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# The interaction mechanism of dolomite mineral and phosphoric acid and its impact on the Ti alloy corrosion

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The interaction mechanism of dolomite rock mineral as a novel corrosion inhibitor for titanium alloy in phosphoric acid (0.001–0.5 M) has been investigated for the first time using electrochemical impedance spectroscopy (EIS) and potentiodynamic polarization measurements. Unlike prior studies on synthetic inhibitors, this work uniquely highlights the potential of a naturally abundant, eco-friendly mineral for corrosion mitigation. EIS results reveal a distinctive shift in the inhibition mechanism—from adsorption to diffusion-controlled processes—with increasing phosphoric acid concentration, corroborated by in-situ Raman spectroscopy. This is the first report showing that dolomite effectively reduces the corrosion rate of titanium alloy by forming a stable passive film, with optimal performance observed in low acid concentrations (0.001–0.01 M). Furthermore, at higher acid concentrations (0.1–0.5 M), the mechanism transition influences the film performance. Surface characterization using SEM, EDS, FTIR, XRD, XRF, and in-situ Raman spectroscopy confirms the protective film formation and elucidates the interaction chemistry. The study establishes dolomite as a sustainable alternative to synthetic inhibitors, offering new insights into corrosion protection strategies for titanium alloys.

**Keywords** Dolomite, Phosphoric acid, Corrosion inhibitor, Passive film, Ti alloy, EIS measurements

Titanium and its alloys are very reactive in an aerated environment; a thin titanium dioxide compact, stable, and adherent film is fabricated. This is responsible for their excellent corrosion resistance<sup>1–3</sup>. The superior properties and excellent corrosion resistance of titanium alloys make them ideal for marine usages such as submersible vehicles and in risers in offshore oil drilling<sup>4</sup>. Furthermore, their biocompatibility and acceptable degradation rates enables their use in medical implants and other biomedical applications<sup>5</sup>. Titanium alloys can be utilized as metallic biomaterials due to its acceptable degradation rate<sup>5–9</sup>. Despite these advantages, titanium or its alloys undergo corrosion in violent surroundings forming uniform corrosion and localized corrosion: crevice and pitting, hydrogen embrittlement, stress-corrosion cracking, fretting corrosion and erosion<sup>1</sup>.

Phosphoric acid is a vital industrial acid due to its greater chemical properties and acts as an aggressive medium whereas it poses a significant challenge to the stability of titanium's passive film<sup>10–12</sup>. The acid has been extensively used to study the active passive behavior of titanium and its alloys, shedding light on their electrochemical properties in reducing acid solution<sup>13,14</sup>. Corrosion inhibitors can be utilized to prevent metal corrosion and reduce hydrogen evolution in corrosive media<sup>15–19</sup>.

To combat these issues, various studies of corrosion behavior and the type of passive film have been conducted on titanium and its alloys<sup>20,21</sup>. Organic inhibitors, including heterocyclic molecules, aminoacids, and biopolymers, have demonstrated efficacy in reducing corrosion rates by forming protective films on the alloy surface<sup>14,19</sup>. Inorganic compounds like phosphate based and silicate based inhibitors have also been widely studied for their effectiveness in enhancing the stability of the passive film<sup>22,23</sup>. The electrochemical behavior of titanium was explained in terms of a two-layer passive film structure. Due to its low cationic conductivity and porous outer layer, a barrier film next to the metal prevents the metal from dissolving<sup>13</sup>. Many studies have looked at titanium's active dissolving and passive behavior in mineral acids such as sulfuric, hydrochloric, phosphoric, and perchloric acids<sup>24</sup>. The instability and considerable failure of Ti and its alloys in reducing acid solution have been prompting researchers to investigate the electrochemical behavior of Ti in such a medium<sup>14</sup>.

Dolomite  $\text{CaMg}(\text{CO}_3)_2$ , one of the most abundant rocks, appears to be the cheapest and most readily available raw material. It is well known as double carbonates, which are good sorbents of heavy and nonferrous metals<sup>22</sup>.

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The most common use for dolomite is in the construction industry. Corrosion protection is one of the many uses for dolomite, particularly in water distribution systems. Water filters, pipe linings, and backfill are all made of dolomite<sup>23</sup>. Raw dolomite powder was discovered to be a highly effective adsorbent for removing low levels of phosphates from a variety of water and waste water matrices<sup>25</sup>.

Generally, dolomite and apatites (hydroxylapatite, fluorapatite, francolite, etc.) are the principal phosphate minerals of sedimentary phosphates ores<sup>26</sup>. Amongst these minerals, apatite, calcite, dolomite, barite, fluorite, scheelite, and any salt-type minerals which are highly soluble in aqueous media may be simply separated by flotation from oxides and silicates<sup>27,28</sup>. Therefore, the separation of these minerals are significant. The separation of apatite/calcite was studied previously by depressing apatite and preventing collector's adsorption on its surface<sup>28–30</sup>. Usually, phosphoric acid is used as a depressor of the apatite when the calcite and dolomite are floated using the fatty acid-based reagents as collector<sup>31–34</sup>.

The investigation into dolomite's ability to prevent titanium alloy corrosion in phosphoric acid contributes to the expanding corpus of work on environmentally friendly corrosion prevention techniques. Prior research has examined a range of synthetic and organic inhibitors, emphasizing the importance of material selection for environmental sustainability and efficacy. Bhat et al.<sup>35–39</sup> investigated the use of organic coatings and electrochemical methods and demonstrated significant promise in preventing corrosion on Zn-Ni or Zn-Fe alloys. In the creation of protective films in dolomite, Bhat et al. analysis the relationship between coating thickness and corrosion resistance may be helpful which provided information on how protective films functioned in different situations. He also emphasizes the role of material properties and environmental factors in corrosion behavior and made a transition from adsorption to diffusion-controlled inhibition with increasing acid concentration. Whereas, these findings align with the broader discussion on the role of inhibitors in altering surface reactions and extending the lifespan of materials exposed to aggressive environments. Moreover, He provides further context on the challenges and advantages of natural versus synthetic corrosion inhibitors, underlining the importance of developing cost-effective and eco-friendly alternatives, a concept central to your study's innovation.

To fill this research gap, we will investigate how dolomite interacts with the surfaces of titanium alloys and phosphoric acid to provide corrosion resistance. Herein, it is wanted to study the mechanism of the reaction for dolomite with phosphoric acid and titanium alloy surface to protect it from corrosion. For our knowledge, no corrosion studies have been done using dolomite mineral as an inhibitor for Ti-alloy in phosphoric acid solution. The aim of this work is to evaluate the performance of dolomite material as a natural environmental inhibitor for Ti-6Al-4 V alloy in phosphoric acid with different concentrations (0.001–0.5 M).

## Materials and methods

Dolomite quarries have been obtained from Cairo – Suez road, Egypt. After that, the rock has been crushed, grounded and then grinded to be fine using mortar; and then it is used. Dolomite composition done by XRF is given in Table 1A.

Ti-6Al-4 V alloy was utilized with a cross-sectional area of 0.196 cm<sup>2</sup>. It is manufactured by Johnson and Matthey (England) with elemental compositions given in Table 1B.

The electrode was successively polished with SiC papers (600–2400 grit); and then washed with triply distilled water; and cleaned in an ultrasonic bath containing acetone; and then left to dry in air. A conventional three-electrode cell containing Ti-6Al-4 V alloy as working electrode (WE), a platinum sheet as a counter electrode (CE) and Ag/AgCl as the reference electrode (RE) is used.

The working electrode (Ti alloy) is immersed in 100 mL solution of water and/or different acid concentrations (0.001 M to 0.5 M) with and without dolomite mineral. The amount of mineral as an inhibitor added is 3 g (3%). It has been added in solution with agitation of 245 rpm for 1 h of immersion.

The test solution was prepared from phosphoric acid in different concentrations 0.001–0.5 M (Sigma Aldrich) with or without 3% dolomite inhibitor. All solutions were manufactured by triple distilled water.

Electrochemical measurements were performed utilizing EC-Lab SP 150 Potentiostat electrochemical unit. Potentiodynamic polarization experiments were achieved at a scan rate of 1 mV/s. Before beginning the potentiodynamic measurements, the open-circuit potentials (OCP) were stabilized for 60 min. Electrochemical impedance spectroscopy (EIS) was performed in the frequency range of 100mHz–100 kHz, with a perturbation amplitude of 5 mV. All experiments were repeated two or three times to have reproducible results. A Quanta 250 FEG (Field Emission Gun) scanning electron microscope (SEM) with an Energy Dispersive X-ray spectroscopy (EDX) unit (FEI business, Netherlands) was employed. Also, GeoRessources - SCMEM Tescan VEGA3 was used. A IR-Affinity-1, serial A21374701135S1 made in Japan from Shimadzu corporation used to measure FTIR. PANanalytical X-Ray Diffraction equipment model X-Pert PRO with secondary Monochromator, Cu-radiation ( $\lambda = 1.542\text{\AA}$ ) at 45 kV, 35 mA. and scanning speed 0.04°/s were used.

Mg	Al	Si	P	S	Cl	K	Ca	Ti	Mn	Fe	Zn
(A)											
8.2	0.9	2.2	0.4	0.7	2.1	0.3	47.5	0.03	0.2	1.44	2.18
Al	V	Fe	C	O		N		Ti			
(B)											
5.7	3.85	0.18	0.038	0.106	0.035		Balance				

**Table 1.** Chemical composition of (A) dolomite by XRF analysis and (B) Ti-6Al-4 V.

A Philips X-ray fluorescence equipment, model Philips PW/2404, with Rh target and six analyzing crystals has been used for determining major and trace elements. Crystals (LIF – 200), (LIF – 220) were used for estimating Ca, Fe, K, Ti, Mn and other trace elements from Nickel to Uranium while crystal (TIAP) was used for determining Mg and Na. Crystal (Ge) was used for estimating P and crystal (PET) for determining Si and Al and PXI for determining Na and Mg. The concentration of the analyzed elements was determined by using software Kernal X-44.

The pH of the solution was measured by a pH meter (Hanna) after calibration with standard solutions of pH 4.01, 7.00 and 10.0.

Raman spectra were collected at wavenumber in the range of 200–4000  $\text{cm}^{-1}$  by RXN4 spectrometer (Kaiser) at a laser excitation wave-length of 532 nm, with a fiber-optic Raman sampling probe. The sapphire-head probe gives Raman peaks at 377, 417, 447, 576 and 749  $\text{cm}^{-1}$ .

The Raman spectroscopy experiment was done for 3 g of the sample of dolomite in 100 mL of deionized water and/or phosphoric acid solutions at concentrations of 0.001, 0.01, 0.1, 0.25 and 0.5 M, and stirred at 200–300 rpm for 1 h.

The dolomite sample chemical and mineralogical compositions are determined to check its purity. The XRD (Fig. 1) analysis shows that the sample is relatively pure and doesn't contain other minerals in significant proportion except for a small amount of quartz and tremolite.

The dolomite raw samples' primary peaks were indexed using the dolomite  $\text{MgCa}(\text{CO}_3)_2$  reference pattern (JCPDS Card No. 075-1761). With regards to the unprocessed dolomite samples, it was verified that rhombohedral dolomite was present with reflections at  $2\theta \sim 31.02^\circ$ ,  $33.68^\circ$ ,  $35.43^\circ$ ,  $43.93^\circ$ , and  $52.354^\circ$  which corresponded to the planes (104), (006), (015), (021), and (205).

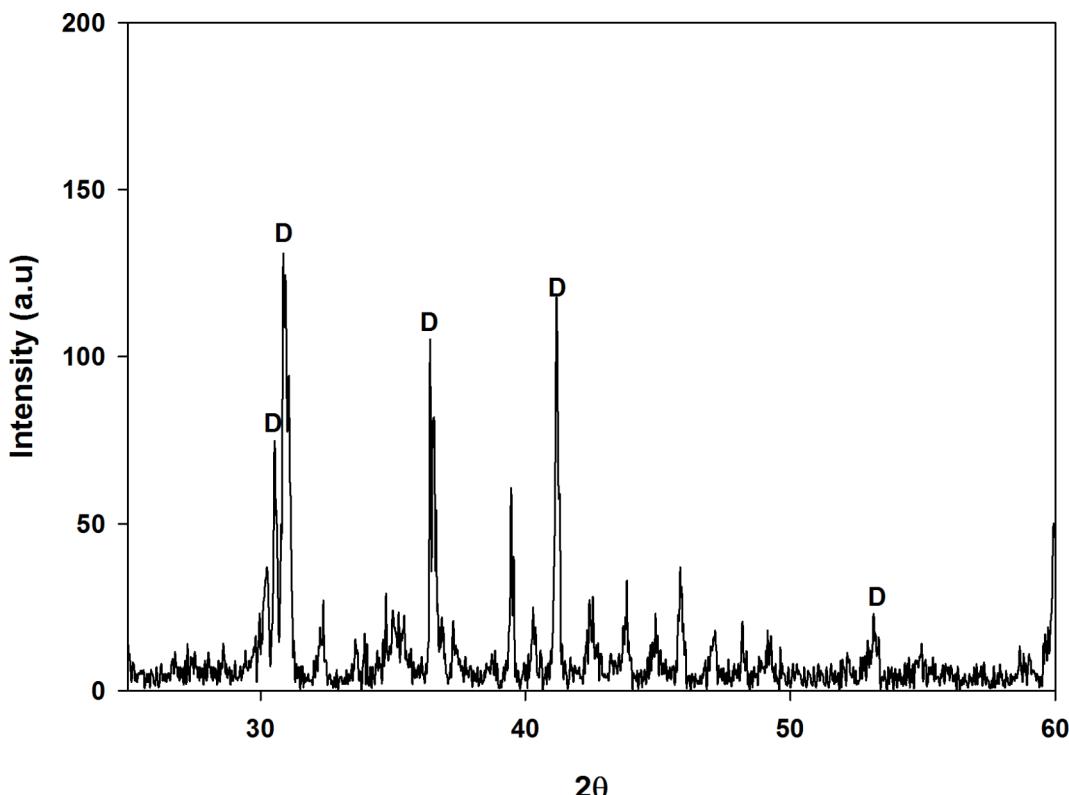
## Results and discussion

### Surfaces morphology and characterization

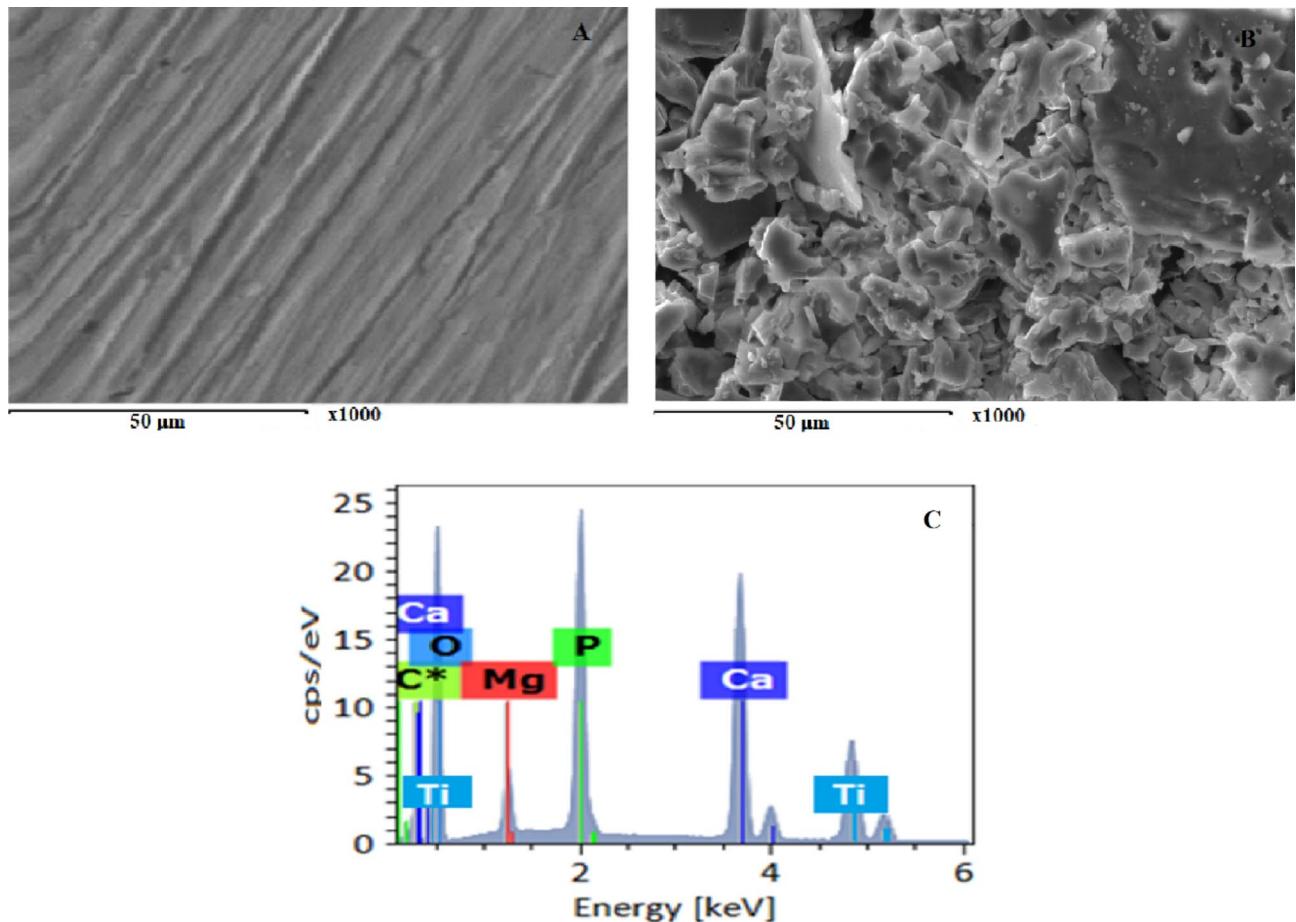
Figure 2A shows the SEM photos of Ti-alloy surface without any addition as a bare electrode<sup>40</sup>. Figure 2B shows the alloy after treatment with dolomite rock mineral in phosphoric acid solution for 1 h at 0.01 M showing the mineral adsorbed film. It seems to be better due to the formation of Ca and Mg phosphate film at this concentration with  $\text{TiO}_2$  film<sup>1–5,32</sup>. Figure 2C is the EDX analysis confirming the presence of all elements forming this film.

### Electrochemical impedance measurements

Electrochemical impedance spectroscopy (EIS) technique was performed for Ti-6Al-4 V alloy in both pure water and phosphoric acid of different concentrations at 1 h of immersion without dolomite, respectively and the results are presented in Fig. 3A, B.



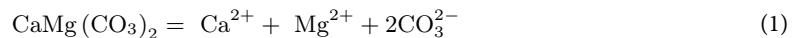
**Fig. 1.** XRD analysis of dolomite sample.



**Fig. 2.** SEM images of (A) Ti alloy surface<sup>40</sup> and (B) after adding dolomite sample in 0.01 M phosphoric acid solution after 1 h of immersion with its EDX analysis (C).

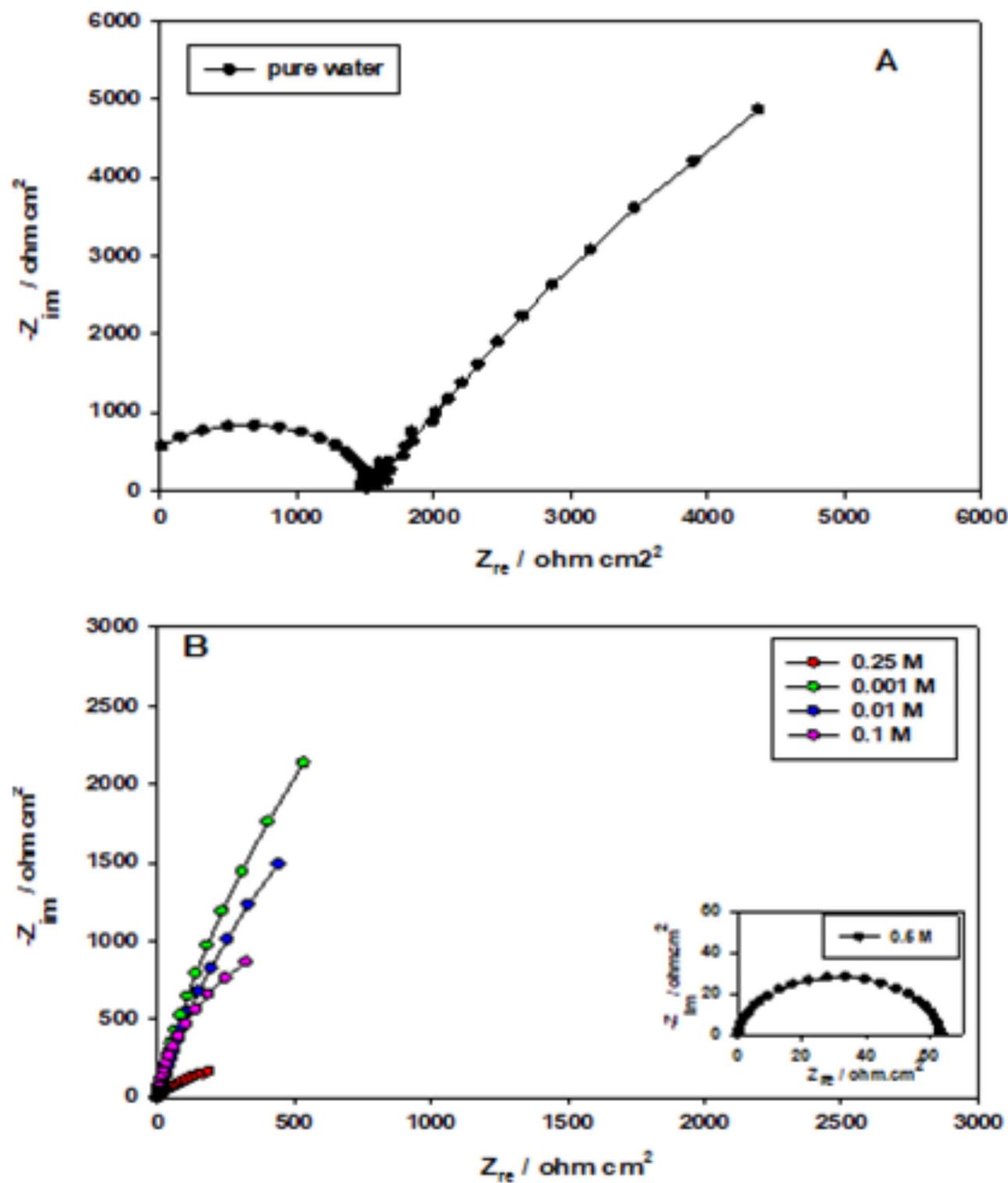
The alloy in water works with both adsorption (semicircle) and diffusion (line) mechanism, however, on phosphoric acid, it is only adsorption mechanism (semicircle)<sup>34,41</sup>. The impedance value in acidic medium decreases than in water due to its violence.

The Nyquist plots (Fig. 4) after adding dolomite show higher impedance value than without dolomite with immersion time owing to the dissolution and precipitation of dolomite in water as carbonate and thickening of passive  $\text{TiO}_2$  film on the alloy surface<sup>1,6,42–45</sup>.



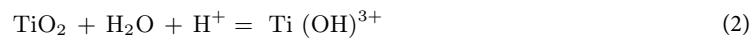
Dolomite dissolves with release of  $\text{Ca}^{2+}$  and  $\text{Mg}^{2+}$  cations that cause precipitates to be developed on the alloy surface, forming a protective coating of either Mg or Ca phosphates depending on phosphoric acid concentration. Hard water with high calcium and magnesium content is less corrosive than soft water forming protective coating<sup>45,46</sup>. Here is the same case where water containing dolomite has high calcium and magnesium content that are adsorbed on alloy surface leading to an increase in impedance value.

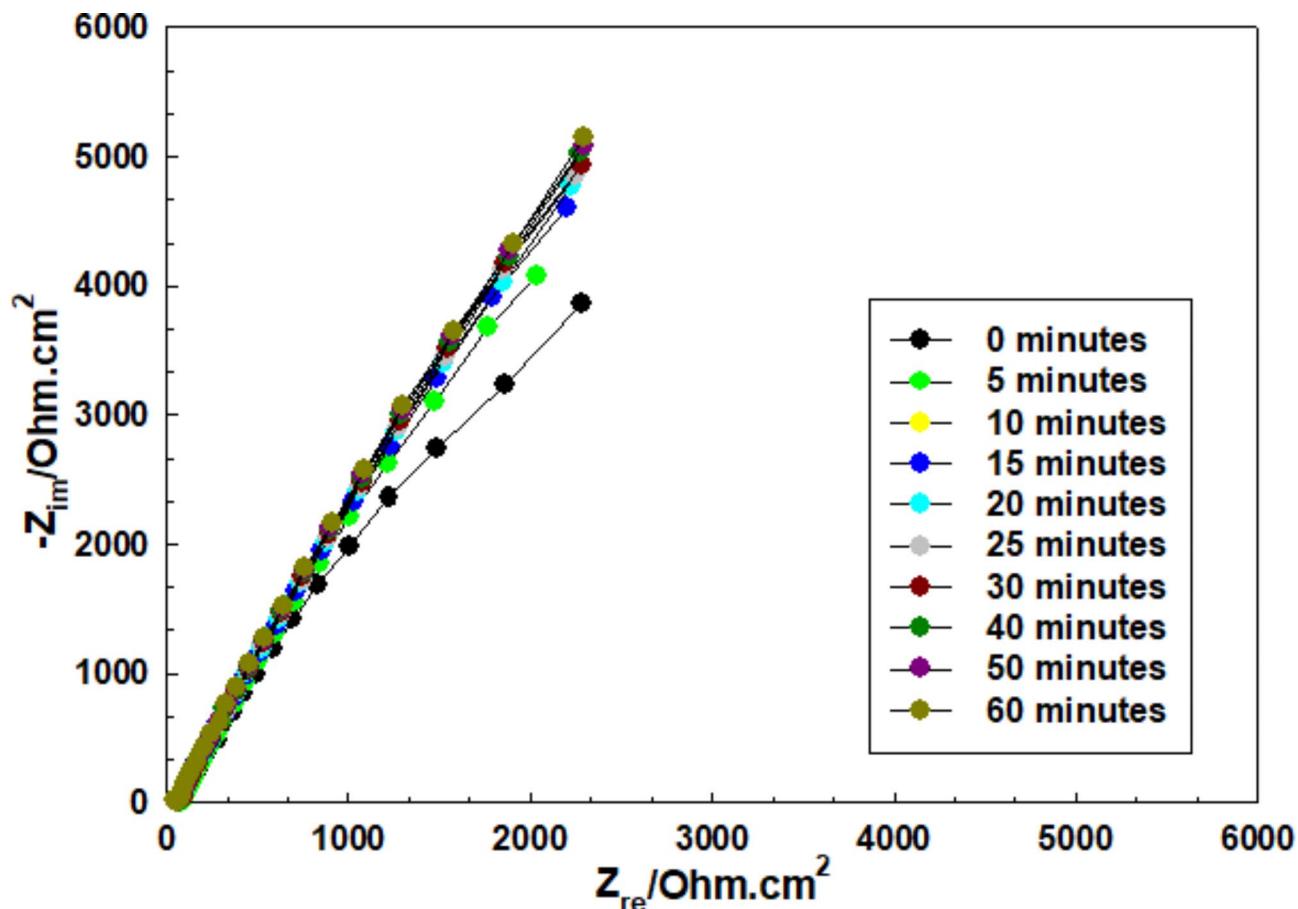
In order to study the protective effect and mechanism of dolomite mineral in phosphoric acid. The impedance diagrams for Ti-6Al-4 V alloy in phosphoric acid with different concentrations (0.001–0.5 M) containing 3% dolomite as a function of immersion time represented in Fig. 5(a–e). The Nyquist plots show the same trend with considerable decrease in impedance value than that of pure water due to the corrosive and aggressiveness effect of phosphoric acid. They show the highest corrosion resistance of dolomite mineral as an inhibitor for Ti-6Al-4 V alloy at the lower  $\text{H}_3\text{PO}_4$  concentration up to 0.01 M indicating the formation of a thicker compact protective film (SEM image in Fig. 2B) by adsorption of dolomite on the alloy surface by forming magnesium and calcium phosphate salts. On further increasing  $\text{H}_3\text{PO}_4$  concentration, the mechanism of adsorption on titanium alloy of undissolved dolomite changes to diffusion and charge transfer resistance. This is due to adsorption resistance decreases and diffusion resistance increases slightly. The film starts to deteriorate at 0.1 M concentration of phosphoric acid due to forming magnesium and calcium phosphate porous film<sup>10</sup>. It is well known that magnesium phosphate is less soluble than calcium phosphate. On increasing acid concentration calcium phosphate mostly dissolve<sup>34</sup>. Also, the compact film converted to porous film with increasing diffusion phenomena and hydrogen evolution rate reaching the top at 0.5 M phosphoric acid solution. Therefore, the



**Fig. 3.** Nyquist plots of Ti-6Al-4 V alloy (A) in pure water and (B) in phosphoric acid of different concentrations at 1 h of immersion without dolomite<sup>2</sup>.

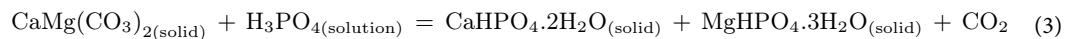
corrosion performance for Ti-6Al-4 V alloy reflects the dissolution of the protective oxide film on the alloy surface as the corrosiveness of the acidic medium is increased. The dissolution reaction for the oxide film on Ti-6Al-4 V alloy in phosphoric acid solution might be written as<sup>6</sup>:





**Fig. 4.** Nyquist plots of Ti-6Al-4 V alloy in pure water with 3% dolomite as a function of immersion time.

The reaction between dolomite and dilute phosphoric acid at room temperature can be described by the equation:



The extent of the reaction depends on  $\text{H}_3\text{PO}_4$  concentration, pH, relative amounts of solid and liquid phases, temperature, and reaction time<sup>10,46</sup>.

The best-fit model for Nyquist plots, used to obtain the parameters in Tables 2 and 3, is the Randles circuit with two time constant. It is represented in Fig. 6 and used for water or phosphoric acid without and with dolomite mineral using EC software. It comprises of a parallel combination of a capacitors ( $C_1, C_2$ ) as a constant phase element (CPE) that is parallel to the surface film's polarization resistance ( $R_1$ ). Warburg impedance ( $W$ ) is due to diffusion mechanism and is connected to the solution resistance ( $R_s$ ). Because of the non-homogeneity of the electrode surface, CPE<sup>34,47–49</sup> is utilized instead of capacitances in the two circuits ( $C_1$  and  $C_2$ ).

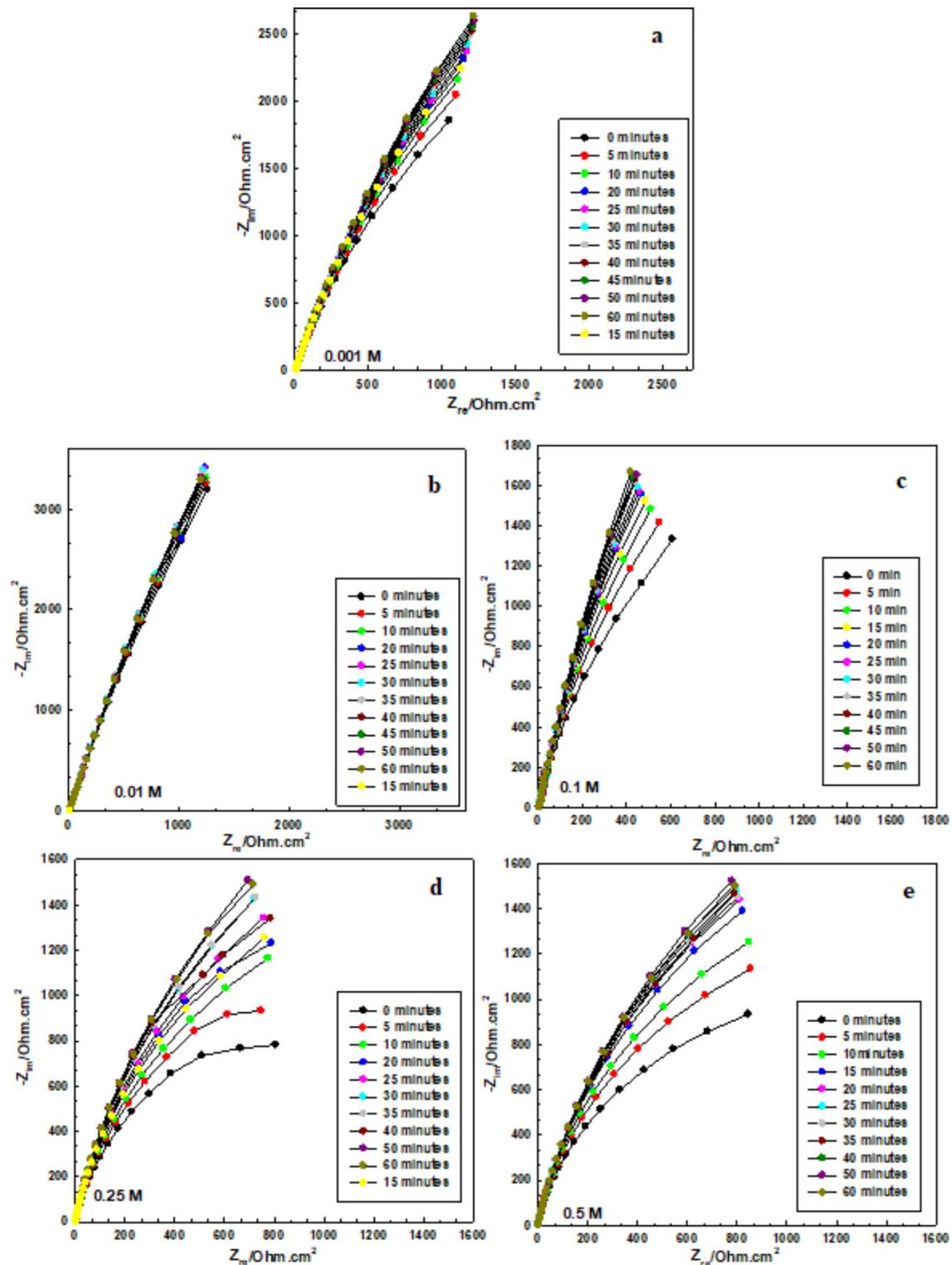
$$Z_{\text{CPE}} = [C (j\omega)^{\alpha}]^{-1} \quad (4)$$

Where  $\alpha$  is an exponent account for surface heterogeneity,  $0 \leq \alpha \leq 1$ ,  $j$  is the imaginary number ( $j = (-1)^{1/2}$ ),  $\omega = 2\pi f$  is the angular frequency in rad/s,  $f$  is the frequency in Hz =  $s^{-1}$ .

The impact of dolomite as an inhibitor is examined in  $\text{H}_3\text{PO}_4$  acid solution of different concentrations (0.001–0.5 M) without and with 3% dolomite as represented in Figs. 3 and 5, respectively. The results showed that the presence of 3% dolomite in phosphoric acid concentrations improved the resistance of the titanium electrode. The corrosion parameters are given in Tables 2 and 3.

This ensures the effectiveness of dolomite rock material that works with adsorption mechanism at low aggressive medium and converts to diffusion especially at high concentration of phosphoric acid giving higher Warburg impedance values. However, even it gives the lowest resistance at higher concentration of phosphoric acid but still working as a good inhibitor (Fig. 7). This is because of forming two films of Ca and Mg phosphate compared to the electrode without dolomite mineral at the same concentration.

Here (Fig. 7) are so clear that the best concentration is 0.01 M concentration after 1 h of immersion, it is better than water starting at 25 min of immersion. Generally, all the films are stabilized after 10 to 20 min. Also, the resistance at 0.25 M and 0.5 M nearly reaches to the same resistance starting from 15 min to 1 h of



**Fig. 5. (a-e).** Nyquist plots of Ti-6Al-4 V alloy in phosphoric acid of different concentrations (0.001-0.5 M) containing 3% dolomite as a function of immersion time.

immersion. The same is for the capacitance value which is reversible proportional to the resistance confirming that the thickness of the film increases with time at all concentrations.

#### Raman kinetic analysis of dolomite

Impedance results can be confirmed by Insitu-Raman spectroscopy. As shown in Fig. 8, It is seen, Raman spectra of phosphoric acid with different concentrations without dolomite.

Conc. M	$R_s$ $\Omega \text{ cm}^2$	$R_t$ $\Omega \text{ cm}^2$	$C_1$ $\mu\text{F cm}^{-2}$	W $\Omega \text{ cm}^2 \text{ s}^{-1/2}$	$C_2$ $\mu\text{F cm}^{-2}$
0.000	2632	4811	2.30	19.5	3.2
0.001	2102	3562	10.5	17.8	3.7
0.010	362.0	2810	11.9	15.3	4.2
0.100	38.20	1412	13.5	12.5	6.8
0.250	24.10	532.0	19.0	9.30	8.1
0.500	15.10	63.7	23.1	6.50	8.2

**Table 2.** Electrochemical impedance parameters of Ti-6Al-4 V alloy in phosphoric acid of different concentrations (0.001–0.5 M) after immersion for 1 h.

The Raman spectra of the phosphoric acid solutions with various concentrations (without mineral existence) are displayed in Fig. 8. The peaks containing this sign (\*) are the probe peaks that is used for measurements. There are four peaks appear at 501, 890  $\text{cm}^{-1}$ , 1075  $\text{cm}^{-1}$  and 1175  $\text{cm}^{-1}$  which increase in intensity with the acid concentration: the main one is at 890  $\text{cm}^{-1}$  while the others are of lower intensity. These peaks appear only at phosphoric acid concentrations higher than 0.01 M and are related to symmetric vibrations of  $\text{PO}_4$  group<sup>34</sup>.

Raman spectra after adding dolomite mineral to phosphoric acid are displayed in Figs. 9, 10 and 11.

Figure 9 displays the Raman spectra of dolomite after contact with phosphoric acid for 1 h of immersion. The main peak at 876  $\text{cm}^{-1}$  that is related to  $\text{H}_2\text{PO}_4$  vibration, it starts to appear at 0.1 M phosphoric acid concentration as a new peak changed its place from 880 to 870  $\text{cm}^{-1}$  (10  $\text{cm}^{-1}$  displacement change) as seen in the inset. It upsurges in intensity with increasing phosphoric acid concentration upper than 0.1 M (Figs. 10 and 11). Also, the same happened at 501 to 518  $\text{cm}^{-1}$  (17  $\text{cm}^{-1}$  displacement change) like that at 890  $\text{cm}^{-1}$  with immersion time. The peak at 1075  $\text{cm}^{-1}$  is related also to phosphoric acid concentration. In addition, it gives the highest intensity at 0.5 M phosphoric acid due to forming two films of Ca and Mg phosphate. This is due to the strong reaction between phosphoric acid and dolomite mineral at high concentrations. The peak at 1175  $\text{cm}^{-1}$  is broadened with time at all concentrations and disappeared at 0.1 M concentrations or lower. The peak at 1097  $\text{cm}^{-1}$  is due to carbonate vibrations. Peaks with low intensity are detected at 277  $\text{cm}^{-1}$  owing to  $\text{CaO}$  or  $\text{MgO}$  symmetric stretching and formation<sup>34,50</sup> and it helps in protecting alloy surfaces. All other peaks are related to the probe (417, 450, 577, 750  $\text{cm}^{-1}$ ). This ensures the formation of phosphate coating especially at high concentrations and decrease of carbonate peaks at 277 and 1097  $\text{cm}^{-1}$  All of these are due to the change of pH of the solution as given in Table 4.

The alteration of pH with time of a solution comprising dolomite in suspension at various concentrations of phosphoric acid are presented in Fig. 12. Outcomes display that the pH of the solution enlarged quickly in the first minute after adding of the mineral and then stays constant, slightly rises or slightly falls. This is due to adsorption precipitation of the hydrolysis products. Also, pH decreases for all sizes for dolomite with increasing phosphoric acid concentration.

Phosphate minerals are sparingly soluble<sup>51,52</sup> and adding of dolomite to the phosphoric acid solution increases the pH, indicating a proton consumption by the dolomite surface and/or by dolomite dissolution products. Change of pH ( $\Delta\text{pH} = \text{pH}_{15 \text{ min}} - \text{pH}_{t=0}$ , without mineral) given in Fig. 13, or proton consumption on dolomite is maximum at  $10^{-3}$  M and then decreases at 0.1 M followed by sharp minimization. It is the sharper reduction or faster dissolution for fine particles than coarser ones owing to the higher area and more passivation of fine particles. Also, proton consumption on dolomite is reduced by enlarging phosphoric acid concentration due to passivation of the dolomite surface by forming  $\text{Mg}_3(\text{PO}_4)_2$  or  $\text{Ca}_3(\text{PO}_4)_2$ .

Here are SEM images in Fig. 14A-D for samples of Dolomite before and after treatment with phosphoric acid solution. Figure 14A is for pure dolomite without any treatment. Figure 14B is for dolomite mineral after immersion in 0.1 M phosphoric acid for 1 h and doing Raman spectroscopy experiment and then filtration to see the film of Ca or Mg phosphate. The film started to be more porous in 0.25 M concentration (Fig. 14C) and with the highest porosity at 0.5 M concentration (Fig. 14D). These are the same results obtained on the surface of Ti alloy Fig. 2A-C. They confirmed Raman spectroscopy results that the intensity is higher with increasing acid concentration owing to increasing film porosity and thickness. Thus, the porosity decreases the film impedance, however, still a good porous thick film that protects the alloy surface from corrosion.

### Potentiodynamic polarization measurements

To confirm further the EIS results, SEM images and Raman spectroscopy experiments, potentiodynamic polarization technique is utilized to study the Ti-6Al-4 V alloy in  $\text{H}_3\text{PO}_4$  acid solution of different concentrations (0.001–0.5 M) without and with 3% dolomite as represented in Figs. 15 and 16, respectively.

The corrosion potential ( $E_{\text{corr}}$ ) moves towards more negative potential as the concentration of the acid growths (Fig. 15). In addition, corrosion current density ( $I_{\text{corr}}$ ) increased with growing acid concentration. It is clear from Fig. 15 that the addition of 3% dolomite leads to a decrease in the current density value. This means that dolomite inhibitor lowering the corrosion and hydrogen evolution rate by decreasing the cathodic branch protecting the surface due to good adsorption of dolomite on Ti-6Al-4 V alloy in  $\text{H}_3\text{PO}_4$  acid.

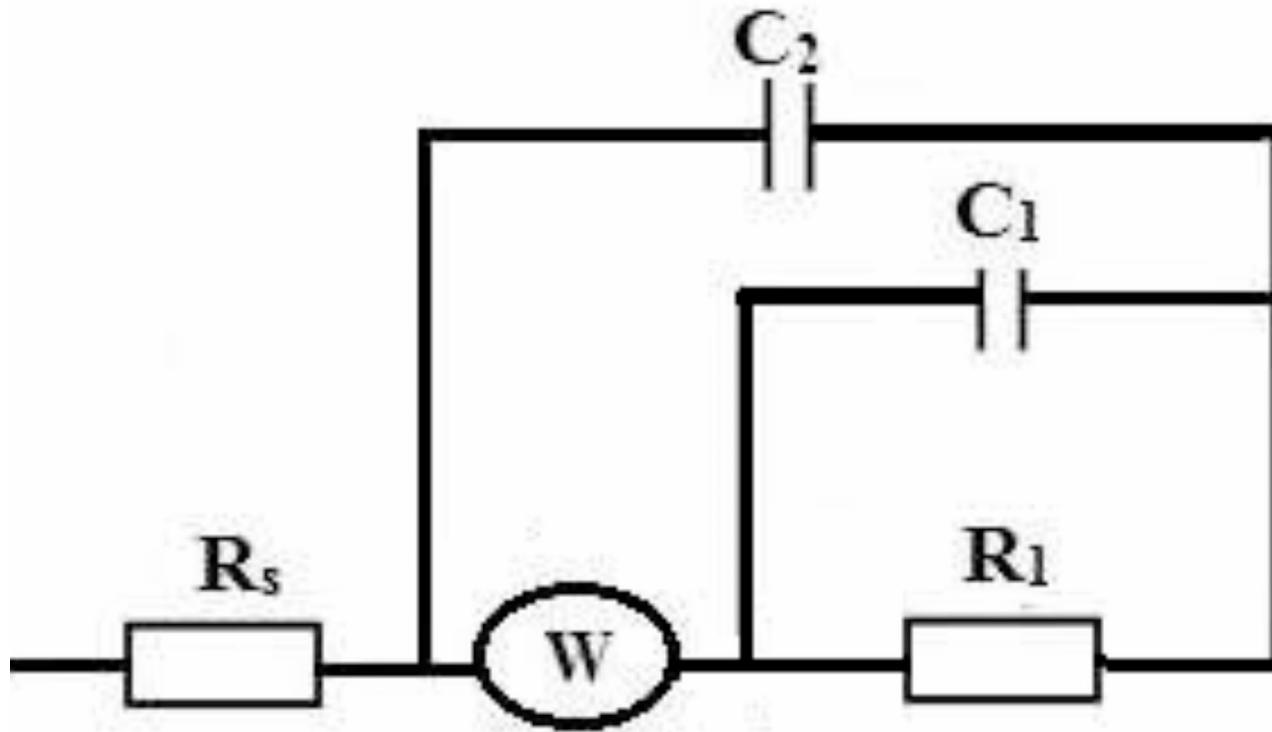
The corrosion inhibition efficiency IE % of the Ti-6Al-4 V alloy in phosphoric acid of different concentrations (0.001–0.5 M) containing 3% dolomite were calculated from corrosion current densities value in Table 5 using the following Eq. (5).

Conc. M	Time	$R_s$ $\Omega \text{ cm}^2$	$R_1$ $\Omega \text{ cm}^2$	$C_1$ $\mu\text{F cm}^{-2}$	W $\Omega \text{ cm}^2 \text{ s}^{-1/2}$	$C_2$ $\mu\text{F cm}^{-2}$
0.00	0	583.44	4811	2.109	0.997	2.112
	5	656.11	5060	2.050	0.874	2.109
	10	660.50	5123	2.030	0.870	2.800
	15	683.12	5132	2.011	0.870	2.790
	20	717.55	5339	1.991	0.867	2.785
	25	753.16	5342	1.983	0.864	2.784
	30	760.13	5411	1.978	0.860	2.781
	40	777.18	5490	1.972	0.856	2.780
	50	789.18	5512	1.965	0.854	2.778
	60	812.14	5631	1.957	0.851	2.775
0.001	0	349.48	3350	5.397	1.348	3.795
	5	356.13	3431	5.106	1.207	3.662
	10	354.29	3540	4.844	1.134	3.597
	15	353.16	3578	4.607	1.055	3.439
	20	427.65	3764	4.469	1.024	3.457
	25	355.64	3899	4.800	1.083	3.496
	30	414.56	4210	4.299	0.978	3.388
	40	419.91	4231	4.401	0.995	3.446
	50	412.37	4335	4.063	0.917	3.323
	60	415.55	4411	4.139	0.929	3.303
0.01	0	397.19	4560	3.474	0.715	3.234
	5	392.39	4934	3.532	0.668	3.241
	10	392.06	5050	3.542	0.666	3.206
	15	393.59	5112	3.571	0.663	3.160
	20	390.01	5225	3.568	0.665	3.204
	25	390.11	5357	3.581	0.650	3.217
	30	391.77	5450	3.597	0.648	3.235
	40	393.36	5649	3.628	0.658	3.228
	50	388.25	5831	3.691	0.672	3.276
	60	396.11	6090	3.691	0.668	3.258
0.1	0	123.67	2533	7.468	1.242	9.956
	5	103.12	2870	9.462	1.608	8.165
	10	109.01	3005	9.309	1.692	8.173
	15	108.69	3232	8.772	1.226	9.129
	20	106.27	3443	9.602	1.402	8.418
	25	107.55	3645	9.202	1.246	8.994
	30	107.39	3650	8.928	1.218	8.913
	40	108.85	3731	9.830	1.305	8.388
	50	114.41	3769	9.115	1.103	8.992
	60	116.01	3921	9.006	1.082	8.906
0.25	0	16.328	1533	8.614	2.724	5.694
	5	16.245	1831	8.597	2.417	5.783
	10	16.293	2211	8.853	1.945	6.149
	15	16.569	2400	8.811	1.808	6.285
	20	18.618	2453	8.802	1.796	6.153
	25	19.823	2530	8.844	1.664	6.203
	30	32.221	2678	8.674	1.569	6.485
	40	32.528	2733	8.807	1.544	6.354
	50	44.632	2831	8.766	1.446	6.409
	60	44.754	2930	8.821	1.458	6.396

Continued

Conc. M	Time	$R_s$ $\Omega \text{ cm}^2$	$R_1$ $\Omega \text{ cm}^2$	$C_1$ $\mu\text{F cm}^{-2}$	W $\Omega \text{ cm}^2 \text{ s}^{-1/2}$	$C_2$ $\mu\text{F cm}^{-2}$
0.5	0	8.0063	2100	8.821	2.341	4.743
	5	5.2945	2234	8.468	2.011	4.966
	10	5.6984	2302	7.855	1.671	5.154
	15	5.1983	2423	8.318	1.853	5.048
	20	3.7443	2450	7.993	1.665	5.229
	25	5.1483	2523	8.247	1.761	5.072
	30	3.8492	2676	7.965	1.598	5.214
	40	4.3773	2717	8.076	1.652	5.168
	50	4.1122	2821	8.007	1.619	5.189
	60	4.1091	2919	7.993	1.611	5.155

**Table 3.** Electrochemical impedance parameters of 3% dolomite in phosphoric acid solution of different concentrations, after immersion for 1 h.

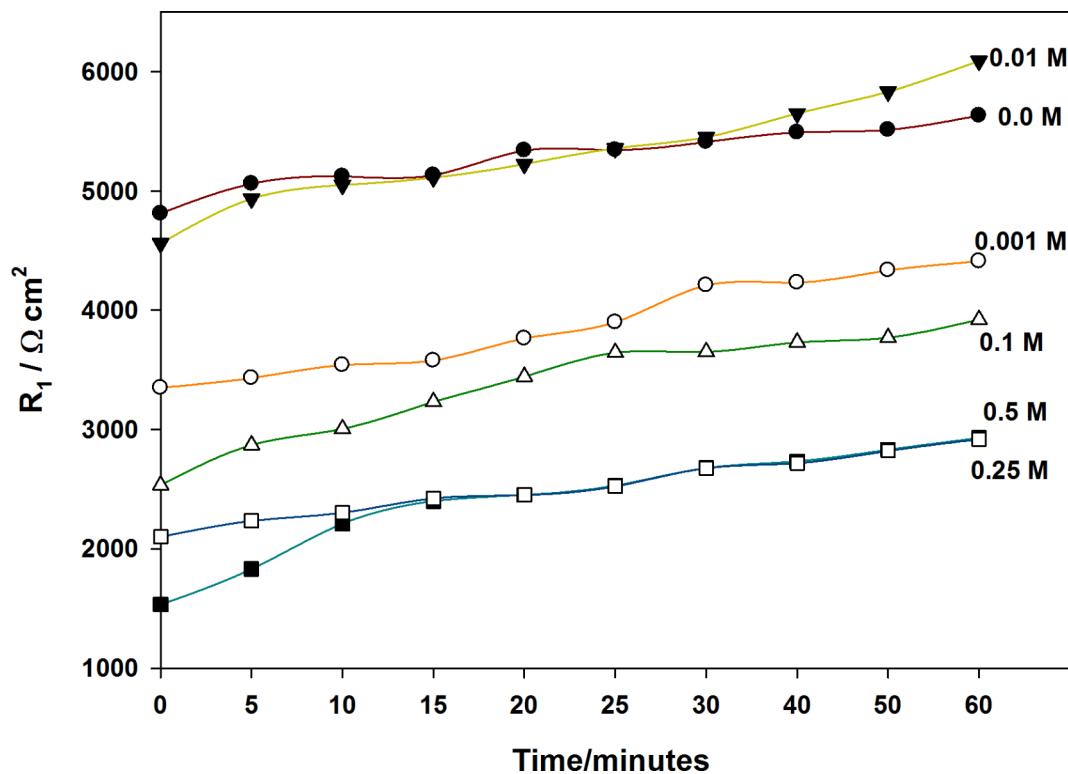


**Fig. 6.** Equivalent circuit used for modeling of experimental EIS data.

$$\text{IE \%} = 1 - \frac{i_{\text{corr}}}{i_{\text{corr}0}} \times 100 \quad (5)$$

Where  $i_{\text{corr}0}$  and  $i_{\text{corr}}$  are the corrosion current densities for the Ti-6Al-4 V alloy without and/or with the dolomite inhibitor, respectively<sup>43,45</sup>. The corrosion parameters derived from the polarization curves such as corrosion potential ( $E_{\text{corr}}$ ), corrosion current density ( $I_{\text{corr}}$ ), and inhibition efficiency IE% were recorded in Table 5.

It was deduced that the highest IE% is 94.3% at 0.01 M and then decreased at higher concentrations but it is all nearly good at all concentrations. This is due to at 0.01 M, the film porosity is the lowest with the more positive potential. Also, due to the more adsorption of dolomite mineral. Thus, dolomite mineral works well as an environmental inhibitor. However, at higher concentration diffusion mechanism changed the film to be more porous, so this decreases the inhibition efficiency with increasing aggressive acid concentration. The inhibitor seems to affect mainly on the anodic branch and so it is an anodic type inhibitor.



**Fig. 7.** The relation between the surface film's polarization resistance ( $R_1$ ) and the time.

#### FTIR analysis for phosphoric acid without and with 3% dolomite

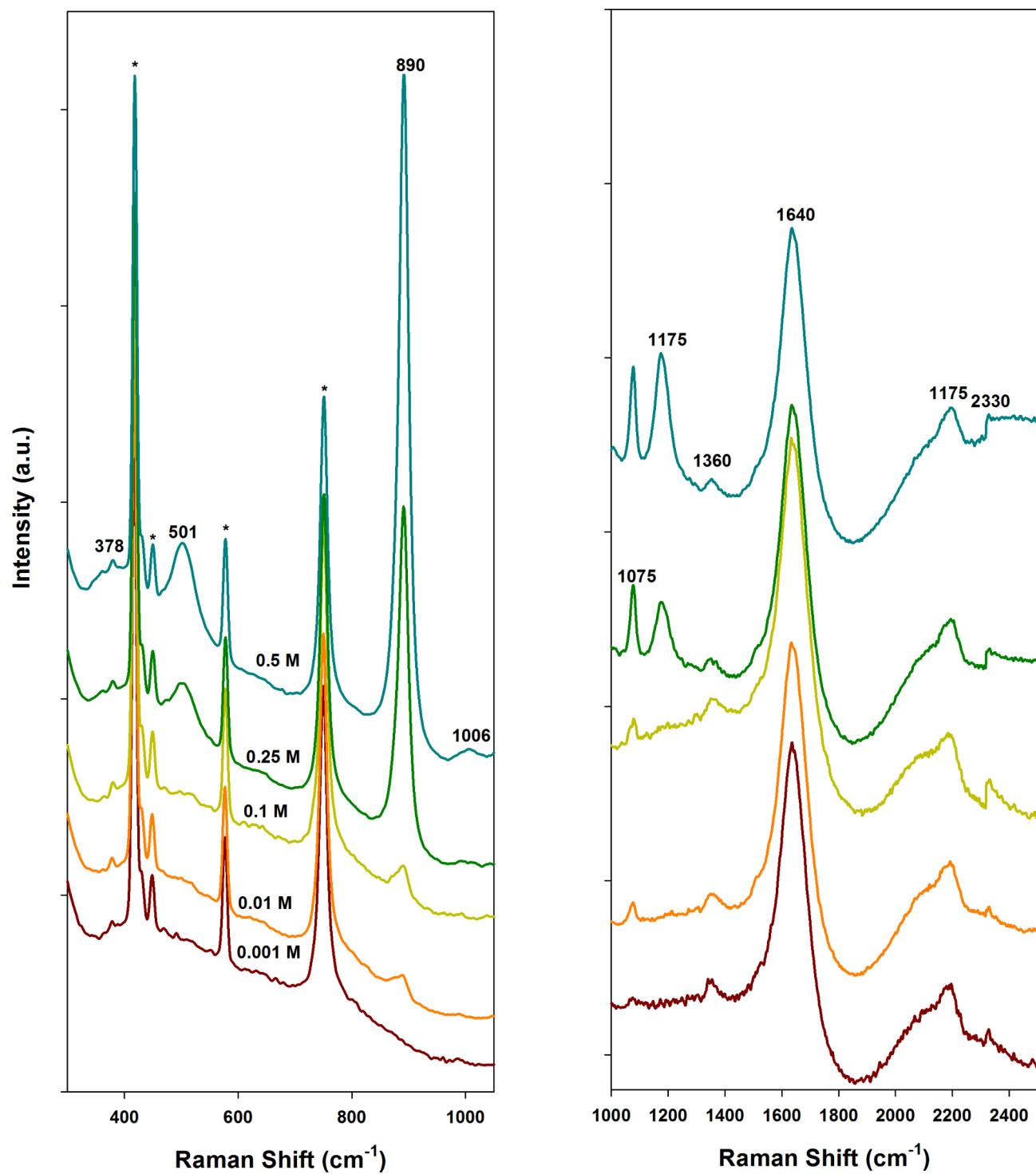
The Fourier transform infrared spectra (FTIR) are shown in Fig. 17, and were recorded to understand the chemical and molecular interactions between 0.5 M phosphoric acid and 3% dolomite. It is clear that after adding dolomite, there are new bands at 894, 748, 540  $\text{cm}^{-1}$  due to magnesium phosphate (struvite, Cattite, Bobierrite; respectively) formation<sup>34,52</sup>. Other bands at 1111 and 1419  $\text{cm}^{-1}$  are due to carbonates. Also, all bands higher than 1600  $\text{cm}^{-1}$  are due to H-O-H bending or stretching. This confirms calcium or magnesium phosphate film formation on reaction between phosphoric acid solution and 3% dolomite.

Finally, all results obtained using all techniques and analyses ensured the excellent protective behavior of dolomite inhibitor for Ti-6Al-4 V alloy in phosphoric acid of different concentrations (0.001-0.5 M) and considered as a novel natural, low cost and widely available material used in industrial applications.

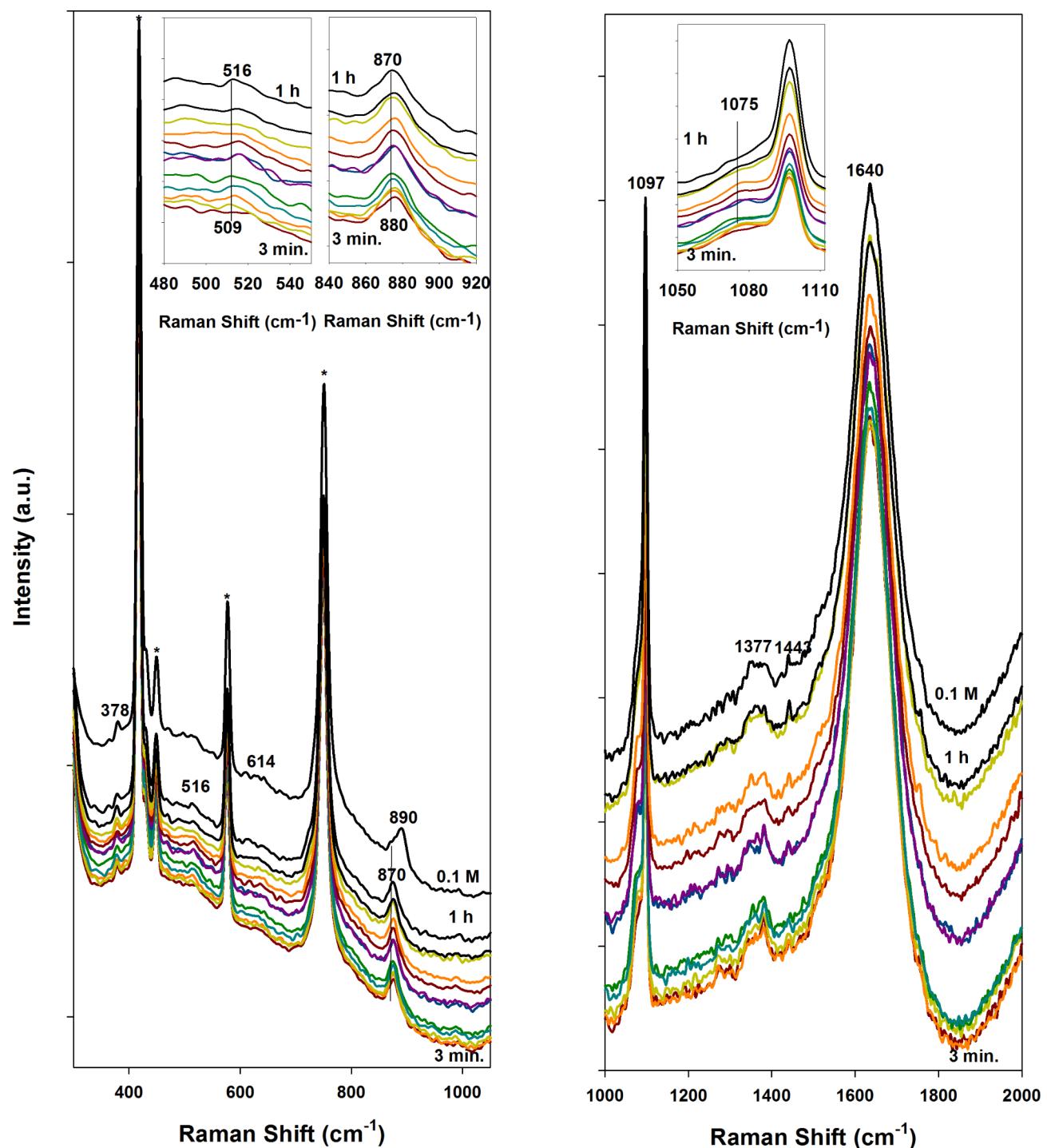
#### Conclusions

Based on the results of potentiodynamic polarization and EIS measurements, the following conclusions are drawn in this study:

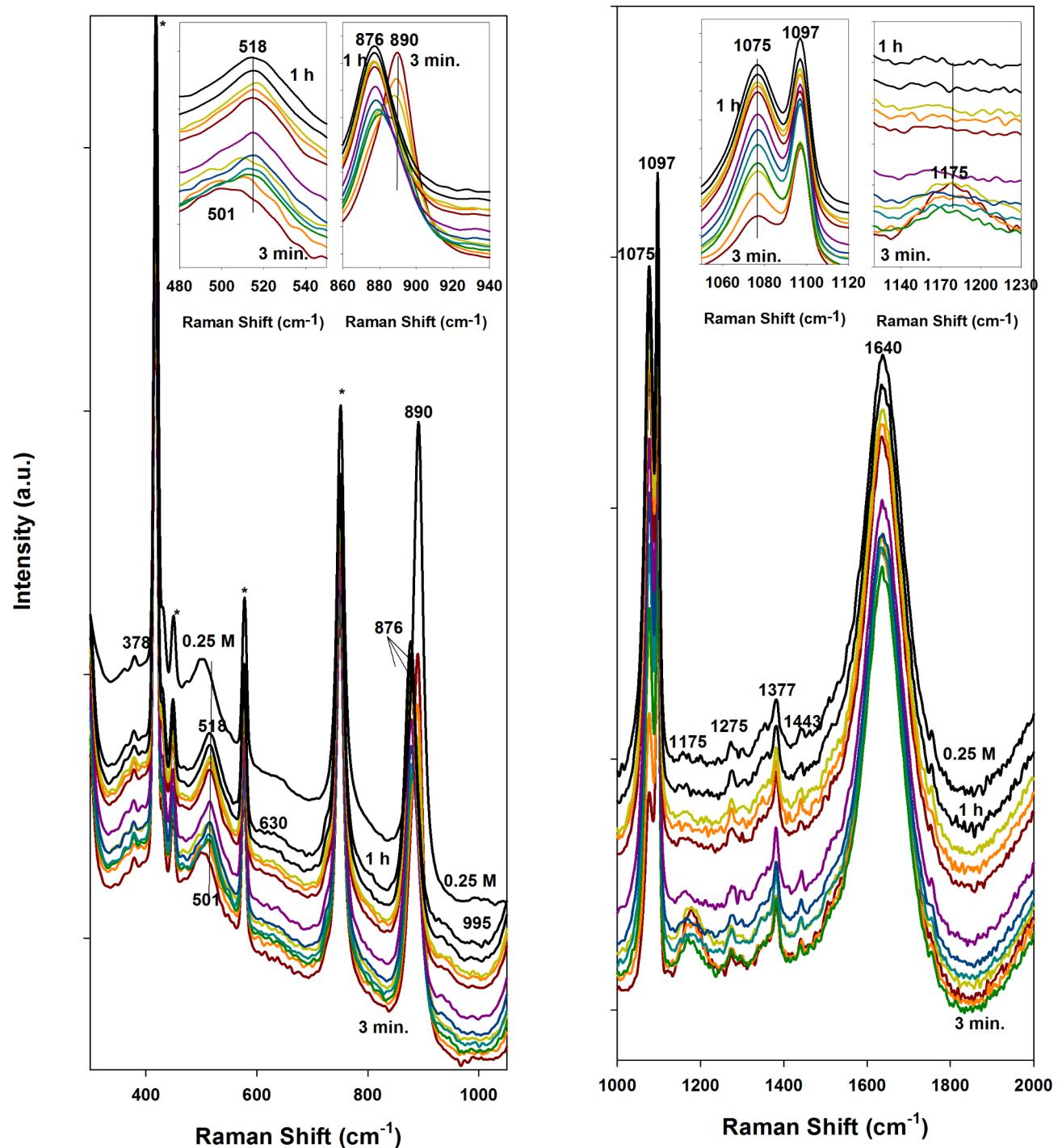
- The electrochemical corrosion resistance of the Ti-6Al-4 V alloy in phosphoric acid increases with immersion time and is the highest at 0.01 M phosphoric acid.
- The corrosion resistance of Ti-6Al-4 V alloy in phosphoric acid shows that dolomite mineral behaves as an effective inhibitor and increases the corrosion resistance of the alloy by protecting it from corrosive medium and forming a protective film on the alloy surface.
- It was found that the inhibition efficiency of dolomite, from polarization technique after adding to water or phosphoric acid, is 92.0% in water, 92.8% at 0.001 M, 94.3% at 0.01 M, 85.8% at 0.1 M, 85.7% at 0.25 M and 85.6% at 0.5 M with the same trend of impedance values obtained from EIS technique.
- This confirms well that dolomite rock material is an effective inhibitor that works with adsorption mechanism at low aggressive medium and converts to diffusion especially at high concentration of phosphoric acid.



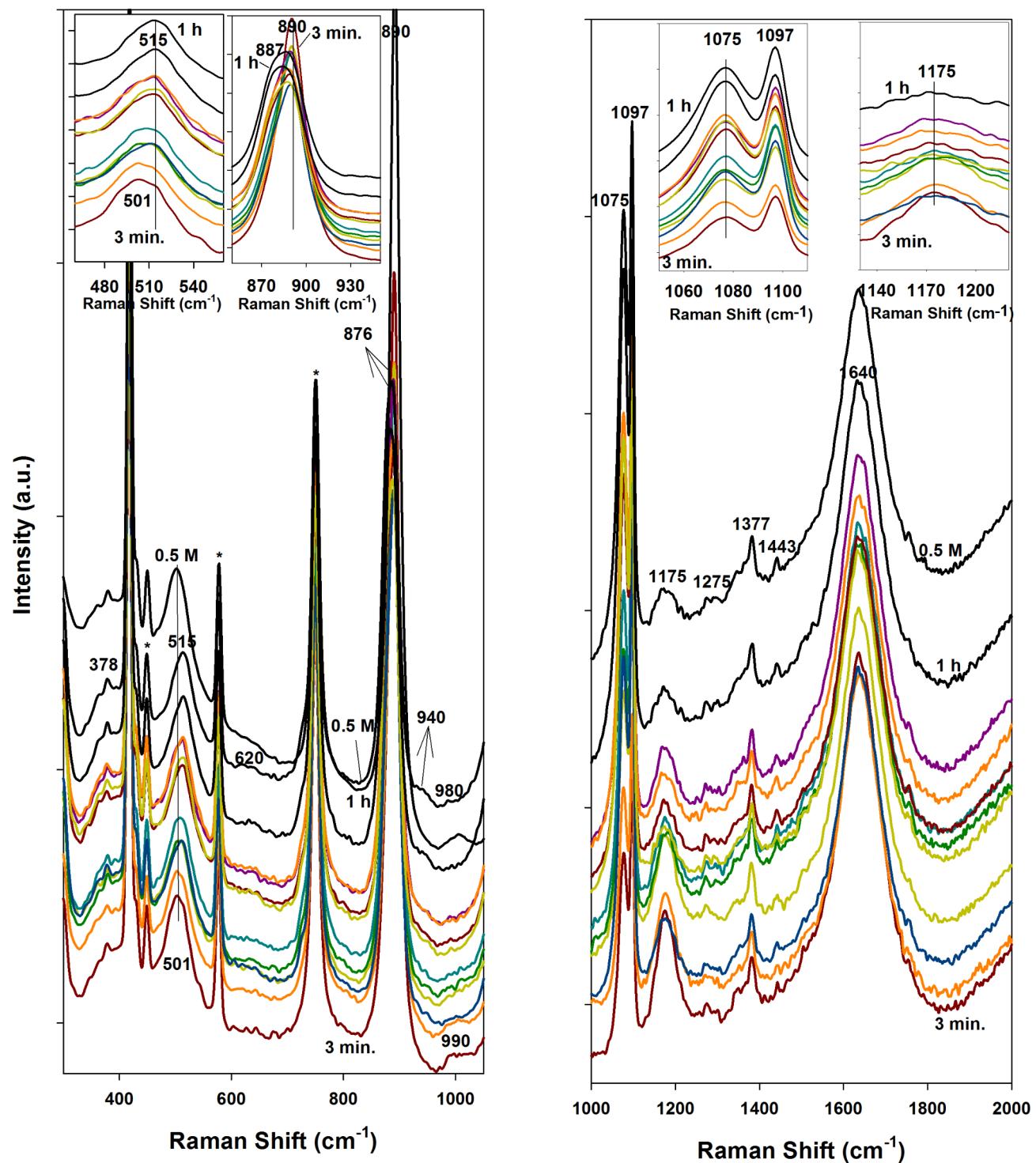
**Fig. 8.** Raman spectra of phosphoric acid in deionized water at different acid concentrations.



**Fig. 9.** Raman spectra of Dolomite with time for 1 h of contact in 0.1 M phosphoric acid solution.



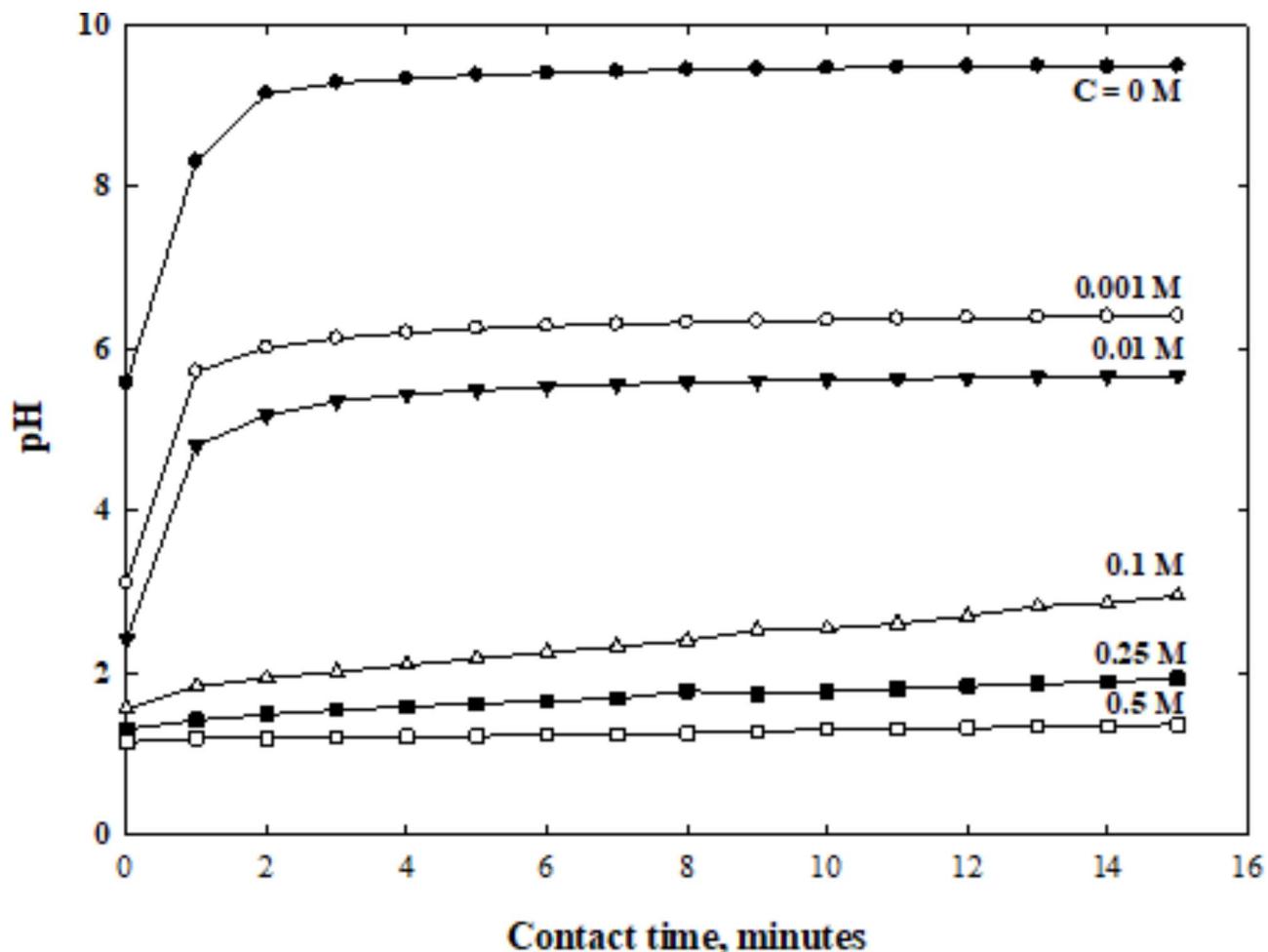
**Fig. 10.** Raman spectra of Dolomite with time for 1 h of contact in 0.25 M phosphoric acid solution.



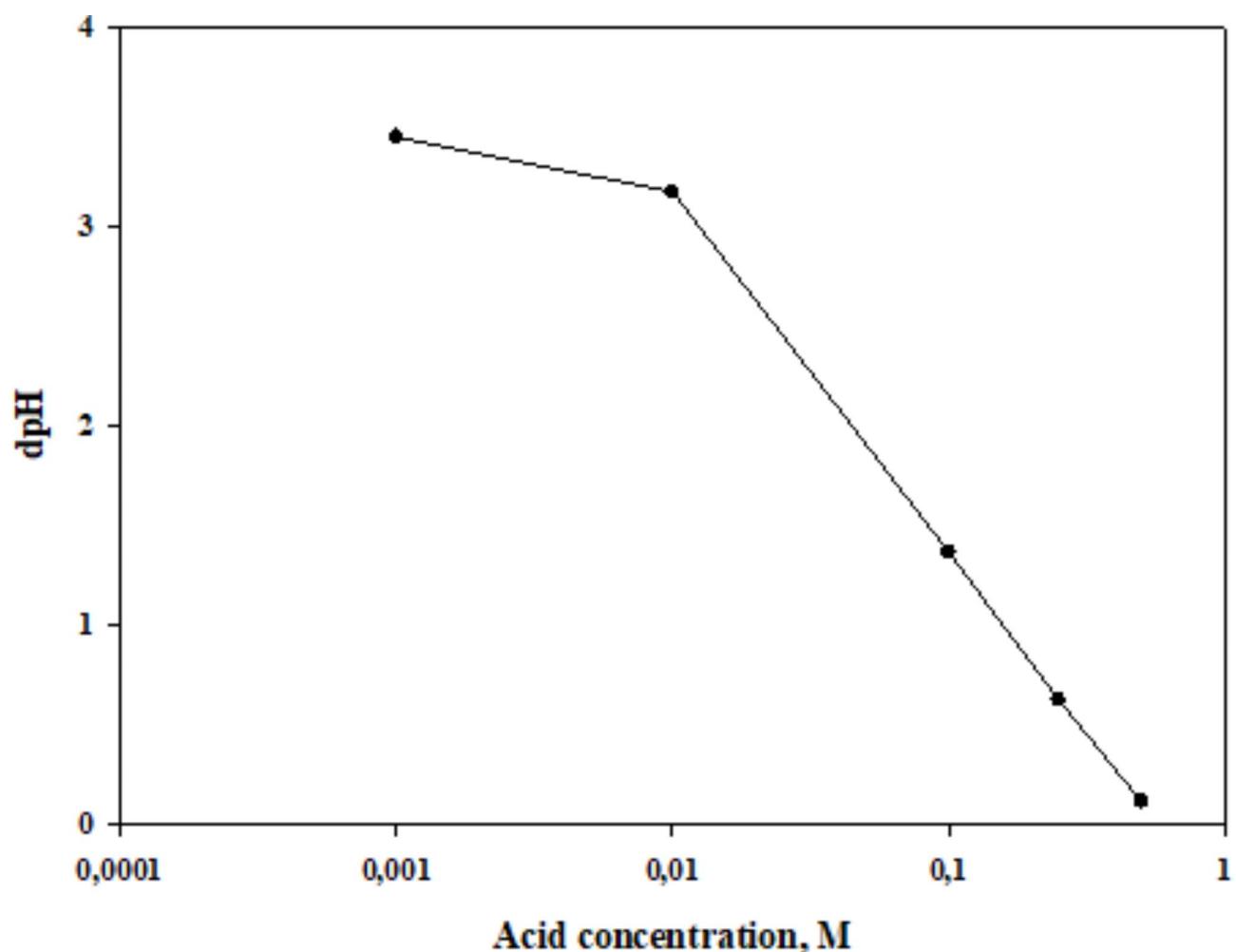
**Fig. 11.** Raman spectra of Dolomite with time for 15 min of contact in 0.5 M phosphoric acid solution.

Phosphoric acid conc. / M	pH	Calcium or magnesium Phosphates formed
0	5.8–9.8	-
0.001	6.9–7.4	$\text{CaHPO}_4^0$ or $\text{MgHPO}_4^0$
0.01	5.7–5.9	$\text{CaH}_2\text{PO}_4^+$ or $\text{MgH}_2\text{PO}_4^+$
0.1	2.4–4.4	$\text{CaH}_2\text{PO}_4^+$ or $\text{MgH}_2\text{PO}_4^+$
0.25	1.6–3.5	$\text{CaH}_2\text{PO}_4^+$ or $\text{MgH}_2\text{PO}_4^+$
0.5	1.4–2.2	$\text{CaH}_2\text{PO}_4^+$ or $\text{MgH}_2\text{PO}_4^+$

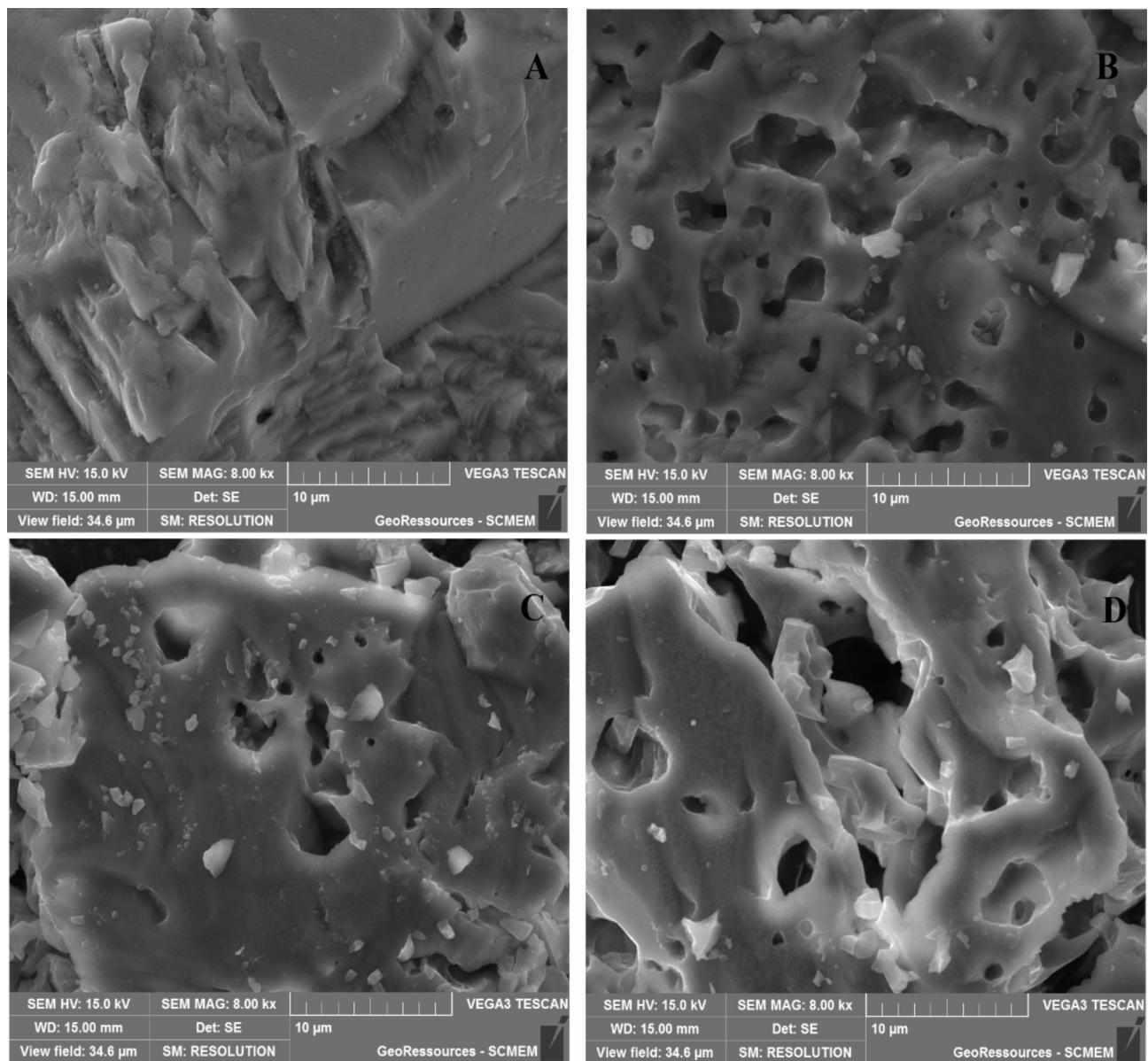
**Table 4.** The compositions formed for both calcium and magnesium phosphates using pourbaix diagram with phosphoric acid pH<sup>40</sup>.



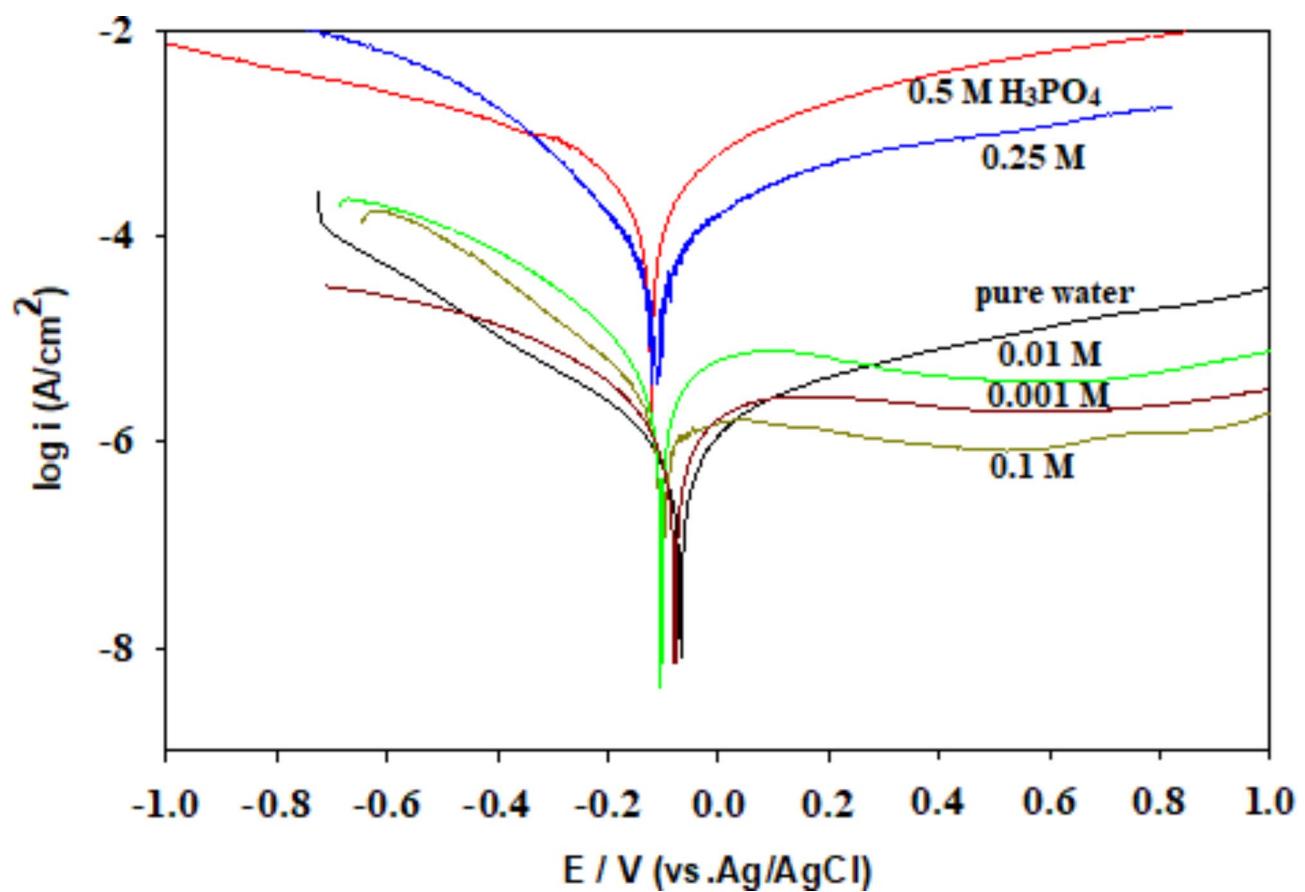
**Fig. 12.** Variation of Dolomite pulp pH with time for 15 min in various phosphoric acid solution concentrations.



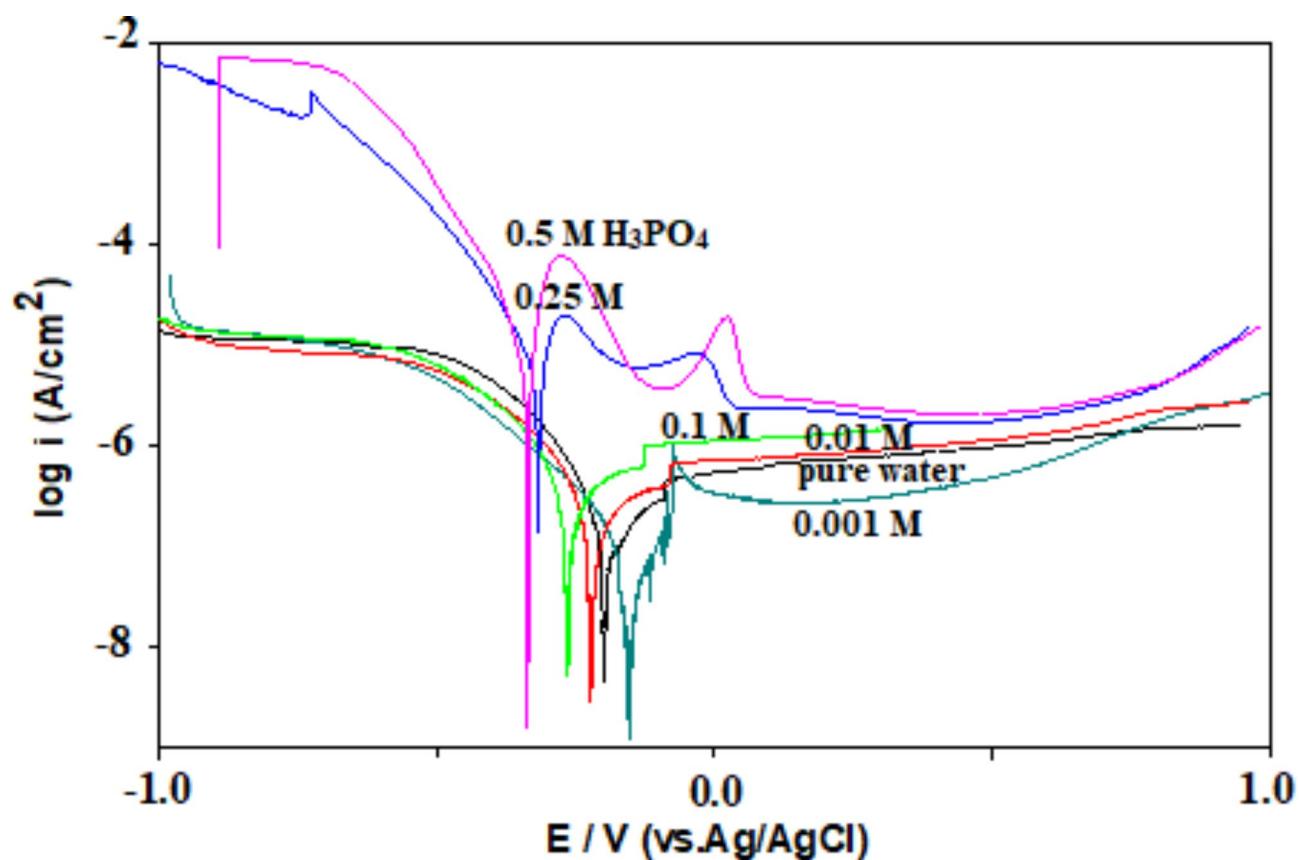
**Fig. 13.** Variation in pulp pH ( $\Delta\text{pH} = \text{pH}_{t=15\text{ min}} - \text{pH}_{t=0, \text{ without mineral}}$ ) for dolomite as a function of phosphoric acid concentrations.



**Fig. 14.** SEM images of (A) Pure Dolomite; (B) in 0.1 M; (C) in 0.25 M and (D) in 0.5 M phosphoric acid solution after 1 h of immersion and filtration.



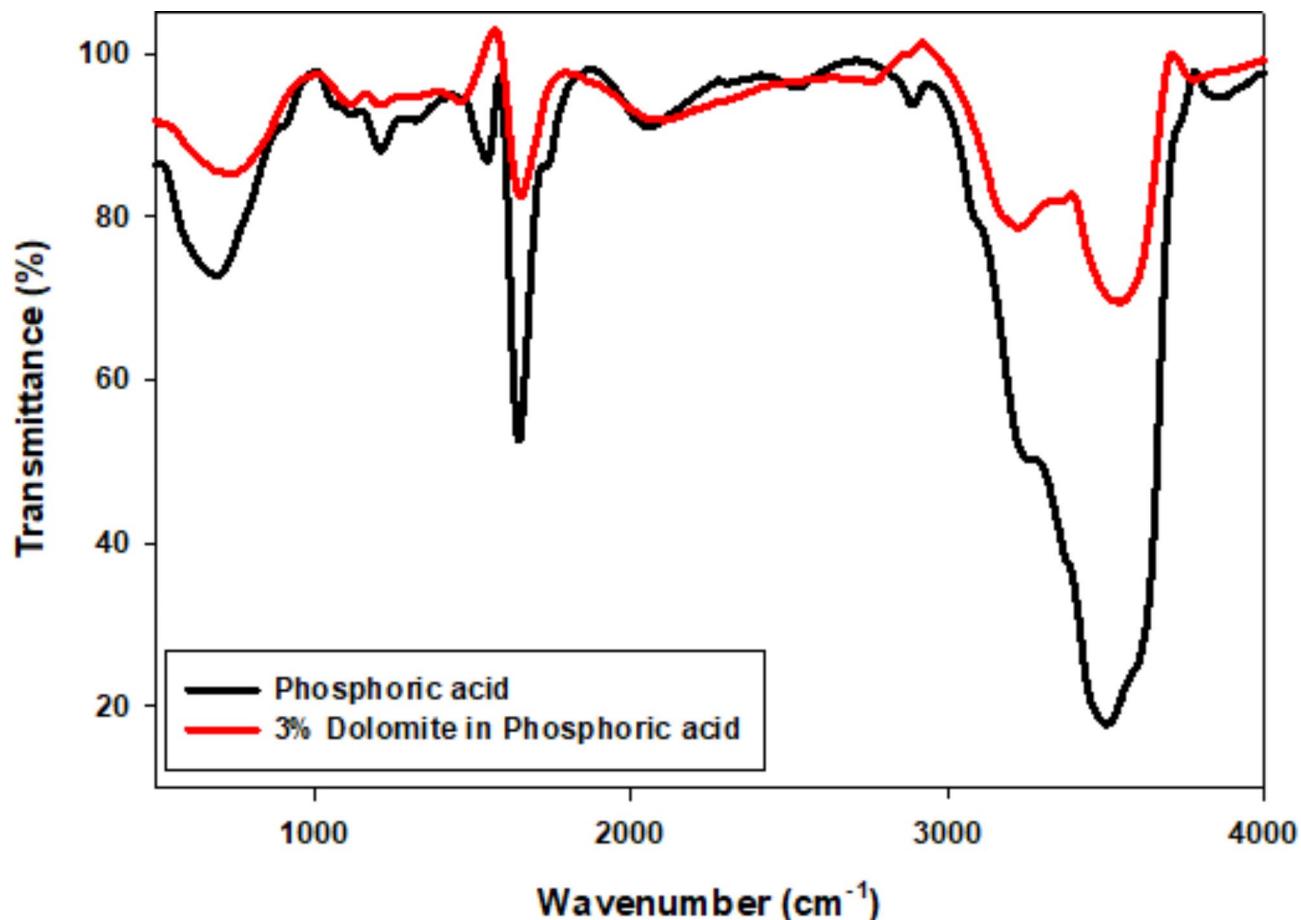
**Fig. 15.** Polarization curves of Ti-6Al-4 V alloy in phosphoric acid of different concentrations (0.001-0.5 M).



**Fig. 16.** polarization curves of Ti-6Al-4 V alloy in phosphoric acid of different concentrations (0.001-0.5 M) containing 3% dolomite.

Conc/M	$E_{corr}$ / mV	$I_{corr}$ / $\mu A/cm^2$	IE%
0.00	-71	1.25	-
0.001	-80	1.10	-
0.01	-104	6.87	-
0.10	-101	2.88	-
0.25	-108	134	-
0.50	-117	219	-
0.00	-188	0.10	92.0
0.001	-150	0.079	92.8
0.01	-230	0.39	94.3
0.10	-250	0.41	85.8
0.25	-305	19.2	85.7
0.50	-331	31.6	85.6

**Table 5.** Electrochemical parameters and inhibition efficiency obtained by polarization tests before and after dolomite addition in acid medium.



**Fig. 17.** FTIR for phosphoric acid without and with 3% dolomite.

### Data availability

The datasets used and/or analysed during the current study available from the corresponding author on reasonable request. (Amany M. Fekry)

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Conceptualization: I.V Filippova, L.O. Filippov, S.S. Medany and A.M. Fekry; Data curation: I.V. Filippova, S.S. Medany and A.M. Fekry; Investigation: S.S. Medany and A.M. Fekry; Methodology: I.V. Filippova, S.S. Medany and A.M. Fekry; Project administration and Supervision: A.M. Fekry and L.O. Filippov; Validation: I.V. Filippova, L.O. Filippov, S.S. Medany and A.M. Fekry; Visualization: I.V. Filippova, S.S. Medany, A.M. Fekry and L.O. Filippov; Roles/Writing - original draft: A.M. Fekry; Writing - review & editing: S.S. Medany, A.M. Fekry, I. Filippova and L.O. Filippov.

## Declarations

### Competing interests

The authors declare no competing interests.

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