared with respect to their molecular weights to obtain the final concentration at the desired level (Table 1). Each type of PEG solution (i.e., MMW and HMW) was mixed with each CHT solution (i.e., MMW and HMW) and magnetically stirred for 90 min. at  $60 \pm 5$  °C. Once a homogenous solution was formed and the temperature of the mixture was dropped to  $30 \pm 5$  °C, the polymer mixture was poured (10 ml) into the plastic Petri dishes (55 mm diameter) and left to dry without lids in ambient conditions (i.e.,  $25 \pm 2$  °C) for up to 120 h to develop the placebo xerogels (X1–X6). Of these placebos, xerogels X1, X2, and X7, X8 served as controls containing each CHT type only without or with DH, while X3–X6 were prepared as placebos. Each sample was prepared in at least triplicate. The illustrative image of the formulation procedure is presented in Fig. 1.

### Preparation of drug-loaded xerogels

As described above, polymers (CHT and PEG) were mixed at  $60 \pm 5$  °C. Once the solution temperature dropped to  $30 \pm 5$  °C, the drug solution was added (Table 1), and the mixture was further stirred for 30 min. with occasional overhead stirring (MS 5060D, MTOPS, Misung Scientific, Korea) since it became viscous at room temperature. After forming a homogenous gel matrix, the drug-polymer gels were dried as described above. A DH control with each CHT type only was also prepared similarly. Each sample was prepared in at least triplicate.

### Storage of the xerogels

All xerogel samples prepared in the Petri dishes were covered with the lids and stored in the sealed desiccator at room temperature (i.e.,  $25 \pm 2$  °C).

### pH Measurement of the gels before drying

The pH of the gel samples before drying was recorded with the help of a digital pH meter (Elmetron CP501, Poland) using a glass electrode and a temperature probe. The instrument's calibration was performed using solutions prepared from commercially available buffer tablets of pH 4.00 and 7.00 (Merck, Germany).

### Drying Time and drying rate of the xerogels

The placebo and DH-containing solutions poured into the Petri dishes were weighed on a weighing balance (ALS 220-4, KERN, Germany) at 0, 24, 48, 72, 78, 84, 90, 96, 102, 108, 114, and 120 h. The weighing was continued until the samples reached a constant weight. The following formula determined the drying rate:

$$N = (W^0 - W^D) / t$$

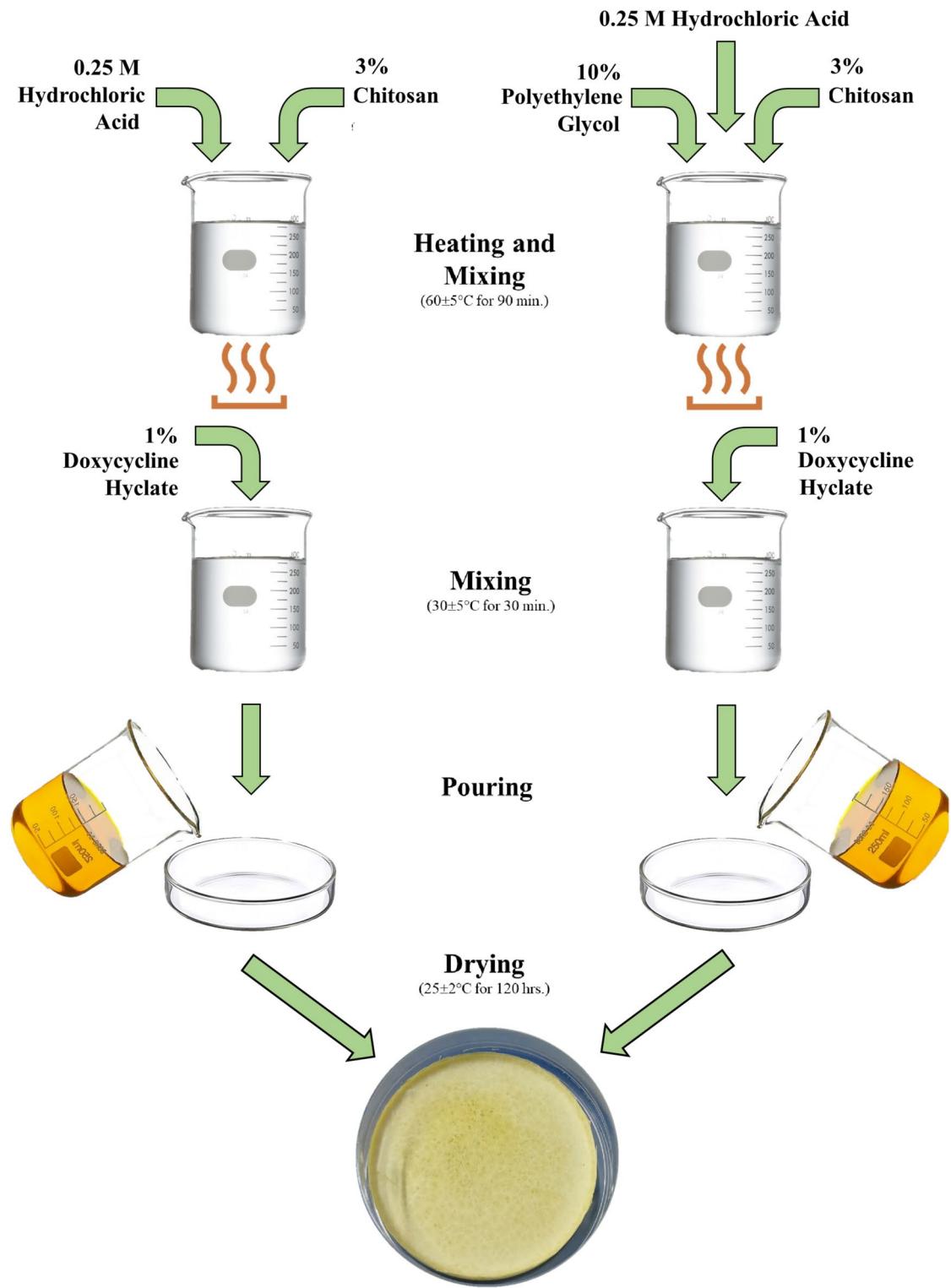
where N = Drying rate,  $W^0$  = Weight of the sample at zero-time (minus the weight of the Petri dish),  $W^D$  = Weight of the sample after drying (minus the weight of the Petri dish), t = Drying time.

### Thickness of the xerogels

The thickness of the xerogels was measured using a calibrated digital vernier caliper (Insize, 1108-150, China). Each sample measurement was made in a triplet.

Ingredients	Composition percentage (%)											
	X1	X2	X3	X4	X5	X6	X7	X8	X9	X10	X11	X12
CHT MMW	3	0	3	3	0	0	3	0	3	3	0	0
CHT HMW	0	3	0	0	3	3	0	3	0	0	3	3
PEG MMW	0	0	10	0	10	0	0	0	10	0	10	0
PEG HMW	0	0	0	10	0	10	0	0	0	10	0	10
DH	0	0	0	0	0	0	1	1	1	1	1	1

**Table 1.** The composition of the xerogels.



**Fig. 1.** Illustrative image showing the steps involved in the preparation of xerogels.

#### Moisture content determination of the xerogels

Moisture content (MC) was calculated by pre-weighing  $10 \times 10$  mm of the xerogel and thoroughly drying it at  $100^{\circ}\text{C}$  in a hot air oven (Memmert, Germany) for 60 min. The MC of the dried xerogel samples was determined after reweighing the samples using the following formula<sup>34–36</sup>:

$$\text{MC } (\%) = [(W_{\text{gel}} - W_{\text{dry}})/W_{\text{gel}}] \times 100$$

where  $W_{\text{gel}}$  is the pre-drying weight of the xerogel,  $W_{\text{dry}}$  is the post-drying weight of the xerogel.

### Swelling index of the xerogels

The basic steps in calculating the swelling index (SI) ratio are the same as those discussed in the section on MC. The dried samples were weighed and placed on a porous aluminum foil (porous to act as a sieve). The foil containing the dried sample was placed in a Petri dish filled with simulated saliva (pH 6.8) and left for 15 min. to swell. The additional saliva was removed, and the sample was reweighed. The following formula was used to determine the SI ratio<sup>37–41</sup>:

$$\text{SI} = [(W_{\text{swell}} - W_{\text{dry}}) / W_{\text{dry}}]$$

where  $W_{\text{swell}}$  is the weight of the swelled xerogel sample,  $W_{\text{dry}}$  is the weight of the dried xerogel sample.

### Organoleptic studies

The organoleptic parameters of the xerogels were analyzed using a descriptive sensory analysis. A total of ten panelists with normal sensory acuity were included, trained, and oriented. They received a questionnaire (Table S1) with instructions on recording the perceived results<sup>42</sup>. The parameters of the questionnaire included fragrance, taste (taste and aftertaste), and texture [visual (color, gloss, smoothness, appearance) and tactile characteristics (roughness, hardness)]. The visual analysis was performed with the naked eye. The tactile elements were determined by retrieving a small piece and feeling it by holding it between the index finger and thumb, whereas other properties were evaluated visually. The study followed the Helsinki guidelines after approval from the Institutional Research and Ethics Board of Baqai Medical University 'Ref: BMU-EC/03-2022', and informed consent was obtained from all the participants.

### Scanning electron microscopy (SEM) analysis of the xerogels

The surface topography of placebo and DH-loaded xerogels was performed using a scanning electron microscope (JSM-6380A, JEOL, Japan) at an accelerating voltage of 10 kV. Before the analysis, the samples were fixed on a 0.5-inch aluminum stub using carbon adhesives. Sputter coating (JFC-1500, JEOL, Japan) of the samples was performed using gold to make the samples electrically conducive. Various magnifications were chosen to obtain the best details of the samples.

### Fourier transform infrared (FTIR) spectrometric analysis of the xerogels

FTIR spectrometer (Nicolet iS5, Thermo Fisher Scientific, USA) was used to record the IR spectra of all the raw materials, their physical mixtures, and all xerogel samples. The physical mixtures were prepared in the same concentration as the formulations by mixing the drug powder with the polymer in a Petri dish with the help of a spatula for five minutes. The samples were placed on the diamond crystal optical base (iD7 ATR, ThermoFisher Scientific, Great Britain). Each spectra was recorded at 4 cm<sup>-1</sup> with 128 scans between 4000 and 700 cm<sup>-1</sup>. The spectra were analyzed using Omnic software (version 9).

### Differential scanning calorimetry (DSC) analysis of the xerogels

The thermal characteristics of the pure, physical mixtures and xerogel samples were assessed using a differential scanning calorimeter (DSC 100, Lab-Kits, Hong Kong). DSC was purged with 20 ml/min nitrogen, and the data was processed using built-in Pyris software. Standard indium and zinc were used to calibrate the instrument. The DH-loaded xerogel samples were weighed in ceramic crucibles and heated to 400 °C at a rate of 10 °C per min. The samples were cooled at 20°C per min. and the data was recorded until they reached room temperature.

### In-vitro drug release kinetics of the xerogels

Drug release from the xerogels was performed using a dissolution apparatus basket method (Type I, GDT-6L, Galvano Scientific, Pakistan)<sup>43</sup>. Each xerogel sample (55 mm diameter) was placed into the basket that was rotated at a speed of 200 rpm into the dissolution vessels containing 600 ml of simulated saliva kept at a constant temperature of 37 ± 0.5 °C. The samples (5 ml each) were periodically withdrawn at 0, 2, 5, 10, 15, 30, 60, 120, 240, 360, and 600 min. The volume of the withdrawn sample was replaced with a fresh medium to maintain the sink conditions. The samples were subjected to UV spectrometric assay (Evolution 201, ThermoFisher, USA) to quantify the amount of DH released from the xerogels.

The kinetics of the drug release from the xerogels was evaluated using the following equations<sup>44–46</sup>:

$$\text{Zero - order : } C_0 - C_t = k_0 t$$

$$\text{First - order : } \ln (C_0 / C_t) = kt$$

$$\text{Higuchi : } C_t = k_H t^{1/2}$$

$$\text{Hixson - Crowell : } C_0^{1/3} - C_t^{1/3} = kt$$

$$\text{Korsmeyer - Peppas : } C_t / C_0 = k t^n$$

where  $C_0$  = initial concentration;  $C_t$  = concentration at time  $t$ ;  $k_H$  = Higuchi dissolution constant;  $k$  = release rate constant;  $n$  = slope.

### Antibacterial analysis

The bacterial strains used in the study were *Escherichia coli* (ATCC 25,922), *Staphylococcus aureus* (ATCC 6538), *Pseudomonas aeruginosa* (ATCC 27,853), and *Streptococcus mutans* (ATCC 35,668). The antibacterial activity of the samples was evaluated using the disc diffusion method. Tryptone soya agar (TSA) was used for *E. coli*, *S. aureus*, and *P. aeruginosa*, and blood agar (i.e., TSA + human blood) was used for *S. mutans*<sup>47</sup>. A uniform amount (i.e., 19 ml) of media was poured into each petri dish to maintain the thickness of 4 mm in each plate. Each culture was standardized using the 0.5 McFarland standard. A uniform lawn of each bacterial culture was made on each plate using a sterile cotton swab. DH-loaded xerogels and placebo xerogel discs (5 × 5 mm) were inoculated at equal distances in the plates. The plates were incubated in an aerobic environment at 37 °C for 24 h, and the diameter of the growth inhibition zones surrounding each xerogel disc was measured in mm. DH solution (1%) was used as a control, and the discs (5 mm diameter) were prepared by pouring 5 µl of the solution onto it using a sterile micropipette. Each experiment was performed in triplicate for each sample.

## Results and discussion

### Development of the xerogels

The formulation ingredients possess good in-silico compatibility and are suitable for developing xerogels for periodontal applications<sup>31</sup>. The drug and polymer solutions prepared in 0.25 M HCl appeared homogenous. A total of 12 xerogels were prepared (Table 1) and subjected to various physical and chemical characterizations.

### pH Measurement of the gels before drying

After the ingredients of the xerogels were carefully mixed, the pH of each gel was measured before it was placed in a Petri dish to begin the drying process. The placebo and DH-loaded gels showed nearly identical pH values (Table 2). No pH adjustment was made to the gels before drying. The paired sample two-tailed *t*-test comparing the pH values between the placebo and drug-loaded gels did not reveal any significant difference ( $p=0.363$ ). This finding implies that adding DH to the xerogel matrix did not result in a discernible shift in the medium acidity or alkalinity compared to the placebos.

### Drying time and drying rate of the xerogels

The drying time, defined as the time a sample needs to lose its moisture at ambient temperature (i.e., 25 ± 2 °C), is influenced mainly by the polymer composition<sup>48</sup>. The drying time and rate are significantly influenced by temperature, humidity, and surface area<sup>49–51</sup>. It is crucial to remember that the raw materials' moisture content may affect the drying time<sup>52</sup>. In this study, the samples were periodically weighed for a maximum of 120 h to determine the drying time, with the recorded drying time signifying the time taken for the samples to reach complete dryness (Table 2).

The drying time of the placebo xerogels was observed between 90 and 102 h, while that of the drug-loaded xerogels was achieved between 66 and 86 h (Table 2). The high drying time is due to the viscous nature of the gels, which requires adequate time to transform into xerogels in ambient conditions. On the other hand, the time difference can be attributed to the supplementary mixing of the drug-loaded gels during their preparation, which results in some moisture loss. However, the drying rates showed no significant difference between the placebo and drug-containing xerogels of a similar type (i.e., X1 with X7, X2 with X8, etc.), and this slight difference can be attributed to factors such as those mentioned during the drying of the gels (Table 2).

It has also been observed that the drying rate is comparatively lower in most of the xerogels containing higher molecular weight polymers. These observations can be explained by the molecular weight of CHT, which influences the moisture release characteristics. Increasing the molecular weight of CHT leads to a decrease in hydration and, consequently, a reduction in drying rate<sup>31,53,54</sup>. Similarly, PEG is known for its hygroscopic nature, and its molecular weight is inversely related to its hygroscopicity. Thus, increasing the molecular weight of PEG

Xerogels	pH	Rate ± SD (mg/h)	Time ± SD (h)	Mean thickness ± SD (mm)	MC ± SD (%)	SI ± SD
X1	0.9	87.64 ± 2.28	92 ± 7.48	0.24 ± 0.02	13.13 ± 0.40	4.91 ± 1.91
X2	0.9	81.37 ± 0.00	102 ± 4.90	0.23 ± 0.02	10.31 ± 3.76	5.04 ± 0.43
X3	1.5	83.42 ± 2.84	98 ± 10.20	0.55 ± 0.06	3.94 ± 1.33	3.77 ± 0.59
X4	1.3	84.19 ± 5.39	96 ± 9.80	0.55 ± 0.07	10.76 ± 0.59	3.90 ± 1.54
X5	1.3	80.38 ± 1.32	90 ± 9.80	0.54 ± 0.10	5.21 ± 0.90	3.63 ± 0.45
X6	1.1	79.00 ± 2.47	92 ± 10.20	0.53 ± 0.08	11.85 ± 0.62	3.97 ± 0.45
X7	1.0	93.29 ± 4.74	66 ± 12.96	0.27 ± 0.02	11.97 ± 0.85	4.78 ± 2.70
X8	1.0	87.44 ± 0.65	70 ± 17.20	0.24 ± 0.02	8.65 ± 1.60	4.86 ± 2.63
X9	1.4	83.30 ± 3.17	86 ± 10.20	0.58 ± 0.03	2.97 ± 1.47	3.60 ± 0.73
X10	1.3	93.86 ± 1.81	80 ± 7.48	0.55 ± 0.13	7.82 ± 1.60	3.48 ± 0.22
X11	1.4	81.22 ± 2.59	82 ± 7.48	0.55 ± 0.04	3.97 ± 0.44	3.39 ± 0.31
X12	1.1	94.08 ± 6.23	70 ± 15.75	0.54 ± 0.07	9.52 ± 1.51	3.37 ± 1.02

**Table 2.** Physical characterization data of the placebo (X1–X6) and drug-loaded (X7–X12) xerogels. SD standard deviation, MC moisture content, SI swelling index.

reduces the hygroscopic potential, which reduces the drying rate<sup>55,56</sup>. It is, therefore, essential to understand variations in drying time to maintain consistent xerogel quality and improve drug delivery performance.

### Thickness of the xerogels

Xerogels containing both types of polymers are thicker than those containing CHT only (Table 2). The higher molecular weight of the polymers seems to have little effect on the thickness of the xerogels (Table 2). These findings suggest that multi-polymer xerogel formulations may have advantages in periodontal applications. This could facilitate their application due to their thicker consistency without causing local irritation while maintaining steady drug delivery performance.

### Moisture content of the xerogels

It is crucial to perform MC evaluation to predict the synthesized samples' biocompatibility<sup>57</sup>. Generally, a higher MC in a sample indicates the presence of hydrophilic agents inside the xerogel matrix<sup>58</sup>. Additionally, high MC can predict increased microbial growth and decreased sample purity<sup>59</sup>. This study observed approximately 3–13% MC in the xerogels (Table 2). It is known that CHT possesses hygroscopic characteristics, mainly when dry<sup>60</sup>, while for PEG, it is reported that their hygroscopicity decreases with increasing molecular weight and that, in general, all PEGs with higher molecular weight to 4000 are non-hygroscopic<sup>60</sup>.

After heating the xerogels in a hot air oven, the CHT-containing xerogels showed no visible changes in their appearance. However, samples containing PEG showed a slight alteration in color, as PEGs are known to discolor at high temperatures<sup>60</sup>. Additionally, all samples were more fragile after being removed from the hot air oven. Furthermore, xerogel samples composed of MMW CHT and PEG displayed lower MC than xerogels with HMW (Table 2). This phenomenon can be explained by the longer molecular chains in HMW CHT and PEG, which results in better water absorption within the molecular structure due to hydrogen ligation with the terminal –OH groups<sup>56</sup>.

Surprisingly, the xerogels containing MMW PEG had a very low MC. This phenomenon can be explained by the possibility that the glass transition temperature of MMW PEG is lower than that of HMW PEG (Table 2). Therefore, MMW PEG can present a more soft and flexible polymeric structure, facilitating water absorption and forming compact and solid xerogels. This, in turn, decreases water dispersion and reduces moisture loss during heating. On the other hand, xerogels containing DH presented a lower MC than the placebo samples, probably due to the greater uniformity of xerogels obtained due to excessive mixing, as discussed earlier in the Drying Rate and Drying Time Section (Table 2). Understanding MC is essential to ensure adequate drug delivery, highlighting its importance in xerogel-based therapies. It is, therefore, necessary to evaluate the MC to ensure the purity and biocompatibility of xerogels, especially in local drug delivery applications in periodontal pockets. The observed MC, influenced by CHT and PEG, affects the physical characteristics and yield of xerogels.

### Swelling index of the xerogels

PEG, CHT, and DH in the xerogels can absorb significant amounts of water<sup>61–63</sup>. The SI plays a vital role in evaluating the ability of dried xerogel samples to absorb moisture or water during storage<sup>64</sup> or when exposed to the oral cavity. This parameter also provides information about the release characteristics of the drug from the xerogel<sup>65</sup>. The hydrophilic functional groups in the xerogel matrix, i.e., –OH, –NH<sub>2</sub>, and –COOH, play an essential role in the swelling behavior, allowing the xerogel to absorb significant fluid volumes<sup>66</sup>.

In this study, xerogel samples were dried according to the instructions provided in the section on MC, which resulted in physical changes similar to those mentioned previously. After immersion in artificial saliva, there was an increase in the weight and volume of the xerogel samples (Table 2). Xerogels containing CHT-only showed a more significant weight change with comparatively higher SI levels (X1 and X2, X7 and X8) than the samples containing both polymers (Table 2). It may have occurred due to comparatively increased water absorption in CHT since the  $pK_a$  of CHT is  $\sim 6.5$ <sup>67</sup>. The results are in accordance with those observed in the MC analysis, and the reasons could be the same as those already explained in that section.

On the other hand, adding PEG caused a decrease in the SI ratio of xerogels (Table 2), probably due to its presence in greater quantities (Table 1), making the polymeric matrix more dense and less sensitive to water. The SI relationship between MMW CHT and PEG was lower than that between HMW CHT and PEG (Table 2). This might have occurred due to the comparatively stronger binding of water with HMW polymers through hydrogen bonding<sup>56</sup>, as also observed in the MC analysis (Table 2). Xerogels containing DH showed comparatively lower SI than placebos, which may be attributed to the hydrogen bonding of DH with the polymers, which allows for relatively lower interaction between saliva and polymers (Table 2). A two-tailed *t*-test revealed a significant ( $p=0.012$ ) difference between placebo and drug-loaded xerogels. The study highlights the impact of polymer concentration variations, particularly CHT and PEGs, on the SI. The prepared xerogels thus show good swelling capability, enabling them to withstand the saliva and also avoid removal of the drug by the saliva from the periodontal pockets. It is essential to capture this dynamic to optimize xerogel formulations to ensure efficient drug delivery and achieve positive therapeutic outcomes during periodontal therapies.

### Organoleptic studies

All xerogel samples were subjected to a thorough organoleptic analysis to assess the appearance and tactile properties, including the texture of the xerogels. Additionally, their sensory characteristics, such as odor, taste, and aftertaste, were evaluated.

#### Visual evaluation

The visual assessment of the xerogel samples revealed distinct characteristics (Table 3). MMW and HMW CHT-only xerogels (X1 and X2) exhibited mixed color shades of '5' and '6', featuring a matte, smooth, and unblemished

Xerogel	Visual		Tactile	Olfactory	Gustatory	
	Color	Appearance			Taste	Aftertaste
X1	6	MT, SM, UF	SM, HD	NS	B	B
X2	6	MT, SM, UF	SM, HD	NS	B	B
X3	4	MT, GR, UF	SM, HD	CH	B/S	B
X4	5	MT, SM, UF	SM, SO	CH	S	B
X5	4	MT, GR, UF	SM, HD	CH	B/S	B
X6	5	MT, SM, UF	SM, SO	CH	S	B
X7	6	MT, SM, UF	SM, HD	NS	B	B
X8	6	MT, SM, UF	SM, HD	NS	B	B
X9	5	MT, GR, UF	SM, HD	CH	B/S	B
X10	9	MT, SM, UF	SM, SO	CH	S	B
X11	5	MT, GR, UF	SM, HD	CH	B/S	B
X12	9	MT, SM, UF	SM, SO	CH	S	B

**Table 3.** Organoleptic evaluation of the xerogels. *MT* Matt, *SM* smooth, *GR* grainy, *UF* un-fissured, *HD* hard, *SO* soft, *NS* no smell, *CH* chemical, *S* sour, *B* bitter.

surface. The xerogels with MMW CHT and MMW PEG (X3), as well as those with HMW CHT and MMW PEG (X5), were rated '4' and displayed a matte, grainy texture but remained unblemished. The xerogels with MMW CHT and HMW PEG (X4), as well as those with HMW CHT and HMW PEG (X6), earned a color rating of '5' and exhibited a matte, smooth, and unblemished appearance. On the other hand, DH-loaded xerogels exhibited the same appearance but with significant color change. The color of DH-loaded MMW CHT xerogels and HMW CHT xerogels (X7 and X8) is rated '6'. The DH-loaded xerogel with MMW CHT and MMW PEG (X9) and those with HMW CHT and MMW PEG (X10) were rated '5' and '6', respectively. The DH-loaded xerogels with MMW CHT and HMW PEG (X11) and HMW CHT and HMW PEG (X12) earned a color rating of '9'. The visual color comparison of the xerogels is shown in Fig. S1.

#### Tactile evaluation

The tactile examination of the xerogel samples revealed variations in surface texture (Table 3). All the xerogel samples (X1, X2, X3, and X5) showcased smooth and hard surfaces, except xerogels containing MMW CHT with HMW PEG (X4) and HMW CHT with HMW PEG (X6), which were characterized by a smooth and soft texture. The DH-loaded xerogels exhibited the same textures as those of the placebo xerogels. The variations observed could be due to the different molecular weights of PEG since HMW PEGs impart more remarkable plasticity or smoothness<sup>60</sup>, as observed in this study.

#### Olfactory evaluation

The olfactory assessment revealed a slight chemical (acid-like) smell in several xerogel samples (X3, X4, X5, X6) (Table 3). In contrast, no distinct smell was detected in CHT xerogels (X1 and X2). The DH-loaded xerogels exhibited the same smells as the placebo xerogels.

#### Gustatory evaluation

The gustatory evaluation of the xerogels revealed various taste characteristics (Table 3). CHT xerogels had a bitter taste (X1 and X2). Xerogels of MMW CHT with MMW PEG (X3) and HMW CHT with MMW PEG (X5) had a complex taste profile, including bitter and sour notes. Xerogels of MMW CHT with HMW PEG (X4) and HMW CHT with HMW PEG (X6) had a sour taste. The DH-loaded xerogels exhibited the same tastes as the placebo xerogels. All the xerogels are found to have a bitter aftertaste.

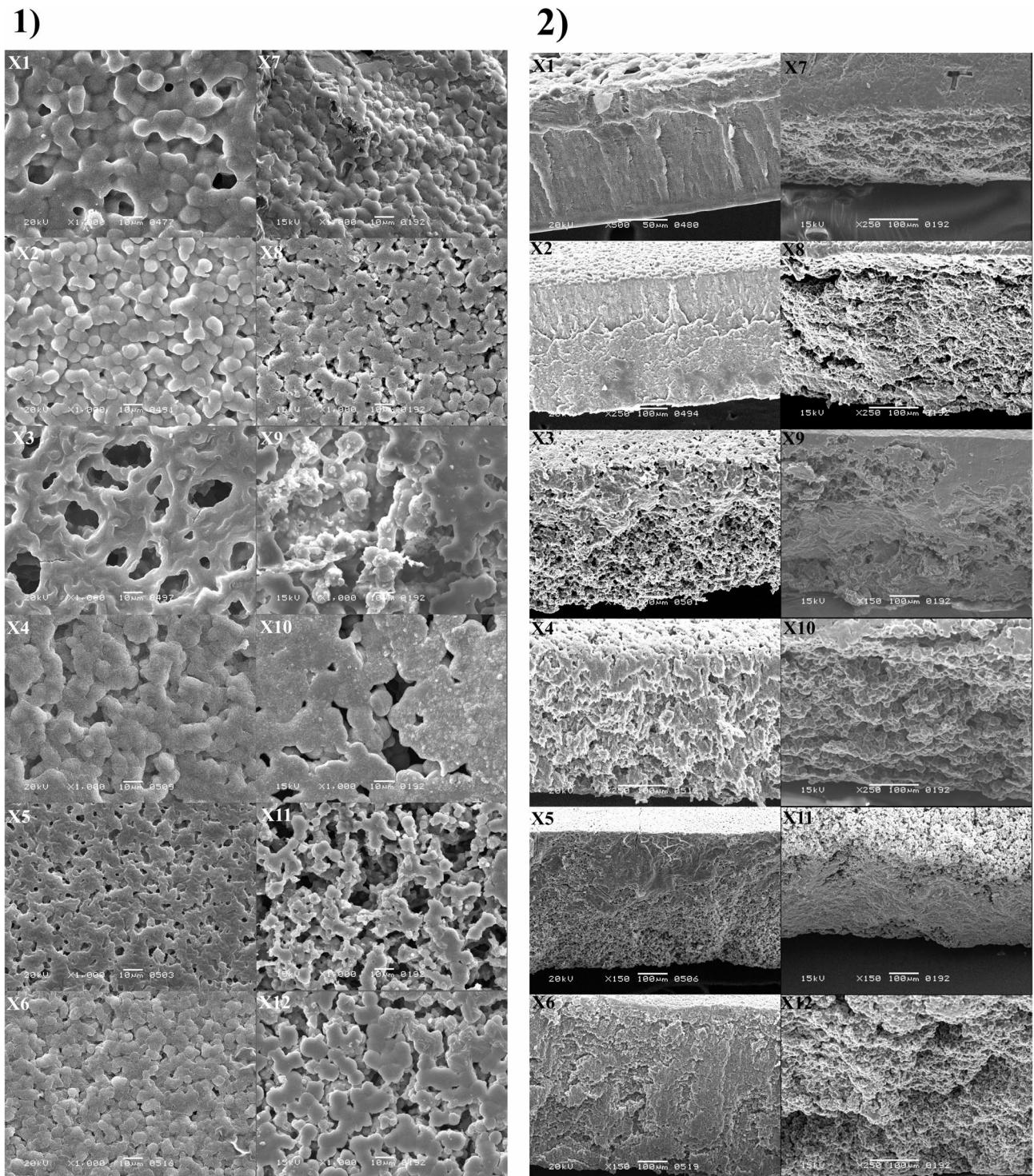
#### SEM analysis of the xerogels

The morphology and surface topography of placebo and drug-loaded xerogels have been examined using SEM, and their differences have been highlighted. According to SEM analysis, all xerogels exhibited a porous texture, consistent with previous research results<sup>31,68</sup>. The pore formation decreases with increasing molecular weights of both polymers (Fig. 2.1). The presence of the drug on the surface of the xerogels can also be visualized, particularly in the image of samples X7, X9, and X10 (Fig. 2.1).

Analysis of the cross-sectional area of the xerogel samples using SEM revealed rough and irregular surfaces, which became denser as the molecular weight of the polymers increased (Fig. 2.2). The origin of these differences lies in the different chemical reactions that occur during the formation phase of the xerogel. The use of HCl induces acetylation of CHT, which increases the molecular weight of CHT<sup>69,70</sup>, giving it a more dense appearance. The intermolecular interactions and hydrogen bonding between polymeric chains create voids, thus creating a porous structure<sup>71,72</sup>.

#### FTIR spectrometric analysis of the xerogels

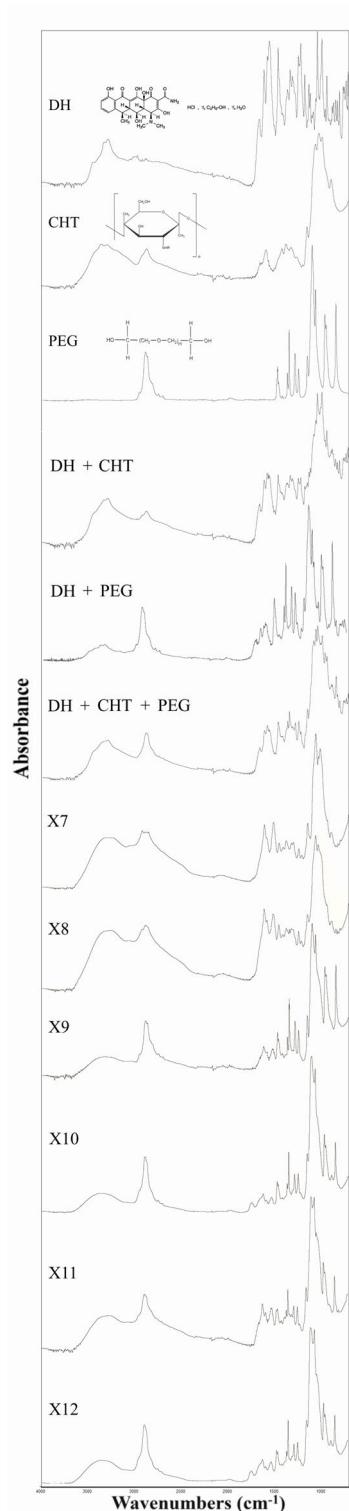
It is of utmost importance to ensure the purity of the active drug and avoid any interaction with other materials when developing a dosage form for any delivery system. It is essential to identify any possible interferences from



**Fig. 2.** The longitudinal (1) and cross-sectional (2) SEM images of placebo and DH-loaded xerogels. X1–X6 are the images of the placebo xerogels, while X7–X12 are the images of the drug-loaded xerogels.

degrading products or unwanted impurities in the sample<sup>73</sup>. FTIR spectrometry was used to ascertain the purity of DH and all other excipients used to develop xerogels. Individual analyses of the raw materials showed that the significant peaks of all compounds (Fig. 3) are in accordance with their chemical structure and correspond to those reported in the literature that have been discussed in detail elsewhere<sup>31</sup>.

The FTIR analysis on the physical mixtures of DH with each excipient (Fig. 3) was also evaluated to identify any molecular interactions between them in the solid state. The physical combination of DH with CHT did not show any prominent changes in the spectrum. All major peaks were observed without significant alterations (Fig. 3). On the other hand, there was a slight broadening between 3500–3100  $\text{cm}^{-1}$  and 3000–2700  $\text{cm}^{-1}$ , which



**Fig. 3.** FTIR analysis of the pure (DH, CHT, and PEG), physical mixtures (DH + CHT, DH + PEG, and DH + CHT + PEG), and drug-loaded xerogel samples (X7–X12).

corresponds to the stretching vibrations of O–H and N–H groups. Some minor variations in the peaks between 1700–1500  $\text{cm}^{-1}$  that is attributed to the C=C, =O, and C–N stretching and N–H bending vibrations and 1100–900  $\text{cm}^{-1}$ , which are characterized by C–H asymmetric stretching, C–O–H bending, C–N stretching, C–O–C, C–H, and C–C ring vibrations have been noted when DH and PEG were mixed physically (Fig. 3). These changes indicate towards the formation of weak hydrogen bonds between the drug and PEG when mixed physically.

On the contrary, significant changes were observed when the xerogels were subjected to FTIR analysis, i.e., broadening, shifting, and disappearance of specific peaks, identified and compared in Table 4. These modifications confirm the formation of hydrogen bonds between the drug and polymers, as well as possible changes in the solid state characteristics of the drug, i.e., the transition from crystalline state to amorphous form during xerogel preparation (Fig. 3). It is evident from the spectra that both the drug and the polymers, particularly PEG, possess specific ionizable groups, i.e., OH and –NH (Fig. 3), which could result in the formation of hydrogen bonds between them (Fig. 3). The transformation in the solid state properties of DH could be confirmed by studying the thermal characteristics using DSC.

#### DSC analysis of the xerogels

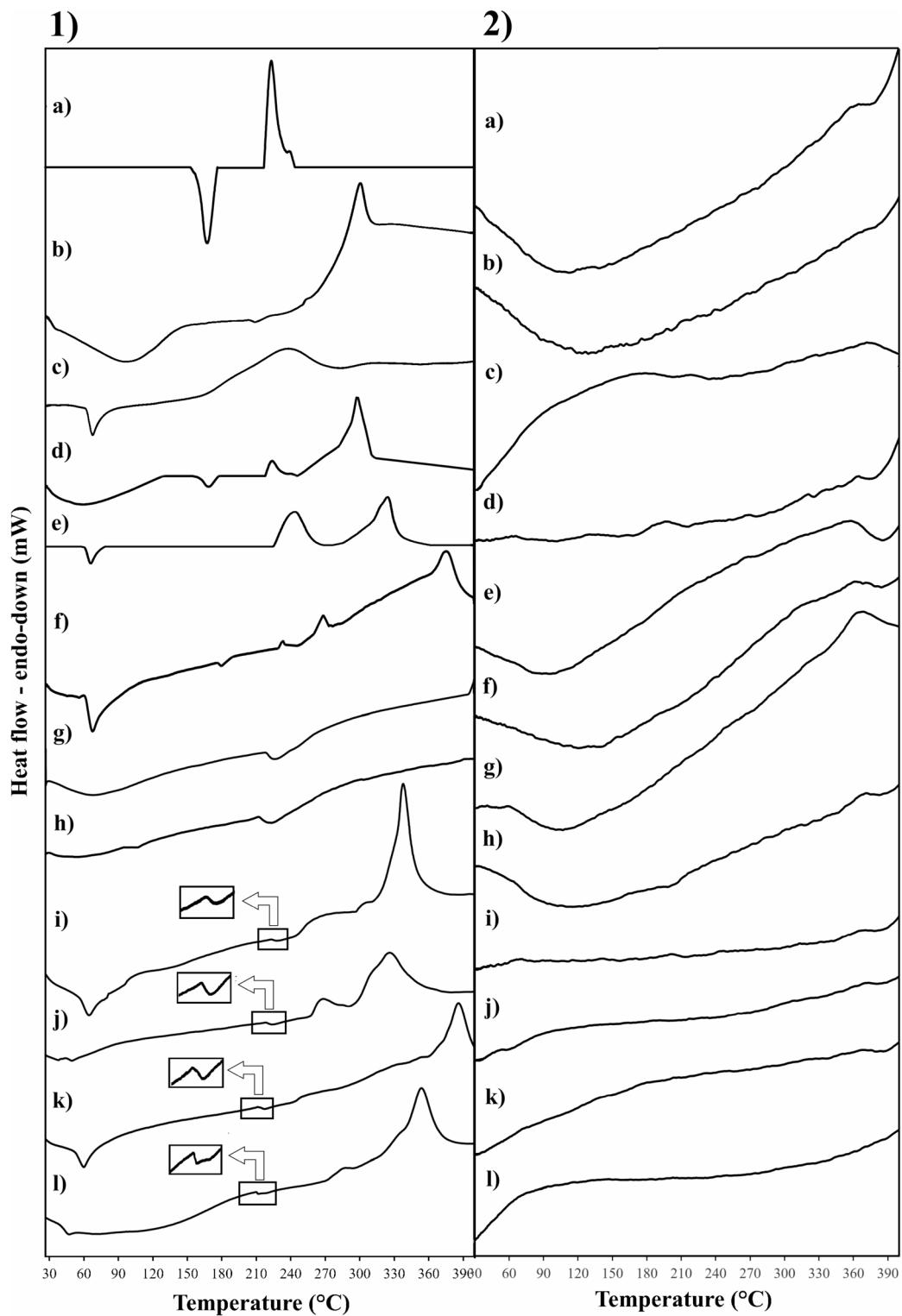
DSC was used to evaluate the thermal characteristics of the raw materials, physical mixtures, and xerogel samples. DH displayed two melting peaks at  $170.87 \pm 3.20$  °C and  $225.98 \pm 2.71$  °C (Fig. 4.1a), corresponding to the melting of crystalline and semicrystalline particles of DH<sup>31,74–78</sup>. The CHT revealed a significant endothermic hump at approximately 50–150 °C (Fig. 4.1b), indicating water loss. In addition, CHT showed a melting peak at approximately  $291.83 \pm 4.89$  °C (Fig. 4.1b), corresponding to the decomposition of their amine groups<sup>79–81</sup>. Interestingly, HMW CHT showed a steeper gradient than MMW CHT, suggesting more significant water loss and a more intense exothermic process<sup>82</sup>. PEG displayed a gradient of  $64.60 \pm 1.70$  °C, suggesting its degradation, while it showed a significant exothermic hump at approximately 200–260 °C (Fig. 4.1c)<sup>83,84</sup>.

Regarding the physical mixtures, there was no significant change in thermal characteristics between DH and CHT (Fig. 4.1d), and there was only a slight decrease in the level of DH peak. On the other hand, shifting of the DH melting peak was observed in the DH and PEG physical mixture (Fig. 4.1e). This change in the thermal peak of DH testifies to an interaction between the two compounds. This interaction was also observed during the FTIR study of the two compounds (Fig. 3). The physical combination of all three compounds also showed a similar thermal graph with reduced DH peaks with a slight shift (Fig. 4.1f). The thermal variations of the physical mixtures of DH and PEG agree with the FTIR data, which further confirms the interaction and formation of hydrogen bonds between the two compounds.

The drug-loaded HCl-based xerogels showed similar changes with more profound results than the physical mixtures (Fig. 4.1g–l). No peak of DH in any of the xerogels has been observed (Fig. 4.1g–l). Thermal data confirmed that the solid-state characteristics of the drug have significantly changed, thus supporting the results of the FTIR analysis. The distinction between the xerogels thermograms of CHT-only (X7 and X8) and those

Wavenumbers (cm <sup>-1</sup> )						
DH	X7	X8	X9	X10	X11	X12
3278	3255	3255	3250	3350	3256	3344
–	2921	2880	2877	2881	2875	2882
1663	–	–	–	–	–	–
1609	1613	1613	1616	1616	1612	1616
1552	1514	1515	1521	1527	1521	1525
1458	1456	1455	1465	1465	1454	1465
1358	1378	1374	1358	–	1358	–
1330	–	–	1340	1341	1340	1341
1307	1302	–	1278	1279	1279	1279
1242	1244	1242	1240	1241	1240	1241
1217	–	–	–	–	–	–
1174	1149	1148	1146	1143	1146	1143
1129	–	–	–	–	–	–
1115	–	–	1095	1097	1093	1104
1084	1062	1060	1058	1060	1060	1060
1038	1016	1034	–	–	–	–
955	–	–	958	961	960	962
937	–	–	–	–	–	–
880	–	–	–	–	–	–
867	–	–	–	–	–	–
843	–	–	840	841	841	841
824	–	–	–	–	–	–
802	–	–	–	–	–	–
761	–	–	–	–	–	–
751	–	–	–	–	–	–
730	–	–	–	–	–	–

**Table 4.** Comparison of the changes in the prominent FTIR peaks of DH in HCl-based xerogels.



**Fig. 4.** DSC heating (1) and cooling (2) plots of (a) DH, (b) CHT, (c) PEG, (d) DH + CHT, (e) PEG + CHT, (f) DH + CHT + PEG, and drug-loaded xerogel samples, i.e., (g) X7, (h) X8, (i) X9, (j) X10, (k) X11, (l) X12.

containing PEG xerogels (X9–X12) is due to the presence of PEG in greater quantities, which highlights the thermal gradient (Fig. 4.i–l).

Studying DSC thermograms during the cooling process can sometimes provide valuable information about recrystallization and the physical and chemical characteristics of the active drug. In this study, no peak of any sample during cooling has been noted, demonstrating good miscibility between the drug and the polymers

(Fig. 4.2). Additionally, this suggests minimal likelihood of drug recrystallization upon moisture absorption, as SI studies also revealed minimal values for moisture absorption.

### In-vitro drug release of the xerogels

The dissolution study was carried out to evaluate the DH release rate from different xerogel formulations. Since xerogels are intended to be used orally in periodontal pathologies, the dissolution tool used in this work was simulated saliva at a nearly neutral pH (pH 6.8). It is essential to consider this pH-dependent solubility when formulating DH for pharmaceutical applications. DH exhibits some intrinsic hydrophilicity, with solubility influenced by pH and other ions or molecules in the solution.

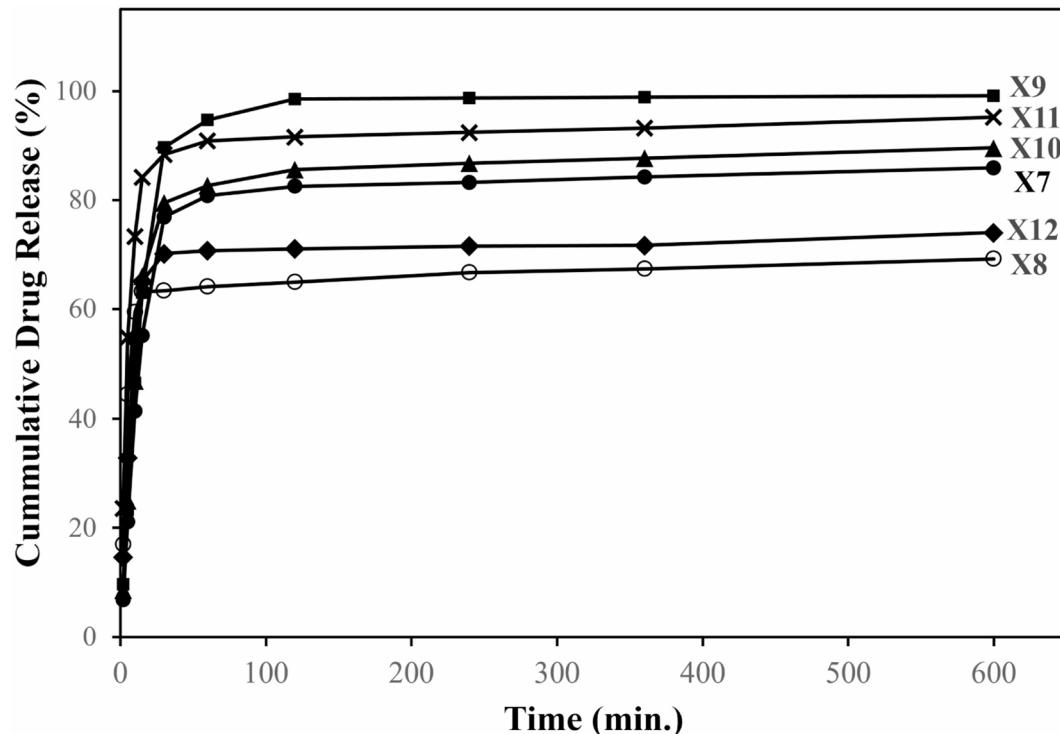
Both CHT and PEG are known to be biodegradable either alone or in combination<sup>60,85–87</sup>. During this study, all drug-loaded xerogels (X7–X12) were dissolved entirely in simulated saliva within 30 min. during dissolution studies. Likewise, the placebo xerogels showed a similar dissolution pattern. These results align with previous reports<sup>85–87</sup>, thus proposing that both CHT-only xerogels (X7 and X8) and CHT-PEG xerogels (X9–X12) will be easily degraded in the periodontal pockets.

In all xerogel formulations, more than 60% of the drug was found to be released during the initial 30 min. (Table S2). Thereafter, a plateau was achieved that showed a consistent slow release of DH till 10 h (Fig. 5). The initial rapid release of DH could be due to its chemical characteristics since it is inherently hydrophilic, with polar functional groups such as hydroxyl (–OH) and amino (–NH) groups. These polar groups can bond hydrogen and water molecules, facilitating DH dissolution in aqueous media<sup>88,89</sup>. The capability of DH to form hydrogen bonding has also been observed and discussed in the FTIR analysis section.

Nevertheless, it is essential to emphasize that the solubility of DH in water varies with pH, with increased solubility in acidic media and reduced solubility in neutral or alkaline media<sup>90,91</sup>. The pH of the gels before drying was acidic, leading to an increased burst of DH release initially. However, over time, the pH effect of the saliva could have impacted the xerogels, leading to a slow release later on. Moreover, the transformation of the drug from crystalline to an amorphous state also leads to an initial rapid release.

The molecular weight of CHT affects the release of DH since it gets slowed with an increase (Table S2). Comparatively, the impact of PEG molecular weight has not influenced the release rate significantly as much as CHT (Table S2). The highest release has been observed in xerogels containing MMW CHT and PEG, while the slowest release has been observed in those containing HMW CHT and PEG (Table S2). However, based on the release profile obtained for DH, it can be envisaged that it will be suitable for clinical use in treating periodontal diseases, particularly where an initial rapid antibiotic release is required.

Different release models were employed to predict the release kinetics of DH from the xerogels. The choice of the most appropriate release model was based on the quality of the fit, represented by the results of  $R^2$  regression (Table 5). It was noted that the cumulative logarithmic release of DH from the xerogels presented a relatively constant relationship with the logarithm of time, which corresponds to the Korsmeyer–Peppas model.



**Fig. 5.** The release profile of DH from various drug-loaded xerogels (X7–X12) at different time intervals (0–600 min.).

Xerogel	Zero-order	First-order	Higuchi model	Hixson–Crowell	Korsmeyer–Peppas
	R <sup>2</sup>				
X7	0.351	0.476	0.552	0.190	0.707
X8	0.223	0.302	0.358	0.101	0.524
X9	0.345	0.618	0.551	0.206	0.735
X10	0.356	0.583	0.851	0.060	0.690
X11	0.251	0.501	0.412	0.145	0.598
X12	0.247	0.316	0.409	0.173	0.614

**Table 5.** Regression values of various models applied for the release of DH from xerogels.

Xerogel	R <sup>2</sup>
X7	0.9493
X8	0.8378
X9	0.9665
X10	0.9629
X11	0.9074
X12	0.9258

**Table 6.** Regression values obtained from the plots of the IVIVC correlation of each sample.

### In-vitro in-vivo correlation (IVIVC)

DH is classified as a Biopharmaceutical Classification System (BCS) Class I drug, characterized by high solubility and high permeability<sup>92,93</sup>. BCS Class I drugs are eligible for biowaivers. IVIVC is based on predictive mathematical models that describe the relationship between an in-vitro property of a dosage form and a relevant in-vivo response, thus avoiding the need to perform in-vivo studies<sup>92,93</sup>. Previously, our group predicted computationally that the DH-loaded CHT-PEG xerogels are highly suitable for localized drug delivery in the periodontium and expected to have no or minimal side effects with negligible systemic ADME. Furthermore, existing local drug delivery systems for periodontal diseases (i.e., Atridox<sup>®</sup>, containing 10% DH) have also demonstrated insignificant systemic absorption (i.e., 0.07–0.1 µg/ml)<sup>94</sup>.

However, based on the availability of in-vivo drug release data of DH from gingival crevicular fluid (GCF) in humans<sup>94</sup>, alongside the in-vitro release profiles obtained in this study, a multiple-level-C IVIVC has been established<sup>95,96</sup> to assess the relationship between the two datasets. A graph was plotted between the log values of in-vivo release data of DH in GCF<sup>94</sup> versus the in-vitro drug release of each drug-loaded sample of this study, and an IVIVC was made<sup>95,96</sup>. The results demonstrated a predominantly strong correlation, as evident from the regression values (Table 6), between in-vivo DH release in GCF and the in-vitro release observed in this study. Based on these findings, it is proposed that DH will have no or negligible systemic effects in the studied xerogel system.

### Antibacterial analysis

The antibacterial activity of both placebo and drug-loaded xerogels was evaluated against the ATCC cultures of *E. coli*, *S. aureus*, *P. aeruginosa*, and *S. mutans*. It has been observed that significant antibacterial activity is present in both placebo and drug-loaded xerogels. CHT is responsible for the antibacterial activity of placebo xerogels<sup>60,97,98</sup>. DH is a versatile antibiotic known for acting against Gram-positive and negative bacteria, including *S. aureus*, *P. aeruginosa*, *S. mutans*, and *E. coli*<sup>99–104</sup>. When comparing the results, it was observed that the activity of the drug-loaded xerogels was more significant than that of the control, as well as that of their placebos against all the tested bacteria (Table 7). It was further confirmed statistically, and a significant difference ( $p < 0.05$ ) was noted between placebo and drug-loaded xerogels when a paired sample two-tailed *t*-test was applied against *E. coli* ( $p = 0.0001$ ), *S. aureus* ( $p = 0.001$ ), *P. aeruginosa* ( $p = 0.011$ ), and *S. mutans* ( $p = 0.00035$ ).

*S. mutans* is considered one of the most critical bacteria forming dental caries. As a result, significant activity against it testifies to the effectiveness of xerogels in treating oral dental conditions. It is possible that the acidic nature of the xerogels played a crucial role in inhibiting the bacteria because their growth is also influenced by the pH of the medium<sup>100,104</sup>.

The MIC<sub>90</sub> (minimum inhibitory concentration for 90% of isolates) of DH against red-complex bacteria that is predominantly responsible for the progression of the periodontal disease is reported to be 0.063 and 0.8 µg/ml<sup>105–108</sup>. In comparison, the MIC<sub>90</sub> values for the bacteria tested in this study are reported<sup>109</sup> as 0.5 µg/ml (*E. coli*), 0.25 µg/ml (*S. aureus*), 4 µg/ml (*P. aeruginosa*), and 3.9 µg/ml (*S. mutans*)<sup>110</sup>. Based on the findings of release and antibacterial analysis of this study, incorporating 1% DH in xerogels (X7–X12) is expected to demonstrate strong clinical efficacy within periodontal pockets. Since the amount released from the xerogels (Table S2) far exceeds the 4 µg/ml threshold, ensuring effective inhibition of pathogenic bacteria. However, further detailed studies in this regard would be beneficial for determining the optimum dose or minimum inhibitory concentration of the drug in the xerogel system.

Samples	Diameter $\pm$ SD (mm)			
	<i>E. coli</i>	<i>S. aureus</i>	<i>P. aeruginosa</i>	<i>S. mutans</i>
DH (Control)	23.25 $\pm$ 0.83	20.00 $\pm$ 3.08	20.50 $\pm$ 1.12	21.25 $\pm$ 0.83
X1	22.67 $\pm$ 2.62	24.00 $\pm$ 1.63	19.00 $\pm$ 1.41	13.67 $\pm$ 0.47
X2	26.00 $\pm$ 0.82	21.33 $\pm$ 1.70	17.00 $\pm$ 0.82	14.00 $\pm$ 0.82
X3	22.33 $\pm$ 0.94	21.67 $\pm$ 0.94	21.33 $\pm$ 2.05	16.67 $\pm$ 0.47
X4	21.33 $\pm$ 1.25	21.00 $\pm$ 4.90	25.00 $\pm$ 2.45	20.00 $\pm$ 4.32
X5	24.00 $\pm$ 1.63	22.33 $\pm$ 0.94	21.00 $\pm$ 3.74	16.67 $\pm$ 1.70
X6	26.00 $\pm$ 0.82	29.67 $\pm$ 2.05	19.33 $\pm$ 2.36	16.67 $\pm$ 2.05
X7	29.67 $\pm$ 1.25	32.67 $\pm$ 2.05	26.00 $\pm$ 1.63	37.33 $\pm$ 1.25
X8	31.33 $\pm$ 1.25	38.67 $\pm$ 1.89	23.33 $\pm$ 2.49	39.33 $\pm$ 0.94
X9	28.67 $\pm$ 0.94	30.33 $\pm$ 0.47	25.00 $\pm$ 2.16	31.67 $\pm$ 1.25
X10	30.67 $\pm$ 1.25	31.00 $\pm$ 0.82	24.33 $\pm$ 1.25	31.33 $\pm$ 1.25
X11	32.33 $\pm$ 2.05	35.67 $\pm$ 4.64	27.67 $\pm$ 3.68	37.33 $\pm$ 1.25
X12	32.67 $\pm$ 1.25	39.00 $\pm$ 1.41	27.00 $\pm$ 0.82	33.67 $\pm$ 1.25

**Table 7.** Inhibition zone of xerogels.

## Conclusion

The placebo and DH-loaded xerogels were developed and characterized using various analytical techniques. The solvent selection is critical in making the xerogels as it may significantly affect the physicochemical properties. The results showed that DH-loaded xerogels displayed different release profiles, having a rapid release of the bulk of the drug in 30 min. The results are promising for the pharmaceutical industry, especially in localized drug delivery within periodontal pockets to treat acute periodontal pathologies. This method can increase bioavailability and reduce systemic side effects, making it a valuable tool for dentists. This approach may improve treatment outcomes while reducing unwanted systemic adverse effects when treating various periodontal conditions. Further studies on stability, cytotoxicity, clinical efficacy, and pharmacokinetics would be of prime importance for commercializing the product.

## Data availability

Data is provided within the manuscript or supplementary information files.

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## Declarations

### Competing interests

The authors declare no competing interests.

### Ethical approval

The study followed the Helsinki guidelines after approval from the Institutional Research and Ethics Board of Baqai Medical University' Ref: BMU-EC/03–2022', and informed consent was obtained from all the participants.

### Additional information

**Supplementary Information** The online version contains supplementary material available at <https://doi.org/10.1038/s41598-025-05894-1>.

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